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# POTASSIUM SUPEROXIDE CANISTER EVALUATION FOR MANNED SPACE VEHICLES

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TECHNICAL DOCUMENTARY REPORT NO. ASD-TDR-62-583

September 1962

Flight Accessories Laboratory
Aeronautical Systems Division
Air Force Systems Command
Wright-Patterson Air Force Base, Ohio

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Project No. 6146, Task No. 614608



(Prepared under Contract No. AF 33(616)-8323 Los Angeles Division of North American Aviation, Los Angeles, California. Author: A. W. Optican.)



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#### FOREWORD

This report\* summarizes an open duct potassium superoxide (KO<sub>2</sub>) canister test program which was performed during the second phase of work on a study of thermal and atmospheric control systems for manned and unmanned space vehicles. The study was conducted over a period of 18 months by the Space and Information Systems Division (S&ID) of North American Aviation, Inc., under contract AF 33(616)-8323 and was sponsored by the Flight Accessories Laboratory of Aeronautical Systems Division. The Los Angeles Division (LAD) of North American Aviation and AiResearch Manufacturing Company were subcontractors in the study effort.

The purpose of the program is to study the problems associated with the control of temperature and atmospheres in military space vehicles. Several hypothetical vehicles have been analyzed and designed to define potential problems of environmental control. These problems are being solved by analytical and laboratory investigation, and their solutions optimized by the physical integration of components and by other means.

This document represents a part of the total program, procedures, designs, and test associated with the selection of space vehicle thermal and atmospheric control systems based on equipment and physiological criteria. The total effort includes the formulation of several hypothetical space vehicles (including supporting systems), the study and selection of compatible and promising thermal and atmospheric control systems, the development of improved methods of analysis and the presentation of a set of specific examples to illustrate supporting background and data, or present other aspects of the equipment and physiological criteria study program. The other reports issued as a result of the second phase study are as follows:

ASD TR 61-164 \*\* Environmental Control Systems
(PART II) Selection for Unmanned Space
Vehicle.Secret Report

ASD TR 61-240 \*\* Environmental Control Systems
(PART II) Selection for Manned Space
Vehicles, Volume I (unclassified)
and Volume II Secret Report

\*Contractor's number NA-62-283

<sup>\*\*</sup>Unclassified Title

ASD TR 61-161 (PART II)	Space Vehicles Environmental Control Requirements Based on Equipment and Physiological Criteria
ASD TR 61-119 (PART II)	Radiation Heat Transfer Analysis for Space Vehicles
ASD TR 61-30 (PART II)	Space Radiator Analysis and Design
ASD TR 61-176 (PART II)	Integration and Optimization Concepts for Space Vehicle Environmental Control Systems
ASD TR 62- (PART I)	Temperature Control Systems for Space Vehicles
ASD TR 61-162 (PART II)	Analytical Methods for Space Vehicle Atmospheric Control Processes
ASD TR 62- (PART I)	Atmospheric Control Systems for Space Vehicles

During the third phase of the program, the above listed reports will be supplemented as new information is made available. Also, additional reports will be prepared on new subjects of interest for thermal and atmospheric control systems.

The thermal and atmospheric control study program was under the direction of Mr. A. L. Ingelfinger and Mr. A. Gross of the Environment Control Section, Flight Accessories Laboratory. Mr. J. P. Allen of the Environment Control Section acted as monitor of this report, and Mr. R. E. Sexton served as the Project Manager at S & I D.

The author gratefully acknowledges the contributions of the following persons from the Engineering Laboratory in carrying out the test program: J. Dunham, C. Hansen, A. McKinstry, D. Rinehart, L. Martens, and S. Tobey. Appreciation is also expressed to M. A. Sulkin, G. W. Campbell, R. A. Paselk, R. A. Sturgill, and J. B. Truett, of the Aero-Thermo Development Group at LAD, and A. Bialecki of S & I D, for their assistance. All drawings and art work in this report were prepared by E. Rush at LAD.

#### ABSTRACT

New experimental techniques, and their results, used to develop and evaluate the effects of simultaneous multiple operating parameters on potassium superoxide (KO2) canisters for life support systems in manned space vehicles, are presented in this document. Actual experiments using rodents, men, and simulated-man, are described and compared. Experimental characteristics of single and dual canister (demand) atmosphere composition control systems are analyzed. The best method for experimental definitive Respiratory Quotient matching is described. A new annular screen device for a chemical canister which prevents blocking of the airflow in a granular solid chemical bed due to high absolute humidity is shown and described. A canister design method was developed to determine basic KO2 canister sizes for atmosphere composition control system tests on a real-time, one-man basis. dominant role of  $PCO_2$  in establishing the  $CO_2$  adsorption rate and the  $O_2$  generation rate is shown, as well as the roles of absolute humidity and catalysts in establishing the O2 generation rate. Empirical induction was used to try and substantiate (unsuccessfully) a previously accepted theoretical equation of a KO2 bed. A series of comparative tables and curves has been analytically and experimentally derived for 02-C02 generation and adsorption support levels for one

#### PUBLICATION REVIEW

William C. Savage

Chief, Environmental Branch Flight Accessories Laboratory

This report has been reviewed and approved.

FOR THE COMMANDER:

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#### SECTION I

#### INTRODUCTION

This document presents the experimental methods, and results, of an open duct test program conducted at the Los Angeles Division of North American Aviation, Inc. to develop and evaluate potassium superoxide canisters for life support systems in manned space vehicles. The contracted open duct KO<sub>2</sub> canister test program was accomplished between July 1961 and February 1962. Results of earlier experimental programs at the Los Angeles Division of North American Aviation, Inc., using KO<sub>2</sub> and KOH in dual-canister closed-circuit life support systems for animals and men, were utilized to orient this research, and are also discussed herein. These earlier tests were conducted between April 1960 and August 1961.

Under the present contract, an open duct solid chemical test facility was developed at LAD, and the difficulties in instrumentation peculiar to these experiments were solved. A new technique for studying the effects of simultaneous multiple operating parameters on solid chemicals for life support systems has been developed in conjunction with the open duct test facility. The open duct solid chemical test facility makes possible full-scale whole-man tests in real time, allows changing of many operating parameters, and the observance of response rates to changing conditions. It also allows for the simultaneous testing and comparison of three different solid chemical canisters against a standard control canister. Such experiments result in the near-optimum operating conditions becoming known for various canisters and chemicals, and provides rapid design evolution for these canisters. An analytical canister design method was also developed to determine basic KO<sub>2</sub> canister sizes for atmosphere composition control system tests on a real-time, one-man basis.

Under the present contract the program provided a new annular screen design for a solid chemical canister which assures alternative air flow paths around used or fused portions of a chemical bed, and through all unused portions of the chemical bed. Useful lifetime of this new annular screen KO<sub>2</sub> canister was extended at least 250% over that of a straight-through axial flow canister. On a weight basis, the availability of oxygen from the KO<sub>2</sub> canister was increased at least 125% by use of the annular screen device. Whereas this device is mainly intended for use in closed-circuit

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respiratory systems for life support of men in sealed cabins, submarines, pressure suits, fire-fighting, and rescue breathing apparatus, it may also be very useful in any chemical process where a gaseous stream is passed through or over, a solid granular chemical bed. In the successful canister (with the annular screen device), the air-flow was maintained for almost 30 hours with almost no increase in pressure drop across the canister. In KO<sub>2</sub> canisters without the annular screen device, the airflow was completely blocked after a few hours by the formation of a solid plug across the inlet of the canister.

Also, under the present contract, a series of comparative tables and curves have been experimentally derived for 02-C02 generation and adsorption support levels (with KO2) on a one man basis. The role of absolute humidity in establishing the rate of O2 evolution was shown in the reactions of the KO2 The dominant role of CO2 partial pressure in establishing the CO2 adsorption rate, and the O2 generation rate, was experimentally shown by means of the test data. The performance of KO2 canisters when KO2 is mixed with a hydrophilic physical adsorber, a CO<sub>2</sub> chemical adsorber (chemisorption) and/or a porous mechanical separator was shown by means of the test data with many interesting results not previously demonstrated. Also, a well-known theoretical equation showing KO2 alone will match man's R.Q. could not be substantiated by empirical induction from the results of these extensive tests.

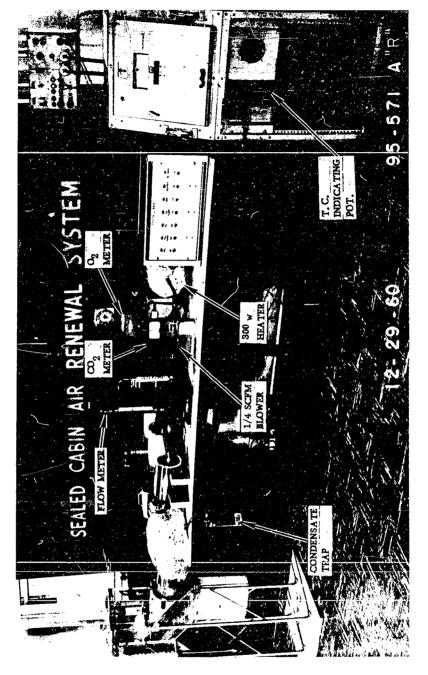
Man-rate support level graphs were plotted with the man-rate production, or adsorption, as a function of time, from the results of integrating (geometrically) the areas under the experimentally derived curves for O<sub>2</sub> generation and CO<sub>2</sub> adsorption. The results show that the ability of a KO<sub>2</sub> canister to support the respiratory requirements of a man is mainly dependent on the bed packing configuration, inlet CO<sub>2</sub> partial pressure, and inlet absolute humidity.

#### BACKGROUND

In 1960, the USAF Ballistic Missile Division, indicated to North American Aviation, Inc., that they were seeking proposals concerned with placing a recoverable monkey-occupied satellite in circular orbit for at least 72 hours. A research program was designed and carried out by the author in 1960 for the evaluation (with animals) of a closed-loop sealed cabin air renewal system utilizing solid chemical

sources. (Reference 1). The results of that experimental tudy demonstrated the ability of a composite solid chemical system (potassium superoxide, potassium hydroxide, and activated charcoal) to renew the air in a sealed cabin with the chemicals in individual canisters. Two female white rats were sustained in excellent health for twenty-five days in a completely sealed enclosure equipped with the solid chemical air renewal system. (Figures 1, 2, 3 and 4). During the twenty-five day test the pressure within the sealed system remained at or near one atmosphere, and the air temperature in the rat enclosure was maintained at 75 ± 9 °F. Control of atmospheric composition (with respect to 02 and CO2) was accomplished with a manual system by limiting the air flow to the potassium superoxide canister in accordance with Oo demand, and increasing a portion of the air flow to the potassium hydroxide canister when the CO2 level rose above the desired limit. A desired relative humidity of 30 to 50% could not be maintained using manual control of a chilled-water heat exchanger, and the average daily relative ... midity gradually increased from about 30% at the start of the test to about 80% at the conclusion of the twenty-five days. Also, intermittent failure of the refrigeration system caused the chemical canisters to inadvertently be used for water removal, which resulted in the formulation of a caustic liquid in the KOH canister. Entrainment of the liquid into the downstream connections and aluminum tubing caused considerable corrosion of the aluminum and a buildup of salt deposits. An example is shown in Figure 5. Airflow was thus restricted and the test was discontinued after the twenty-fifth day.

In September of 1961, at the Los Angeles Division, three men were maintained in a sealed chamber for 14 days as part of an investigation of a life support system of the Space and Information Systems Division of North American Aviation, Inc. (Reference 2 and Figures 6 and 7). During this test, two separate atmospheric composition control systems were utilized. The primary system, consisting of stored gaseous oxygen, and lithium hydroxide for CO<sub>2</sub> removal, operated satisfactorily. The secondary system consisted of the same composite solid chemicals used in the rat test, (KO<sub>2</sub>, KOH, and activated charcoal). This system operated satisfactorily until a caustic liquid formed in the (vertical) KOH bed, drained into the canister sump, and filled it to the overflow point within 8 hours; on cooling in the sump, the liquid crystallized and plugged the drain line causing shutdown of this unit. No liquid was formed in the KO<sub>2</sub> canister. However, a 2-3 inch hard plug of salt crystals formed at the top of the (vertical) KO<sub>2</sub> bed, restricting the airflow, and



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Figure 1 Test Facility, Sealed Cabin Air Renewal System

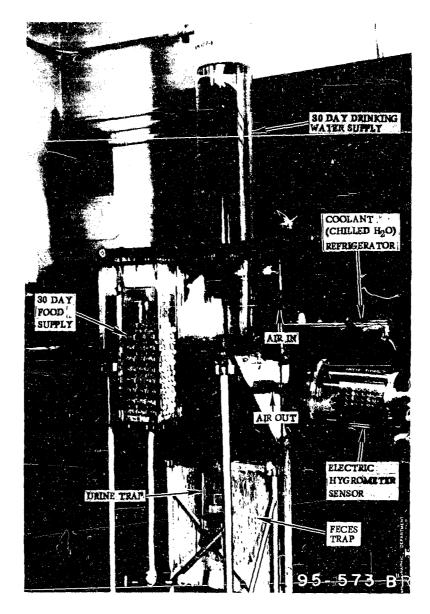
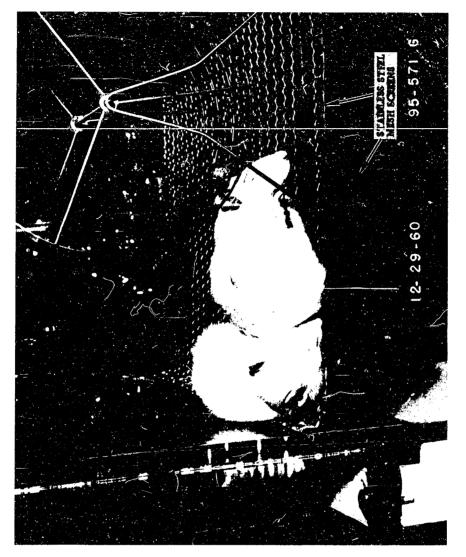


Figure 2 Rat Enclosure Showing Food and Water Water Supply, and Waste Disposal



Rat Enclosure Showing Water Feeder, Exercise Wheel, and Mesh Screens Figure 3

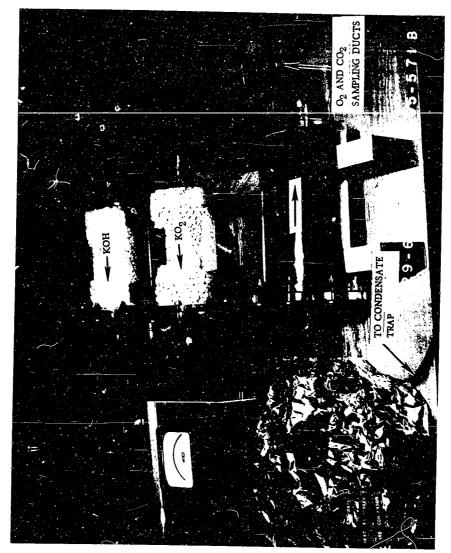


Figure 4 Chemical Canisters for Rat Test

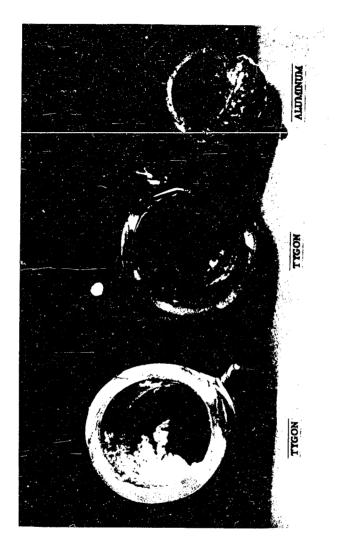


Figure 5 Salt Deposition and Corrosion In Tubing After Rat Test

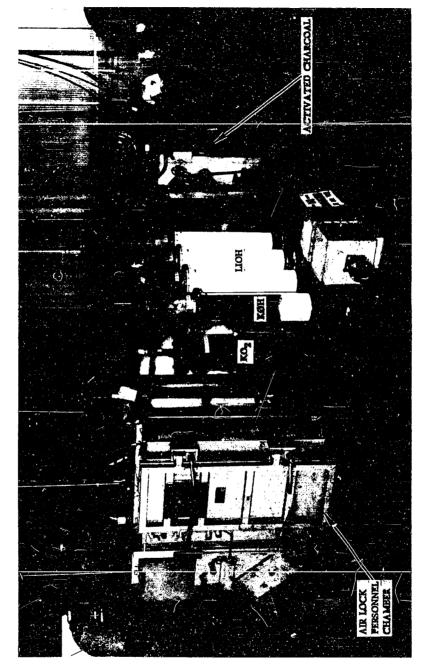


Figure 6 Atmospheric Composition Control System and Personnel Chamber Installation



Figure 7 Interior View of Personnel Chamber

causing shut-down of this unit in about 22 hours. It was thought that the plugging of these canisters was due to excessive moisture content of the air passing through the canisters, and further it was believed that the  ${\rm CO_2}$  level in the system was so low that there was insufficient  ${\rm CO_2}$  available to convert all the very hygroscopic KOH (produced in the  ${\rm KO_2}$  bed from the reaction between  ${\rm KO_2}$  and  ${\rm H_2O}$ ) to carbonate.

### NATURE OF THE PROBLEM

Whereas the closed-loop sealed cabin air renewal system test with the rats successfully demonstrated the ability of a composite solid chemical system to revitalize the air, it did not demonstrate this capability for a single solid chemical (such as  $KO_2$ ) to both absorb  $CO_2$  and generate  $O_2$  in the proper proportion to match the animal's respiratory quotient. A deliberate attempt was made to maintain constant operating parameters (such as temperature, relative humidity, partial pressures, gas velocity and gas residence time in the canisters). Also, only one KO, canister design was used. Therefore, information on the effects of changing the operating parameters, canister dimensions, and/or bed packing configurations (upon CO2 adsorption and O2 generation characteristics of the KO2 bed) could not be readily obtained. It was also hoped that further tests would show what mechanisms caused the plugging of the KO2 canister in the manned test. This called for investigation of various CO2 concentrations and absolute humidities in the (supply) air to KO2 canisters, and investigation of improved KO2 canister designs.

### ADSORPTION AND ABSORPTION

In this report, the terms adsorption and absorption are used synonymously. In a strict sense, however, absorption is a physical taking up of matter in bulk by other matter, or penetration of substances into the bulk of a solid or a liquid. Adsorption is either physical or chemical (chemisorption). Physical adsorption is reversible by physical forces only (no chemical reaction), and is a characteristic surface or interface phenomenon. Chemisorption is an irreversible adsorption by both physical and chemical forces; reaction occurs between the surface of the adsorbent and the adsorbate. Examples of physically adsorptive materials are molecular sieves and activated alumina. Examples of chemically adsorptive materials are KO<sub>2</sub> and LiOH.

#### SECTION II

#### OBJECTIVES OF THE PROGRAM

There were several objectives to the open duct KO<sub>2</sub> canister test program, as follows:

- 1. To attempt to prove or disprove the feasibility of using a single chemical, KO<sub>2</sub>, for both CO<sub>2</sub> adsorption and O<sub>2</sub> generation in a mol volume ratio which will match man's R.Q. (respiratory quotient).
- 2. To determine operating parameter control requirements necessary for matching man's R.Q. with a single solid chemical, if possible, or with a mixed (single canister) solid chemical system in lieu of a composite (dual canister) solid chemical system.
- 3. To provide experimental data on the effects of operating parameters (temperature, humidity, partial pressures, linear gas velocity, and gas residence time) upon the CO<sub>2</sub> adsorption and O<sub>2</sub> generation characteristics of KO<sub>2</sub> beds.
- 4. To provide test data for the development of design criteria both dimensional and bed packing configuration for a KO<sub>2</sub> canister which will assure alternative air flow paths around used or plugged portions of the chemical bed, and through all unused portions.
- 5. To investigate the performance of KO<sub>2</sub> canisters when the KO<sub>2</sub> is mixed with other chemicals such as a hydrophilic physical adsorber, or a carbon dioxide chemical adsorber.
- 6. To investigate various CO<sub>2</sub> concentrations and absolute humidities in the (supply) air to the KO<sub>2</sub> canisters in an effort to show what mechanisms may have caused plugging of the KO<sub>2</sub> canister in the manned life support system test (Reference 2).

#### SECTION III

#### CONCLUSIONS

The objectives of the open duct KO<sub>2</sub> canister test program were answered as follows:

- 1. The attempt to prove the feasibility of using KO<sub>2</sub> alone for both CO<sub>2</sub> adsorption and O<sub>2</sub> generation in a volume ratio matching man's R.Q. was not successful. However, such a feasibility cannot be ruled out completely because as yet unknown procedures, processes, or catalysts (see page 59) may eventually offer a successful solution.
- 2. It was not possible to determine operating parameter control requirements for matching man's R.Q. with KO<sub>2</sub> alone, or with a mixed solid chemical system (i.e., mixed in a single canister). All indications are, however, that a dual-canister system (see page 16) operating in a closed ecological system (see page 59) would furnish the operating parameter control requirements for matching man's R.Q.
- 3. The test program did provide a great deal of experimental data on the effects of the operating parameters upon the CO<sub>2</sub> adsorption and O<sub>2</sub> generation characteristics of KO<sub>2</sub> beds. (Appendix B).
- 4. The program provided a new design whoth dimensional and bed packing configuration for a solid chemical canister which assures alternative air flow paths around used or fused portions of a chemical bed, and through all unused portions. (See pages 40, 46, 89).
- 5. The performance of KO<sub>2</sub> canisters when KO<sub>2</sub> is mixed with a hydrophilic physical adsorber, a carbon dioxide chemical adsorber (chemisorption), and/or an inert porous mechanical separator (see pages 39 and 40), was investigated with many interesting results not previously demonstrated.

- 6. The test program was successful in showing what mechanisms caused the plugging of the KO<sub>2</sub> canister in the manned life support system test. (See pages 101 and 103).
- 7. An open duct solid chemical test facility was developed, and the difficulties in instrumentation peculiar to these experiments have been solved.
- 8. A new technique for studying the effects of simultaneous multiple operating parameters on solid chemicals for life support systems has been developed.
- 9. A series of comparative tables and curves have been experimentally derived for oxygencarbon dioxide generation and adsorption support levels (with KO<sub>2</sub>) for one man. (Appendix B).
- 10. The dominant role of CO<sub>2</sub> partial pressure in establishing the CO<sub>2</sub> adsorption rate, and the O<sub>2</sub> generation rate, was experimentally shown in the reactions of the KO<sub>2</sub> bed. (See pages 59 and 60).
- 11. The role of absolute humidity in establishing the rate of 0, evolution was shown in the reactions of the KO2 bed. (See pages 59 and 60).
- 12. A well-known theoretical equation showing KO<sub>2</sub> alone will match man's R.Q. could not be substantiated by empirical induction from the results of these extensive tests. (Pages 74, 75, and 76).
- 13. An analytical canister design method was developed to determine basic KO<sub>2</sub> canister sizes for atmosphere composition control system tests on a real-time, one-man basis. (Appendix A).
- 14. A theoretical physico-chemical process for KO<sub>2</sub> with H<sub>2</sub>O and CO<sub>2</sub> was developed. (Page 60).
- A chemical utilization index was developed for KO<sub>2</sub> canisters. (Page 66).

#### Section IV

#### RECOMMENDATIONS

#### ADDITIONAL TEST PROGRAMS

Further test programs are recommended on the basis of the results obtained so far in this experimental development and evaluation of composite solid chemical canisters for atmospheric composition control in manned space vehicles.

# The Open Duct Canister Test Facility

Using the same open duct canister test facility, the following approaches should be taken:

- 1. The effects of several reduced absolute pressures on the performance of KO<sub>2</sub> canisters (the "A" configuration) should be investigated at several levels of temperature, absolute humidity, and \$\mathscr{E}\$ CO<sub>2</sub>. The performance parameters which should be measured are the inlet O<sub>2</sub> and CO<sub>2</sub> concentrations, the inlet absolute humidity, and the static pressures at the inlet and outlet of each canister.
- 2. The effect of various catalysts on the chemical R.Q. of KO<sub>2</sub> should be investigated in an effort to find out whether or not a catalyst can be selected that will control the CO<sub>2</sub> absorption O<sub>2</sub> generation of KO<sub>2</sub> to match man's R.Q.
- 3. Using the existing open duct canister test facility, and a test procedure similar to that of the KO<sub>2</sub> tests, investigations should be made of,
  - (a) the operational parameters of lithium peroxide (Li<sub>2</sub>O<sub>2</sub>) as an O<sub>2</sub> generator and CO<sub>2</sub> absorber for use in atmospheric composition control systems.
  - (b) the feasibility and advantages of combining Li<sub>2</sub>O<sub>2</sub> with lithium superoxide (LiO<sub>2</sub>).

Such tests would be concerned primarily with the acquisition of basic operational data to evaluate performance and R.Q. matching ability of the lithium oxides.

The open duct solid chemical test facility makes possible full-scale whole-man tests in real time, allows changing of many operating parameters, and the observance of response rates to changing conditions. It also allows for the simultaneous testing and comparison of three different solid chemical canisters against a standard control canister. Such experiments result in the near-optimum operating conditions becoming known for various canisters and chemicals, and provides rapid design evolution for these canisters.

## The Sealed Cabin Air Renewal System Test Facility

Using the sealed cabin air renewal system test facility (page  $^4$  ), the following approaches should be taken:

- 1. In order to obtain more definitive R.Q. determinations, a sealed cabin test should be conducted utilizing small animals to demonstrate the feasibility (and reliability) of a dual-canister solid chemical system to support animal life over a long period of time, up to 30 days. This closed system test would utilize the "A" canister configuration (annular screen device) as one canister of the dual-canister atmosphere composition control system. The "A" canister would contain KO<sub>2</sub> with catalyst, supported by the annular screen technique. The other of the two canisters making up the dual-canister system would contain LiOH only.
- 2. If the Li<sub>2</sub>O<sub>2</sub> proves superior to KO<sub>2</sub>, it may be substituted for KO<sub>2</sub> in the dual-canister system above.

#### Section V

#### TEST DESIGN

#### BACKGROUND

A solid chemical system was proposed for atmospheric composition control of a sealed cabin. Compared to gas or cryogenic storage, algal, or chemo-electrolytic regenerative systems, a solid chemical system offers the following operational advantages: lower power consumption, ambient cabin gas pressures, minimum (or no) insulation, low leakage rate, no boil-off, and greater reliability. The open duct KO<sub>2</sub> canister test program was proposed to furnish critical design information needed for KO<sub>2</sub> oxygen generating - carbon dioxide absorbing canister beds intended for long duration use in manned space vehicles.

Potassium superoxide, in common with other solid chemicals used for atmospheric composition control, possesses the advantage of room temperature operation without the requirement of high pressure equipment. Unlike many other solid chemicals, KO<sub>2</sub> may serve a dual function of CO<sub>2</sub> removal and O<sub>2</sub> production. The use of KO<sub>2</sub> for atmospheric control is not new. Units employing KO2 are produced commercially for emergency use in submarines, ships, aircraft, and in mines. Results of the closed-loop sealed cabin air renewal system test suggested the need for more (canister) tests of improved KO, bed packing configurations, with the operating parameters controlled in an attempt to prevent channeling of the airflow through the canister bed, caking of the superoxide, slugging of the ductwork by caustic liquid, and over- or underproduction of oxygen with consequent mismatching of man's R.Q. A survey of the available literature revealed a lack of the information required to answer these design and performance problems. The manned test (Reference 2) further indicated the need for such information.

A number of variables could be expected to effect the performance of a  ${\rm KO}_2$  canister in an atmospheric system. These include:

- 1. Temperature of the inlet air.
- 2. Dewpoint of the inlet air (as a measure of absolute moisture content).
- 3. CO2 concentration at the inlet.

- 4. Air flowrate (bed velocity).
- 5. KO, bed length and diameter (ratio).
- 6. KO, bed-packing configuration.
- 7. Mixture of XO2 with reactive or non-reactive chemicals.
- 8. Grain size of the KO2.
- 9. Ratio of 02 to inert gases in the inlet
- 10. Total (atmospheric) pressure at the inlet.

Due to limitations in time and funds it was decided to evaluate only the effects of items 1 through 7, which were considered to have the strongest influence upon canister performance. If all possible combinations of all ten variables at three levels were considered, about 59,000 tests would have had to be conducted. Seven variables at three levels would still have required over 2,000 individual tests if all possible combinations were tested. Therefore, we reduced (by intuitive judgement) the plan for testing all possible combinations to a reasonable number of individual tests (i.e., 67 tests in the original plan) which we believed would satisfactorily indicate the effects of all test parameters on canister performance. It was decided, at that time, to perform the tests in two phases, i.e., the first phase would consist of 66 short run individual tests (conducted three at a time by paralleling canisters), and the second phase would be a single larger individual canister tested for an extended period of time. After the basic effects of all test parameters were observed in the short run tests, a near-optimum set of operating levels would be selected for the extended duration test of the "best" (selected) canister design. This final canister was to be run to the point of exhaustion to investigate practical design criteria such as channeling, and effective yield of the bed, and to identify problems that might occur during extended operation. During the first phase tests, some of these canisters were to be run to breakthrough. Sequential testing was to be applied throughout, and thus the experimenters could stop after any change of operating conditions, or change of canister design, and examine the accumulated results (to date), before deciding whether to continue the experiment as outlined above, or make changes in the original test plan. In this way, the experimenters were allowed freedom to alter any operating conditions or canister designs (as the test program progressed),

which definitely appeared to be leading us away from nearoptimum operating conditions or near-optimum canister design.

# GENERAL DESCRIPTION

A test was considered to consist of passing room air at controlled levels of temperature and absolute humidity, gas velocity, and CO2 concentration through four KO2 canisters in parallel, for a specified length of time. Three of the canisters were to be test units whose design (bed dimensions and internal packing) were to be varied from one test to the next; the fourth canister was to be a small commerical emergency unit used as a control standard in all tests. First phase testing was to consist of short-run tests of 4 to 5 hours duration, with a portion of these short tests run to exhaustion (about 8 to 10 hours), for various combinations of values for the operating parameters (temperature, dew point, gas velocity, and CO2 concentration). The design parameters (internal packing configuration, and bed length to bed diameter ratio) were to be varied from one test to the next. (See Table 3). Measurements were to be taken intermittently throughout all tests to determine 0, and 00, concentration, humidity, and static pressure at the inlet and outlet of each canister, thereby permitting canister performance to be associated with time as well as with values of the test parameters.

#### THEORETICAL CONSIDERATIONS

The existing literature gives many theoretical reactions of a KO<sub>2</sub> bed with water and carbon dioxide. Some of these are indicated in Table 1. The extent to which these reactions occur is not well known, particularly the hydration of KOH and  $K_2\text{CO}_3$ . The chemical equilibria of an active  $\text{KO}_2$  bed are very complex. The major parameters appear to be the input absolute humidity and % CO<sub>2</sub> concentration, the air mass-flow, and the temperature of the chemical bed. Several of the most likely reactions from Table 1 can be combined into the following overall equation:

$$4 \text{ KO}_2 + 3 \text{ H}_2\text{O} + 2 \text{ CO}_2 = 2 \text{ K}_2\text{CO}_3 \cdot 3 \text{ H}_2\text{O} + 3 \text{ O}_2$$

Water may be absorbed in the  $\mathrm{KO}_2$  bed without equivalent  $\mathrm{O}_2$  evolution, possibly due to hydrate formation. The chemical literature indicates the formation of  $\mathrm{H}_2\mathrm{O}_2$  in a  $\mathrm{KO}_2$  bed provided there is no catalyst present to decompose the peroxide. This peroxide decomposition is highly exothermic

# Table 1 Chemical Reactions of $KO_2$ Bed

$$4 \text{ KO}_{2} + 2 \text{ H}_{2} \text{ O} = 4 \text{ KOH} + 3 \text{ O}_{2}$$

$$2 \text{ KOH} + \text{CO}_{2} = \text{K}_{2} \text{CO}_{3} + \text{H}_{2} \text{O}$$

$$KOH + \text{CO}_{2} = \text{KHCO}_{3}$$

$$KHCO_{3} + \text{KOH} = \text{K}_{2} \text{CO}_{3} + \text{H}_{2} \text{O}$$

$$KOH + 3/4 \text{ H}_{2} \text{O} = \text{KOH} \cdot 3/4 \text{ H}_{2} \text{O}$$

$$KOH + 4/2 \text{O} = \text{KOH} \cdot \text{H}_{2} \text{O}$$

$$KOH + 2 \text{ H}_{2} \text{O} = \text{KOH} \cdot 2 \text{ H}_{2} \text{O}$$

$$2 \text{ KOH} \cdot 2 \text{ H}_{2} \text{O} + \text{CO}_{2} = \text{K}_{2} \text{CO}_{3} \cdot 3/2 \text{ H}_{2} \text{O} + 7/2 \text{ H}_{2} \text{O}$$

$$K_{2} \text{CO}_{3} + 1/2 \text{ H}_{2} \text{O} = \text{K}_{2} \text{CO}_{3} \cdot 1/2 \text{ H}_{2} \text{O}$$

$$k_{2} \text{CO}_{3} + 3/2 \text{ H}_{2} \text{O} = \text{K}_{2} \text{CO}_{3} \cdot 3/2 \text{ H}_{2} \text{O}$$

$$2 \text{ KO}_{2} + 2 \text{ H}_{2} \text{O} = 2 \text{ KOH} + \text{H}_{2} \text{O}_{2} + \text{O}_{2}$$

$$H_{2} \text{O}_{2} \xrightarrow{\text{Cat}} \text{H}_{2} \text{O} + 1/2 \text{ O}_{2}$$

$$K_{2} \text{CO}_{4} + \text{CO}_{2} \xrightarrow{\text{O}-10 \, ^{\circ}\text{C}} \text{ K}_{2} \text{C}_{2} \text{O}_{5} + 1/2 \text{ O}_{2}$$

$$K_{2} \text{CO}_{4} + \text{CO}_{2} \xrightarrow{\text{O}-10 \, ^{\circ}\text{C}} \text{ K}_{2} \text{C}_{2} \text{O}_{6}$$

$$K_{2} \text{CO}_{4} \xrightarrow{\text{heat}} \text{ K}_{2} \text{CO}_{3} + 1/2 \text{ O}_{2}$$

and readily catalyzed by heavy metal oxides. Therefore, to provide a high  $\rm O_2$  generation rate, an  $\rm H_2O_2$  catalyst should be incorporated into the  $\rm KO_2$  bed.

#### Section VI

#### TEST SET-UP AND INSTRUMENTATION

#### DESCRIPTION OF TEST FACILITY AND TEST INSTRUMENTATION

#### Test Set-Up

Three axial blowers were used to supply room air to the test system. These blowers together had a capacity of about 19 SCFM at 18 inches of H<sub>2</sub>O, which was sufficient to compensate for all impedance in the test system, and provide the required airflows through the canisters. The first stage blower was wrapped with a copper coil through which cold tap water flowed carrying away some of the heat of compression. (Figures 8, 9, and 10). Downstream of the third stage blower an aluminum water jacket was welded around the aluminum air duct, and cold tap water passed into, and out of, this cold water jacket further reducing the heat of compression. (Figure 10).

The air next entered the chilled water aircraft-type heat exchanger (Figure 11) where it was cooled to a predetermined dewpoint, and any condensation removed through a condensate drain and trap. The water which circulated through this heat exchanger contained ethylene glycol to depress the freezing point of the liquid coolant (to < 32°F), and prevented freeze-up of the chilled water refrigeration units (Figure 11) which was the cause of the intermittent failure of the refrigeration units in the animal test (page 3). The liquid coolant loop contained a pump, a 3-way valve, and a sump. Originally, the sump was included with the intention of passing liquid CO2 through methyl alcohol (in this sump) for a coolant, but when this idea was discarded (so as not to increase the normal PCO2 in the room) in favor of vapor-cycle (Freon) refrigeration units, the sump thereafter served no real purpose. The 3-way valve was modulated (manually) to provide the required dew-point in tests calling for absolute humidities less than that of the room air. If an absolute humidity was required which was greater than that of the room air, steam was generated at the inlet of the first stage blower. When the absolute humidity was reduced, the excess moisture condensed in the heat exchanger was removed by the trap.

Immediately downstream of the chilled water heat exchanger, CO<sub>2</sub> gas was resulated and metered into the air duct from a pressure bottle. This bottle was charged from a plant liquid CO<sub>2</sub> line to 300 psig before each test run.

Air of the required dew point and % CO<sub>2</sub> concentration was then passed through a 1500 watt finned-strip heater which was activated and regulated by a thermostat as required to provide the proper dry bulb temperature and % relative humidity.

Downstream from the heater was located the sensing element of the electric hygrometer after which the air entered a distribution manifold. From the manifold it was distributed through the aluminum ducting to the four canisters, and also through a by-pass bleed line, to ambient. The air passing through the canisters was discharged into stainless steel ducting from the outlet of each canister (or its stainless steel transition piece) and was vented to ambient. Orifice meters were located in the ducts between the distribution manifold and each canister.

## Pressure Instrumentation

Pressure taps connected to water manometers were used throughout the system. Monitored points included:

P, - upstream of hygrometer sensing element

Po - upstream of orifice, canister No. 1

 $\Delta P_3$  - differential of orifice, canister No. 1

 $P_{10}$  - upstream of canister No. 1

 $\Delta P_{11}$  - differential of canister No. 1

 $P_{\mu}$  - upstream of orifice, canister No. 2

 $\Delta P_5$  - differential of orifice, canister No. 2

P<sub>12</sub> - upstream of canister No. 2

 $\Delta P_{13}$  - differential of canister No. 2

P<sub>6</sub> - upstream of orifice, canister No. 3

 $\Delta P_7$  - differential of orifice, canister No. 3

 $P_{1h}$  - upstream of canister No. 3

 $\Delta P_{15}$  - differential of canister No. 3

P<sub>8</sub> - upstream of orifice, canister No. 4

 $\Delta P_{Q}$  - differential of orifice, canister No. 4

P<sub>16</sub> - upstream of canister No. 4

 $\Delta P_{17}$  - differential of canister No. 4

# Temperature Instrumentation

Iron-Constantan thermocouples connected to a Brown-Electronik recorder were used to indicate wet bulb and dry bulb temperatures throughout the system. Temperature rise across the canisters was monitored by a Honeywell Brown Electronik Potentiometer Pyrometer. Monitored points included:

T, - upstream dry bulb of canister No. 1

To - downstream dry bulb of canister No. 1

T3 - downstream wet bulb of canister No. 1

 $\Delta T_1$  - temperature rise across canister No. 1

 $T_{li}$  - upstream dry bulb of canister No. 2

 $T_5$  - downstream dry bulb of canister No. 2

 $T_6$  - downstream wet bulb of canister No. 2

 $\Delta T_2$  - temperature rise across canister No. 2

 $T_{7}$  - upstream dry bulb of canister No. 3

 $T_{\Omega}$  - downstream dry bulb of canister No. 3

 $T_{\Omega}$  - downstream wet bulb of canister No. 3

 $\Delta \, \mathrm{T_3}$  - temperature rise across canister No. 3

T<sub>10</sub> - upstream dry bulb of canister No. 4

 $T_{33}$  - downstream dry bulb of canister No. 4

 $T_{20}$  - downstream wet bulb of canister No. 4

 $\Delta T_h$  - temperature rise across canister No. 4

T<sub>12</sub> - dry bulb downstream of electric hygrometer

 $T_{\gamma,li}$  - coolant temperature at 3-way control valve

T<sub>15</sub> - by-pass dry bulb

T<sub>16</sub> - by-pass wet bulb

# Water Analysis

The water content of the inlet air was monitored with an electric hygrometer, dry bulb and wet bulb thermocouples in the by-pass line, and a Hygrophil wet bulb-dry bulb "gun". The water content of the outlet air was monitored by dry bulb and wet bulb thermocouples located in the exit ducts, and by the Hygrophil "gun".

# Oxygen and Carbon Dioxide Analysis

Through a series of sampling valves and lines, the inlet air, outlet air, and standardizing bottled gases could be directed through Drierite rubes, Fisher & Porter Rotometers, and the sensing elements of the gas analyzers. A Beckman Model E2 oxygen analyzer, and a Beckman Model 15A infrared carbon dioxide analyzer were used for the gas analysis.

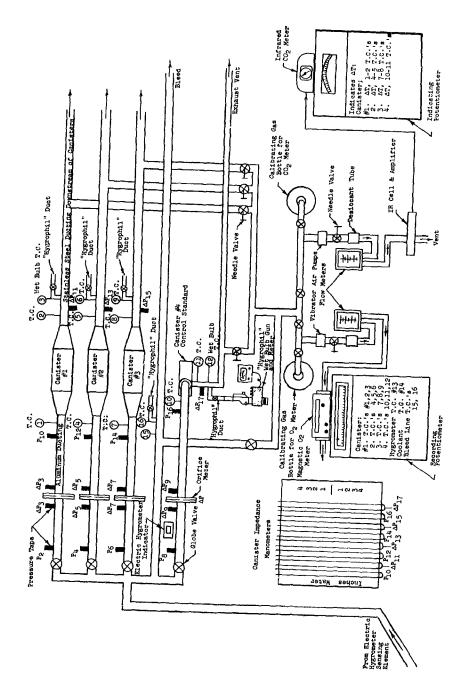


Figure 8 (cont.) Schematic of Open Duct  $KO_2$  Canister Test Facility

## Table 2 Test Instrumentation

Recorder:

Honeywell Brown Electronik Model 153X89-CS-II-III-16A4 S/N 834684

Syn 05-004 Range - 100 to +1100°F Chart No. 6608 FLS 4795 NAA N-241-525

 $\Delta T$  Indicator:

Honeywell Brown Electronik Potentiometer Pyrometer Model  $156 \times 15$ V

S/N 725352 Range 0 to 5mv ELS 2990 TSC 16453

Oxygen Analyzer: Beckman

Model E2

Ranges 0-3, 0-6,0-30% ELS 4829

NAA S-252-998

CO<sub>2</sub> Analyzer:

Beckman Infrared Model 15A ELS 4209 TSC 34006

Hygrometer:

American Instrument Company "Electric Hygrometer" Cat. No. 4-5171 S/N 1075 ELS 3782 NAA 219261

Gun:

Atkins Technical Inc. "A+ Hygrophil Gun" Model 4450 Range 14 to 176° F S/N A-0934 ELS 4848

Sling:

Bacharach Ind. Inst. Co.

Sling Psychrometer

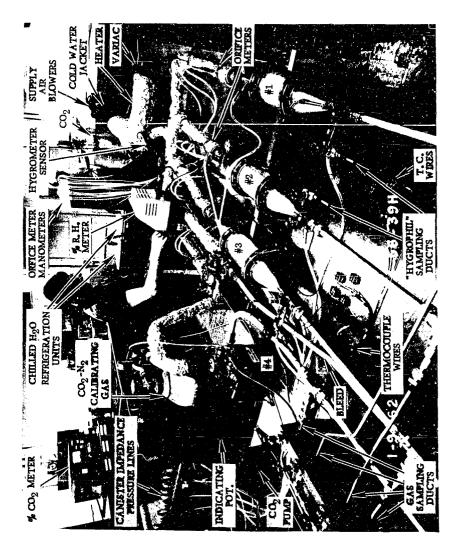


Figure 9 General View of Open Duct  $\mathrm{KO}_2$  Canister Test Facility

Figure 10 Supply Air Blowers

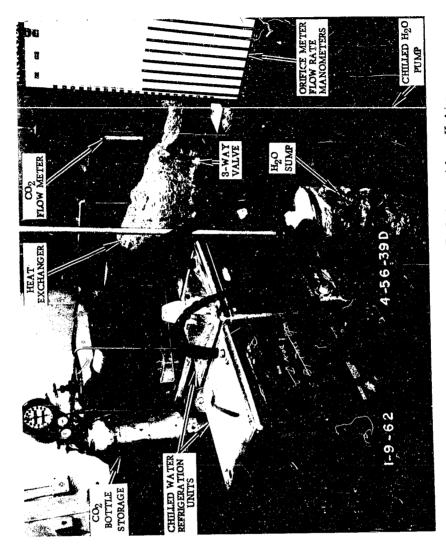


Figure 11 Chilled Water Refrigeration Units

Figure 12 Ducting to Canisters

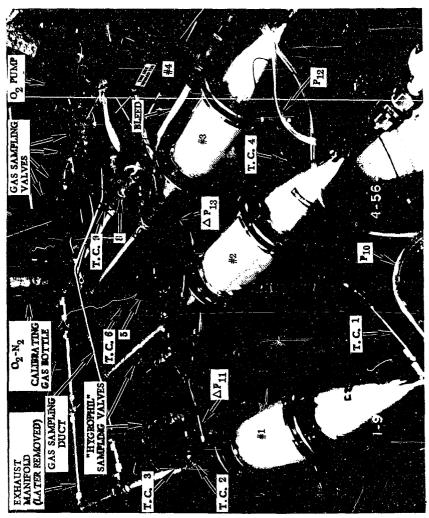
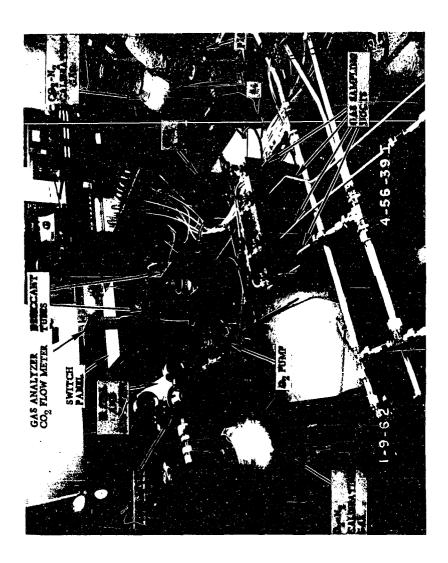


Figure 13 Test Canisters and Control Standard Canister In Test Facility



Taking a Wet Bulb Reading with the Hygrophil "Gun" Figure 14



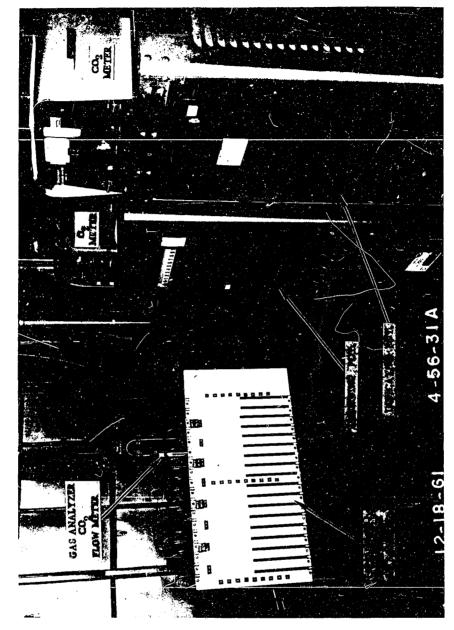


Figure 16 Gas Analyzers and Potentiometers

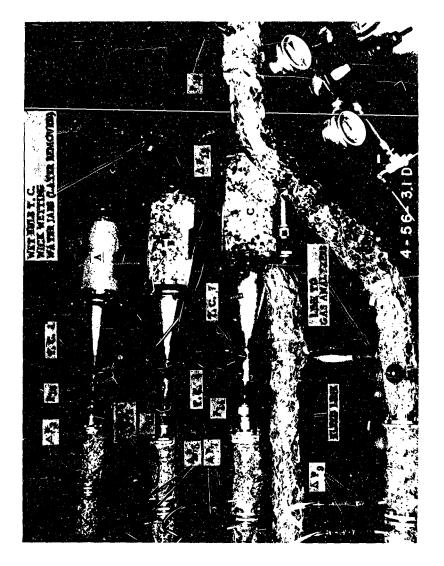


Figure 17 Canisters A, B, and C (Test P-1) in Test Facility



Canisters E, A-8, and F (Test P-10) In Test Facility Figure 18

#### Section VII

#### MATERIALS AND EQUIPMENT TESTED

## CANISTER CONFIGURATIONS

Canisters of various packing and dimensional configurations were tested. Table 3 describes these canisters.

(See Appendix A for derivation of canister sizes, gas flow rates, chemical weights, etc.). Figure 19 shows the Lucite canister shells in modular sizes. Figure 20 shows a typical 10" module with the "A" configuration, while Figure 21 is a sectional view of the same canister with its transitional ducts. Figures 22 and 23 are exploded views of the "A" configuration, and Figures 24 and 25 are view of A-8 canister which consisted of two 10" "A" modules joined in tandem. Figure 26 is typical of the 5" module in either the E or C-2 configuration. Figure 27 is a view of the 10" module in D configuration (four slotted inlet brass probes, and one slotted outlet brass probe). The annular screen is present in all "A" configurations, activated alumina is present in the single "B" configuration, and LiOH is present in all "C" configurations. LiOH is also present in other configurations, but in those cases it is not the dominant design consideration. "D" is a single configuration utilizing the slotted brass probes in conjunction with LiOH. Corrosion is seen on the brass probes. The plugging of "E" and "F" configurations was extremely important in the final evaluation of the role of the annular screen in the successful A and A-8 configurations.

The contents of each canister were packed by a manual loading operation in the proportions indicated in Table 3. The catalyst, copper and manganese oxides, had been impregnated in a portion of the  $\mathrm{KO}_2$  granules during their manufacture by the Mine Safety Appliances Company. The purpose of the catalyst is apparently to decompose  $\mathrm{H}_2\mathrm{O}_2$  formed, and thus increase the rate of  $\mathrm{O}_2$  production from the  $\mathrm{KO}_2$ .

The yellow KO<sub>2</sub> granules were of 2 to 4 mesh size with a density of ~40 lbs/ft³. The black lava rock was a porous volcanic cinder of about 2 to 4 mesh, and had a density of ~51 lbs/ft³. The purpose of the lava rock was to act as a porous inert mechanical separator whose action (it was hoped) would slow down the rate of increase of canister impedance during operation of a canister. This was not too successful as can be seen in the  $\Delta P$  columns (for the "C" canister configuration) of the partially reduced data sheets for tests P-1, P-2, P-3, P-4, and its use was abandoned.

The activated alumina (density =  $60.6~\mathrm{lbs/ft^3}$ ) was chosen for test purposes as a hydrophilic adsorber on the assumption that its use would prevent an excessive collection of either water-of-reaction or atmospheric water in the canisters with resultant slugging and plug formation. From comparison of the  $\Delta P$  columns of the partially reduced data sheets (P-1, P-2, P-3, and P-4) it was seen that its effectiveness was rather limited, as was LiOH (density =  $35.4~\mathrm{lbs/ft^3}$ ) when used for the same purpose. LiOH had the additional advantage of being a  $\mathrm{CO_2}$  adsorber, but it too was unsatisfactory (as a single device) for controlling slugging in the KO<sub>2</sub> bed.

The molecular sieve (Linde X13, density =  $43.0 \text{ lbs/ft}^3$ ) was tried for the same purpose of water control in the canister bed, but its true value as a single controlling device was obscured by the fact that the annular screen was also present in every configuration that contained the molecular sieve. However, since the annular screen by itself (configuration "A") was so successful in controlling plugging, the elimination of the molecular sieve during the development program was substantiated on a least-weight optimization basis alone. Also, the use of molecular sieve was contra-indicated by the fact that the adsorption of water by the molecular sieve inhibited the evolution of  $\rm O_2$  from the  $\rm KO_2$ . A short study was made of other hydrophilic adsorbers, and some promise was indicated for diatomaceous earth, silica gel, Perlite (expanded silica), vermiculite (expanded mica), Drierite (calcium sulfate), chopped blotter-paper, and chopped egg-carton paper. The same consideration for least-weight optimization, as applied to the molecular sieve, might also rule out the above materials.

Although a description of all canister configurations can be found in Table 3, the following paragraphs describe the "A" canisters in detail because the "A" configuration was the most successful design tested, and also the most promising for practical application.

# STRUCTURE OF THE "A" CANISTERS

The canister consisted of a 5" inside diameter Lucite cylinder of 10" (and also 20") overall length. The cylinder was made with flanges at both ends to fit in the open duct system, and between transition pieces with 15° half-angles. Inside the Lucite cylinder there was positioned (by means of a flange on the outer screen, and grooves in the cylinder flanges) a nested cylindrical double-screen insert. The inner screen formed the core. Between the core and the outer screen, the granular chemical was packed. The inlet

end of the double-screen was closed by means of a plate which had four annular slots (with a total cross-sectional free area equal to the 1" duct leading to the transition piece) and which allowed the air to pass into the annulus formed between the Lucite shell and the outer screen. The outlet end of the double screen was closed by a circular screen mounted across the spider and flange which positioned it within the shell. This end-piece screen was welded to the cylindrical double screen.

# OPERATION OF THE "A" CANISTER

The air flowed through the annular slots, passed axially down the annulus between shell and outer screen, flowed transversely and radially across the packed chemical bed into the core, and thence to the outlet. The air could also flow axially through the packed chemical bed directly to the outlet. Its construction provided for a very large area of the chemical bed being exposed to a very low velocity flow of the airstream, instead of the usual small cross-sectional face area presented to the air path in straight-through canisters.

The double canister (see Figure 24) was an example of how the modular feature of the canister can be exploited. Two 10" long canisters were joined in tandem. The inlet plate was not used in the second 10" canister, but was replaced by a circular screen with a rubber plug to force annular flow in the second shell, (see Figure 25).

# COMPOSITION OF THE "A" CANISTER MATERIALS

The outside shell of the canister was made of Lucite for viewing the chemical bed during the tests, but can be made of any non-corrosive material. The mesh screen and inlet plate was used for certain test purposes, but the plate may be solid in the center with no opening for a plug.

# CALCULATED WEIGHTS OF KO2

The weight differences of the  $\mathrm{KO}_2$  in various canisters, as given below, must be taken into account for any final comparison of the effects of operating parameters upon the  $\mathrm{O}_2\text{-}\mathrm{CO}_2$  generating and adsorbing rate characteristics of  $\mathrm{KO}_2$  canisters.

CONFIGURATION	KO <sub>2</sub> lbs
A-3	1.05
В, С	1.47
C-2	1.47
MSA	2.00
D, A-4, A-5, A-6, A-7	2.11
E	2.20
A, A-2	3.16
A-8	6.32
F	8.80

These calculations were made on the basis that a 5" x 10" canister filled entirely with  $\mathrm{KO}_2$  (with or without catalyst) contained 4.4 lbs of  $\mathrm{KO}_2$ . (See Appendix A, paragraph 11). Also, the assumption was made that 30% of the volume available for  $\mathrm{KO}_2$  was taken up by the annular screen device, and the same for the one canister with the brass probes. The bulk density of the  $\mathrm{KO}_2$  was taken as 41 lbs/ft3. The weight of  $\mathrm{KO}_2$  in the MSA canister was measured, not calculated.

Table 3 Canister Configurations

Canister	Description
A.	KO <sub>2</sub> + Catalyst + Annular Screen
в.	$1/3 \text{ KO}_2$ + Cat. + $1/3$ Lava Rock + $1/3$ Act. Alumina
c.	$1/3 \text{ KO}_2 + \text{Cat.} + 1/3 \text{ Lava Rock} + 1/3 \text{ LiOH}$
MSA.	Mine Safety Appliance Standard "Chemox" Canister; KO2 + Cat. + Flat Screens
A-2.	KO <sub>2</sub> (No Cat.) + Annular Screen
A-3.	1/3 KO <sub>2</sub> + Cat. + 1/3 Lava Rock + 1/3 Molecular Sieve + Ann. Screen
D.	2/3 KO <sub>2</sub> + Cat. + 1/3 LiOH + Brass Probes
A-4.	2/3 KO <sub>2</sub> (No Cat.) + 1/3 M.S. + Ann. Screen
A-5.	2/3 KO <sub>2</sub> (No Cat.) + 1/3 LiOH + Ann. Screen
A-6.	2/3 KO <sub>2</sub> + Cat. + 1/3 M.S. + Ann. Screen
A-7.	2/3 KO <sub>2</sub> + Cat. + 1/3 L10H + Ann. Screen
E.	KO <sub>2</sub> + Cat., in 5" long canister (no screen)
A-8.	$KO_2$ + Cat., in two $10^{4}$ long canisters in tandem, with screens, and slotted plate at inlet only.
F.	KO <sub>2</sub> + Cat., in 20" long canister (no screen)
C-2.	2/3 KO <sub>2</sub> + Cat. + 1/3 LiOH, in 5" canister (no screen)

NOTE: All canisters 10" long except as noted in E, A-8, F, and C-2.

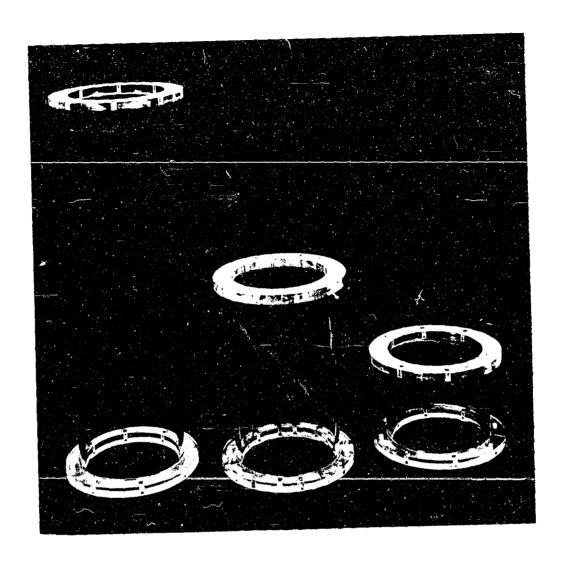


Figure 19 20", 10"and 5" Lucite Canister Shells

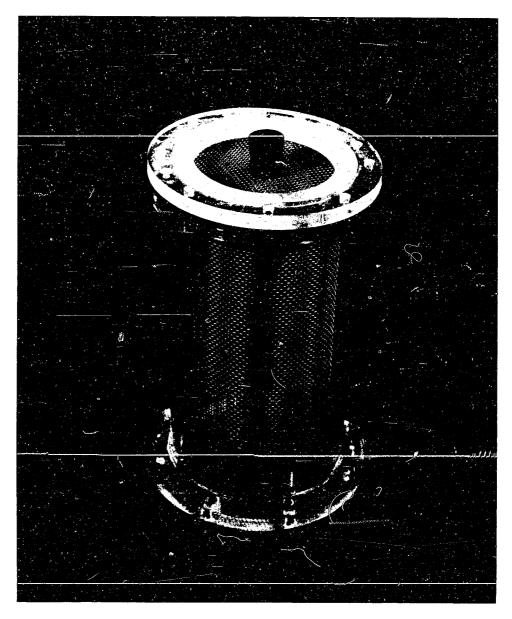
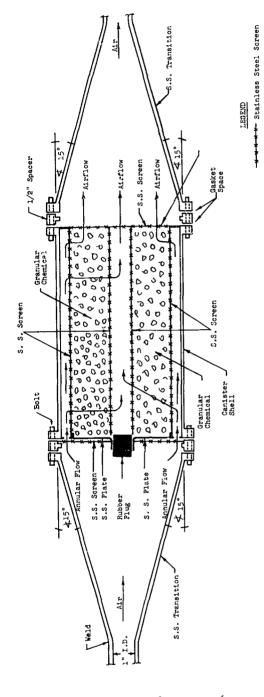


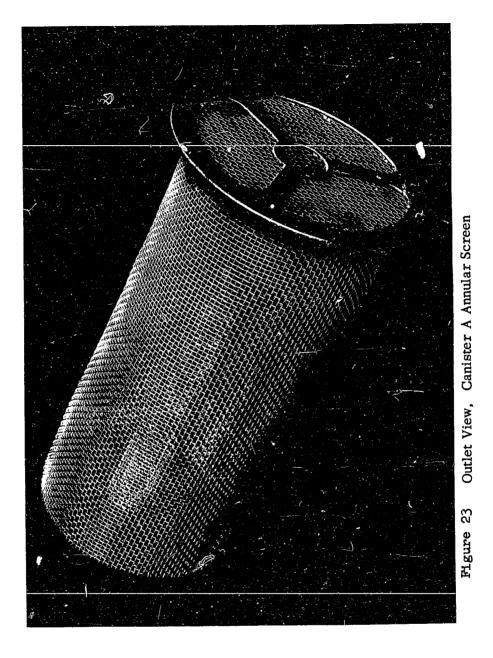
Figure 20 Inlet View, Canister A, Assembled



Sectional View of Canister "A" with Transitional Ducts Figure 21

Exploded View, Canister A

47



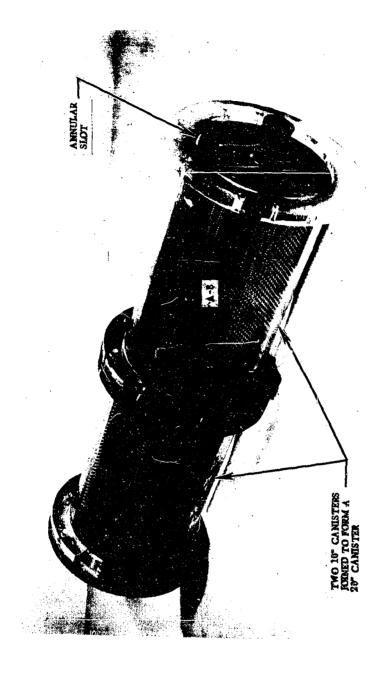


Figure 24 Inlet View, Canister A-8, Partial Assembly

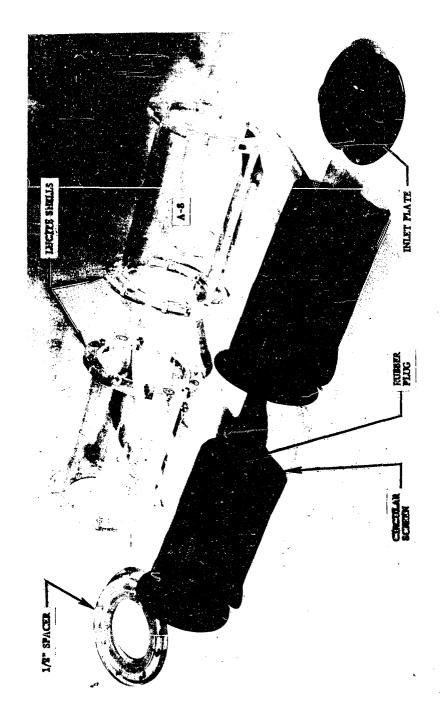


Figure 25 Anister A-8, Exploded View

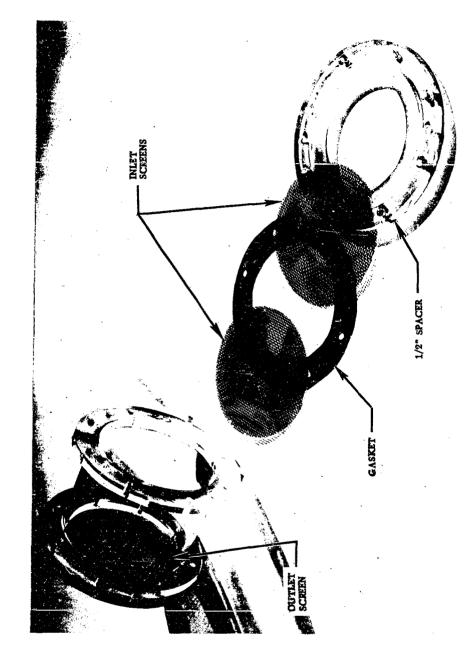


Figure 26 5" Canister, C-2 (and E), Exploded View

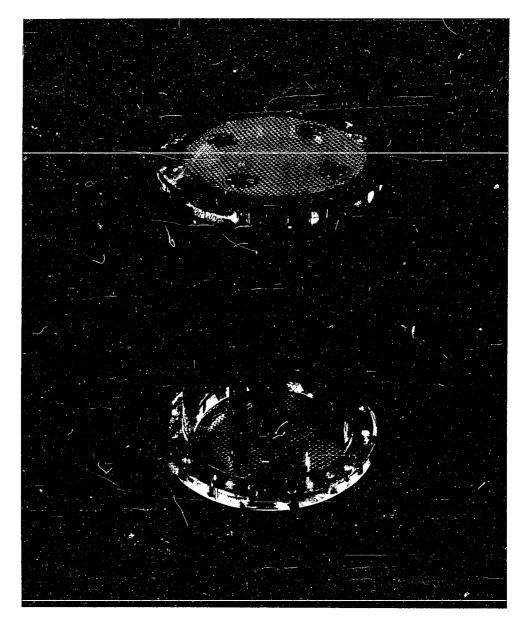


Figure 27 Inlet View, Canister D, Partial Assembly

#### Section VIII

#### TEST PROCEDURES

#### PRE-TEST PROCEDURES

The oxygen analyzer and carbon dioxide analyzer were calibrated against standard gas prior to test runs. Calibration time for standardizing the analyzers was approximately one hour. The battery operated Hygrophil "gun" required recalibration before each reading. Pre-conditioning the system was accomplished by turning on the tap water to the cold water jacket (to prevent heat overload in the system and remove heat created by the blowers), adjusting the refrigeration unit to condition the air passing through the heat exchanger (ethylene glycol solution was used as the cooling agent in the refrigeration system), adjusting the heater to re-heat the air to the required temperature, opening the by-pass line for monitoring the air with the Hygrophil "gun", and adjusting the carbon dioxide flow from the pressurized bottle as required. Three hours were allowed for the pre-conditioning period before the test canisters were valved "on-stream" in the system.

#### TEST RUN

Placing the canisters in operation, i.e., "on-stream", consisted of opening the gate valves at the inlet duct to each canister, monitoring the orifice meter manometer board, adjusting the differential pressures across the orifice to obtain the desired flow rates, and re-adjusting the carbon dioxide flow from the bottle to the desired concentration. Maintaining constant inlet conditions required periodic checking of the refrigeration temperature, re-adjusting the gate valves in the ducts to the canisters as the differential pressure increased across the canisters, and maintaining the desired carbon dioxide concentration in the system.

### DATA RECORDING CYCLE

Twenty-five minutes were required for each recording cycle. The following steps were incorporated in obtaining data for evaluation of the atmospheric composition control systems;

1. Recording inlet conditions of temperature, pressure, relative humidity, and the analysis of oxygen and carbon dioxide concentrations.

- 2. Recording parameters for canisters number 1 and 2: static and differential pressures at each orifice meter; static pressure and drybulb temperature at the inlet to each canister; pressure differential across each canister; drybulb-wet bulb temperatures at the outlet of each canister; analyses of the oxygen production and carbon dioxide adsorption of each canister.
- 3. Recording cycle for canisters number 3 and 4 required the same procedure listed above with the addition of recording the dry bulb-wet bulb temperatures in the by-pass (bleed)line.
- 4. The frequency of the data recording cycles depended on the absolute humidity at each canister inlet.

A high absolute humidity required a data recording cycle each half hour, whereas a low absolute humidity required a cycle only once each hour. These cycles were continued as long as oxygen was generated, or carbon dioxide was adsorbed in the canisters.

#### TEST CONDITIONS

### Series P-1, P-2, P-3, and P-4

Initial test conditions of the first four test series, P-1, P-2, P-3, and P-4, were identical in the following respects:

- 1. The size of the test canisters (not including the MSA control standard) was 5" x 10".
- 2. The face velocity to the test canisters was about 0.55 ft/sec. (It should be noted that the face velocity inside any canister with the annular screen device (series P-1 through P-10) is always much lower than the face velocity to the canister (i.e., to the slotted metal plate across the annular screen). This is because the annular screen area presented to the airflow is about 7 times greater in the 10" long canister, and about 14 times greater in the 20" long canister, than in canisters having axial flow only).

- 3. The (canister) inlet temperatures were about 70-80°F, except series P-3 where the temperatures were about 95-100°F.
- 4. The absolute humidity of the inlet air was in the range of about 45 to 75 grains H<sub>2</sub>O/pound of dry air.

The  $\rm CO_2$  concentration to each canister inlet was about 1.0% for series P-1, P-2, and P-3, and about 0.5% for series P-4.

The main variable under investigation in series P-1, P-2, P-3 and P-4, was the canister bed packing configuration. Canister position #1 had the A configuration, #2 had the B configuration, and #3 had the C configuration. (Table 3).

## Series P-5

Initial test conditions for the canisters in series P-5, were identical in the following respects:

- The size of the test canisters (not including the MSA control standard) was 5" x 10".
- The face velocity to the test canisters was about 0.55 ft/sec.
- The (canister) inlet temperatures were about 80°F.
- 4. The absolute humidity of the inlet air was 60 to 80 grains H<sub>2</sub>0/pound of dry air.
- The CO<sub>2</sub> concentration to each canister inlet was about 0.5%.

The main variable under investigation in series P-5 was the bed packing configuration. Canister position #1 had the A-2 configuration, #2 had the A-3 configuration, and #3 had the D configuration (Table 3).

### Series P-6

Initial test conditions for the canisters in series P-6, were identical in the following respects:

The size of the test canisters (not including the MSA control standard) was 5" x 10".

- 2. The face velocity to the test canisters was about 0.33 ft/sec.
- 3. The (canister) inlet temperatures were about 71-76°F.
- 4. The absolute humidity of the inlet air was about 12 to 28 grains H<sub>2</sub>0/pound of dry air.
- 5. The CO<sub>2</sub> concentration to each canister inlet was about 0.2%.

The main variable under investigation in series P-6 was the bed packing configuration as follows:

#1 , A-2

#2 , A-4

#3 , A-5

### Series P-7

Initial test conditions for the canisters in series P-7, were identical in the following respects:

- 1. The size of the test canisters (not including the MSA control standard) was 5" x 10".
- 2. The face velocity to the test canisters was about 0.33 ft/sec.
- 3. The (canister) inlet temperatures were about 73-79°F.
- 4. The absolute humidity of the inlet air was about 90 to 108 grains H<sub>2</sub>0/pound of dry air.
- 5. The CO<sub>2</sub> concentration to each canister inlet was about 0.2%.

The main variable under investigation in series P-7 was the bed packing configuration as follows:

#1 , A

#2 , A-6

#3 , A-7

### Series P-8

Initial test conditions for the canisters in series P-8, were identical in the following respects:

- 1. The size of the test canisters (not including the MSA control standard) was 5" x 10".
- The face velocity to the test canisters was about 0.33 ft/sec.
- 3. The (canister) inlet temperatures were about 70-80°F.
- 4. The absolute humidity of the inlet air was about 10 to 20 grains H<sub>2</sub>0/pound of dry air. (Near the end of this test the humidity was allowed to rise to about 74 grains H<sub>2</sub>0/pound of dry air).
- 5. The CO<sub>2</sub> concentration to each canister inlet was about 0.5%.

The main variable under investigation in series P-8 was the bed packing configuration as follows:

#1 , A

#2 , A-6

#3 , A-7

# Series P-9

Initial test conditions for the canisters in series P-9 were identical in the following respects:

- 1. The size of the test canisters (not including the MSA control standard) was 5" x 10".
- 2. The face velocity to the test canisters was about 0.55 ft/sec.
- The (canister) inlet temperatures were about 72-80°F.
- 4. The absolute humidity of the inlet air was about 55 to 88 grains H<sub>2</sub>0/pound of dry air.
- 5. The CO<sub>2</sub> concentration to each canister inlet was about 0.2%.

The main variable under investigation in series P-9 was the bed packing configuration as follows:

#1 , A

#2 , A-6

#3 , A-7

# Series P-10

Initial conditions for the canisters in the last, and final, series P-10 were identical in the following respects:

- 1. The face velocity to the test canisters was about 0.33 ft/sec.
- 2. The (canister) inlet temperatures were about 70-80 °F.
- 3. The absolute humidity of the inlet air was about 60 to 87 grains H<sub>2</sub>O/pound of dry air.
- 4. The CO<sub>2</sub> concentration to each canister inlet was about 0.2%.

The two main variables under investigation in series P-10 were the canister size and bed packing configuration as follows:

#1 , E, and C-2 (5" x 5")

#2 , A-8 (5" x 20")

#3 ,  $\mathbf{F}$  (5" x 20")

#### Section IX

#### DISCUSSION

### THE ROLES OF CARBON DIOXIDE, WATER AND CATALYST

The partial pressure of CO2 plays the dominant role in the reactions of the KO2 bed. Its dominance is even greater in establishing the CO, adsorption rate of the KO, bed than in establishing the rate of oxygen evolution. (Figures 155 and 156, and 155 and 160). The oxygen evolution rate may be adequate (for a time) even with a low CO2 partial pressure provided the absolute humidity is sufficiently high (Figure 163). In such cases, however, the CO<sub>2</sub> adsorption rate is inadequate (Figure 164). Since it may be desirable that a sealed cabin provide a low partial pressure of CO2 for respiration, the insufficiency of the KO2 to chemically match the man's R.Q. at a low % CO2, must be in part corrected by an additional CO2 adsorber in a dual canister system. Canister A-8, while remaining unplugged the entire run of 1765 minutes (29 hours and 25 minutes), would not maintain the support level for one man after 375 minutes (6 hours and 15 minutes). The "A" canisters in P-2 and P-3 (Tables 26 and 27) showed a more than sufficient 0, generation rate up to 435 minutes (7 hours and 15 minutes), both of which experienced a high absolute humidity and a high CO2 partial pressure. The PCO<sub>2</sub> in canister A-8 was very low, and probably accounts for this difference. No long duration run was performed with an A-8 configuration using both high absolute humidity and high PCO2. It would have been interesting to see how long the canister would have sustained one man under such conditions. The test run of canister A-8 indicates that a problem still remains in using KO2 to support a man for a long period of time. The effect of the catalyst on the O<sub>2</sub> evolution rate is shown in Table 30, P-6, A-2, A-4, and A-5. In the absence of the catalyst, the O<sub>2</sub> evolution rate drops sharply. Therefore, it is possible that a catalyst (other than that used in this test program), or different operating conditions (such as bed temperature, or total pressure) than those used in these tests may be selected which will allow the man's support level to be reached and maintained by KO<sub>2</sub> over a very long period of time. In the rat test, however, the KO<sub>2</sub> did maintain the physiological support level of the animals for 25 days, which indicates there may be difficulty in obtaining the same R.Q. responses in an open duct system as those experienced in a closed ecological system. Chemical R.Q. determinations apparently are best made in a physiological closed-loop respiratory system.

It is believed that the water and CO2 reacting with the  $KO_2$  forms a  $K_2CO_3$  coating around the  $KO_2$  granule. At a low partial pressure, the  $CO_2$  may not diffuse well through this coating, whereas with a high partial pressure the CO2 can, and thereby the evolution rate of oxygen will be maintained as required for physiological support. Some white-coated (K2CO3) KO2 granules from test run A-8 were broken open and unused yellow KO2 was observed. The more effectively the support level was maintained (in other tests) the more completely white  $(K_2CO_3)$  were the  $KO_2$  granules inside. Figure 28 is a diagramatic representation of this theoretical physicochemical process of a KO2 granule with H2O and CO2, as appears to be indicated by the results of these open duct canister tests. In the first two steps of the diagram,  $(\Delta_1, \Delta_2)$ ,  $H_2O$ initiates the process by first hydrating the  $\rm KO_2$  (evidenced by a color change from yellow to orange), and then with more  $\rm H_2O$ forming an outside layer of white KOH (from this hydrate layer). At the same time, more yellow KO, becomes hydrated in a layer beneath the newly formed white  $K\overline{O}H$ . In step 3,  $(\Delta_3)$ , the outer layer of KOH becomes K2CO3 under the influence of CO2, and the layer beneath changes from hydrate to KOH, etc. In step 4, ( $\Delta\mu$ ), the KOH layer has "moved" beneath two layers of carbonate by virtue of further action of HoO upon orange hydrate  $(KO_2 \cdot H_2O)$  at that depth. The last step, n, represents the ideal end-point if 100% conversion from  $KO_2$  to  $K_2CO_3$  has occurred, and  $\sum_{i=1}^{n} \Delta_i$  represents the summation of all the finite incremental changes between steps 4 and n. Under low PCO2, the KOH layer, as it "moves" deeper in the granule, does not change quickly enough to K2CO3, and since KOH is very hygroscopic, water is not available for further hydration and reaction with the KO2, and consequently the rate of O2 generation falls below the man-support level. Identification of the crystalline compounds at the different layers, and proof of this theory, might be accomplished by X-ray diffraction of micro-slices through the granules, or by the technique of inorganic micro-qualitative analysis.

### THE ROLE OF THE ANNULAR SCREEN

The effectiveness of exposing a large area of the chemical bed to a very low velocity flow airstream (as accomplished by the annular screen in configuration "A") in order to indefinitely prevent plugging of the canister is very well demonstrated in Figure 41, and Table 14. The configurations with the annular screen consistantly showed

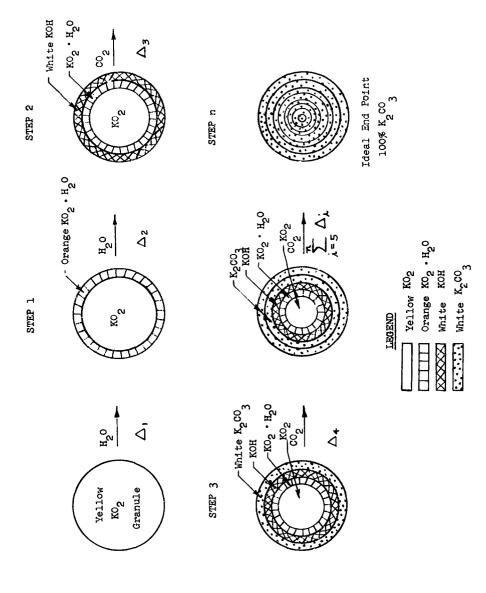


Figure 28 Theoretical Physico-Chemical Frocess of  $m KO_2$  with  $m H_2O$  and  $m CO_2$ 

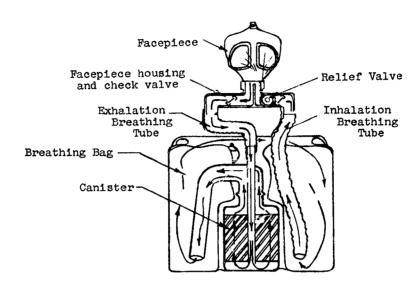
the minimum pressure drop across the canister both at the beginning and at the end of every test run. There was never any important increase in  $\Delta P$  whenever the annular screen was used, i.e., never more than 0.5" H20 increase per linear foot. and usually much less, regardless of the absolute humidity. The increase in resistance to the gas flow is a function of how well the granular state of the chemical is maintained. Water vapor (present in a physiological respiratory circuit) passing through a granular hygroscopic chemical, either directly by physical adsorption or indirectly through chemical adsorptive processes, tends in time to turn the granular chemical either into a soggy mass (Figure 40) which greatly restricts the gas flow, or into a fused solid salt plug (Figures 39 and 41) which completely blocks off all air passage through the canister. This has been the biggest drawback to the use of solid granular hygroscopic chemicals in any chemical process wherein a liquid vapor (such as water vapor) is present in the gas stream. The double canister (Figure 24) is an example of how the modular feature of the annular screen can be exploited. Two 10" long "A" canisters were joined in tandem, in which case the inlet plate was not used in the second 10 canister, but was replaced by a circular screen with a rubber plug to force annular flow in the second shell (Figure 25).

#### ADVANTAGE OF THE ANNULAR SCREEN DEVICE OVER THE CONTROL CANISTER

The control used in all tests was the standard Mine Safety Appliances (MSA) Company "Chemox" canister. This canister is used in the "Chemox Oxygen Breathing Apparatus" (Figure 29) which is a self-contained breathing circuit operating independently of the outside air. The replaceable "Chemox" canister, containing KO<sub>2</sub> with a catalyst, generates (upon contact with moisture in the exhaled breath) a supply of oxygen for breathing requirements for a short period of time, (i.e. about 45 minutes when used in the "Chemox Apparatus"). The evolved oxygen flows upward through the canister and into the breathing bag reservoir where it is cooled, and then passes through the inhalation tube to the wearer's facepiece. The evolvement of the oxygen automatically continues in accordance with breathing rates. An automatic timing device rings a bell at the end of a preset time (45 minutes) to indicate that the wearer should return to fresh air.

In the "Chemox" canister, the air can flow only axially through the packed chemical bed. The consequence of a small cross-sectional face area presented to the air path is the deterioration of the granular nature of the chemical bed, with a resultant increase in resistance to the air-flow through a soggy or fused chemical bed. The NAA annular screen device

## "CHEMOX" APPARATUS ASSEMBLY



"CHEMOX" CANISTER

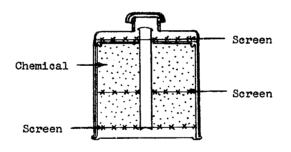


Figure 29 Chemox Oxygen Breathing Apparatus

maintains the chemical bed in the granular condition regardless of the length of time it is operated, even under conditions of high absolute humidity, or low partial pressure of carbon dioxide, in the air stream. The NAA annular screen device should prove very useful in any closed-circuit respiratory system using KO<sub>2</sub> or lithium oxides in the canister. This canister device should also be very useful in any chemical process where a gaseous stream is passed through, or over, a solid granular chemical bed. This may be in a chemical process having nothing at all to do with atmospheric composition control.

#### TEMPERATURE EFFECTS IN GENERAL

It was expected that the inlet temperature would affect the performance of a KO<sub>2</sub> canister in an atmospheric composition control system. In test series P-3 (Table 7) the inlet temperature was raised to about 95-100°F. The favorable effect of this on the O<sub>2</sub> generation of configurations A, B, and C is seen by comparing Figure 149 with Figure 145. At 200 minutes, for example, the average O<sub>2</sub> generation rate is increased about 67%. If configuration A only is compared at 200 minutes (in Figures 145 and 149), the O<sub>2</sub> generation rate is seen to increase about 68%. The annular screen (in configuration A) shows an even greater superiority over C and MSA at the higher temperature in P-3 than at the lower (normal) temperature in P-1. The O<sub>2</sub> generation rate of B is better than either C or MSA at the high temperature, presumably because the water-absorbing capacity of the activated alumina (in B) is negated at the higher temperature.

The effects of increasing temperature on the CO<sub>2</sub> adsorption rate characteristic is seen by comparing Figures 146 and 150. At 200 minutes, configurations A and C appear not to be effected by the increase in temperature in their CO<sub>2</sub> adsorption rates, although the increased temperature did increase their O<sub>2</sub> generation rates. Configuration B has increased its CO<sub>2</sub> adsorption rate about 41% with the increase in temperature, presumably due to negating the influence of the alumina, making more free water available for reaction. The MSA configuration has decreased its CO<sub>2</sub> adsorption rate by about 35% at 200 minutes. This follows along with the slight decrease in O<sub>2</sub> generation rate (of the MSA Configuration) at the higher temperature. Referring back to page 42 we see that B and C have less KO<sub>2</sub> in their configurations than does MSA. A, however, has more than MSA. This order holds true for the O<sub>2</sub> generation rates in Figure 145, but (unaccountably) in Figure 149, this order does not hold true for O<sub>2</sub> generation rates at the higher temperature.

As regards CO<sub>2</sub> adsorption at either the lower (Figure 146) or the higher (Figure 150) temperature it is seen that although C is of a less weight (KO<sub>2</sub>) than A or MSA, it has the highest CO<sub>2</sub> adsorption rate, presumably due to the affinity of the LiOH for CO<sub>2</sub>. Configuration B, although containing a lesser weight of KO<sub>2</sub> than MSA, shows a higher (Figure 150) CO<sub>2</sub> adsorption rate than MSA (at the high temperature). This may be due to additional water being freed from the alumina by the increasing temperature, and becoming available for reaction with KO<sub>2</sub>.

The effects of temperature could not be pursued more fully during this test program, but the results of series P-3 clearly show that in any of the other test series where the 02 generation rates failed to reach the man support level, an increase in the inlet temperature (or KO2 bed temperature) would have increased the Oo generation to the support level required, provided, of course, that the other conditions were satisfactory. These other conditions are parameters such as packing configuration, absolute humidity, partial pressure of CO2, airflow rate, face velocity, etc. It is interesting to note (on page 67) the effect of a higher bed temperature on the chemical utilization (O2 generation) index. In series P-2, this index for canister B is ~ 112%, while in P-1 the index for B is only ~ 40%. The major operating parameter difference between the two was a higher bed temperature for B in P-2 than in P-1. Also, in series P-2, the index for canister C is ~ 117%, while in P-1 the index for C is only ~ 53.5%. Here again, the only operating parameter difference was a higher bed temperature (evidenced by a higher  $\Delta T$ ) in P-2 than in P-1.

The effects of absolute humidity,  $PCO_2$ , and temperature on  $O_2$  generation rates, are graphically shown on pages 68, 69, and 70. The effects of absolute humidity,  $PCO_2$ , and temperature on  $CO_2$  adsorption rates, are graphically shown on pages 71, 72, and 73.

#### FLOWRATE EFFECTS IN GENERAL

Some comparisons may be made concerning the effect of changing velocities in identical canisters. Comparing the canisters in series P-7 and P-9 (both of which series have similar absolute humidity, % CO<sub>2</sub>, and temperature conditions) by inspecting Tables 31 and 33, we note that in the first fifteen minutes of operation, canister configuration A experienced a 150% increase in CO<sub>2</sub> adsorption at the increased airflow rate (P-9), and A-6 experienced about 67% increase in CO<sub>2</sub> adsorption. Similar increases are seen for A-7 and the MSA canister. The O<sub>2</sub> generation rates are not increased as consistently in P-9 (higher velocity) compared to P-7

(lower velocity), but they do show some increase (about 20-25%) in the A and A-7 configurations. Face velocity to the canister will be a less important factor in a closed-circuit respiratory system than in an open duct system. The closed-circuit offers repeated passes of the same atmosphere through the canister, whereas the open duct system used in this test program is a "one-shot" process.

## RELATIVE CALCULATED CHEMICAL UTILIZATION INDEX

Table 4 shows the relative calculated chemical utilization index (%) of each canister configuration (for the best set of operating parameters at which it was tested), based on the weight of KO2 in the canister. Because of economic (and time) considerations, emphasis was placed on testing in such a manner as to most rapidly develop a near-optimum canister design, and to discover as quickly as possible the broad areas of near-optimum operating parameters. Because of these facts it would be difficult to say that any particular canister configuration could not have had a higher relative index number, given more favorable operating conditions. Table 4 must be interpreted with an eye to the "varying operating parameters", which are not shown on the table. Time is one example of a variable operating parameter which must be considered when interpreting Table 4. The "best" index was determined irrespective of time. The "best" indexes for A-8 and MSA, for example, were determined at 2000 minutes (extra-polated past 1765 minutes) and 1120 minutes respectively. Time for other "best" indexes varied from 317 to 575 minutes. Another set of utilization index numbers would be obtained if, for example, all canisters were evaluated for a given common total operating time. The chemical utilization index numbers are based on total operating time (with each canister operating under its own particular conditions), which is not always the same as total <u>life-support</u> time for a man. The cross-over time-point can be seen by comparing the "O2 consumed" and "O2 generated" columns in Tables 25 through 34. Provided the other operating parameters are satisfactory, an increase in either the temperature, the airflow rate, or the CO2 concentration, or a decrease in the superficial velocity, may be all that is required to increase the O2 generation rate of those canisters whose rates are marginal, to the life-support level required by a man. All such variables should be considered when interpreting the relative index values of Table 4. The effects of some of these variables are graphically shown on pages 68, 69, 70, 71, 72, and 73.

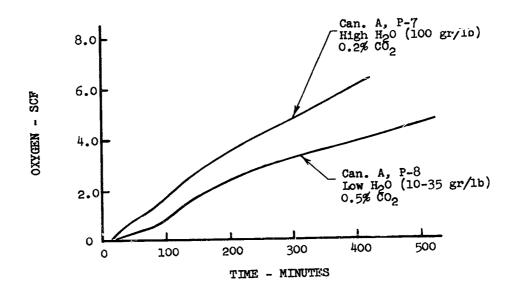
Table 4 Relative Calculated Chemical Utilization Index

CANISTER	CALCULATED WEIGHT OF KO <sub>2</sub> IN CANISTER, 1bs KO <sub>2</sub>	THEORETICAL AVAILABLE O2 (33.7%), lbs O 2	THEORETICAL AVAILABLE 0 <sub>2</sub> (at 1 atmos. and 70°F, 0.0891bs/ 1.077 ft3), SCF 0 <sub>2</sub>	BEST CALCULATED UTILIZATION INDEX, where, Actual SCF 02 Theoretical SCF 02	TOTAL TIME OF CANISTER OPERATION, MINUTES
A~3	1.05	0.354	4.28	2.38 <b>-</b> 55%	317 (P-5)
В	1.47	0.495	5 <b>.</b> 98	6.73 5.98 = 112% *	435 ( <b>P-2</b> )
c	1.47	0.495	5.98	7.00 5.98 = 117% *	440 (P-2)
C-2	1.47	0.495	5.98	4.76 5.98 <b>-</b> 79.7%	365 (P-10)
MSA	5.00	0.674	8.15	<u>5.77</u> ≈ 71% 8.15 ≈ 71%	1120 (P-10)
D	2.11	0.711	8.60	6.45 8.60 = 75%	322 (P-5)
A-4	2.11	0.711	8.60	0.99 8.60 = 11.5%	446 (P-6)
A-5	2.11	0.711	8.60	1.91 8.60 = 22.2%	456 (P-6)
A-6	2.11	0.711	8.60	<u>5.33</u> 8.60 = 62%	565 (P-9)
A-7	5•11	0.711	8.60	5.92 8.60 = 69%	575 (P-9)
R	5.50	0.741	8.96	4.45 8.96 <b>-</b> 49.6%	375 (then plugged) (P-10)
A	3.16	1.065	12.90	9.03 12.90 = 70%	435 (P-3)
A-2	3.16	1.065	12.90	5.79 12.90 <b>4</b> 5%	317 (P-5)
A-8	6.32	2.129		18.50 <b>-</b> 72%	Extrapolated to 2000 (P-10)
F	8.80	2.965	35.80	6.19 35.80 = 17%	380 (then plugged) (P-10)

<sup>\*</sup> NOTE: An index greater than 100% indicates that the measured weight of KO in these particular canisters was probably greater than the calculated weight.

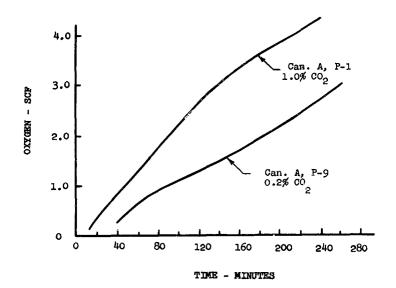
ABSOLUTE HUMIDITY,  $\text{PCO}_2$ , AND TEMPERATURE EFFECTS ON  $\text{O}_2$  GENERATION RATES

Some interesting effects on oxygen generation rates are seen in the following tables and figures. Normal rates are seen in Table 18, while the rate depressing effect of molecular sieve is seen in Table 19 (A-3), Table 20(A-4), and Table 21 (A-6). The rate depressing effect of low absolute humidity is seen in Table 22 (P-8) as compared to Table 20 (A-2, A-4, and A-5). The plugging of configurations E and F are clearly shown in Figures 130 and 136. The man-rate support level graphs (Figures 145 thru 164) are very clear comparative representations of all the factors at various levels. We can compare 02 generation curves from the man-rate support level graphs for the effect of absolute humidity on a particular canister configuration as follows;



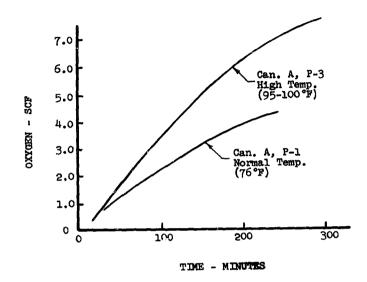
Although the PCO<sub>2</sub> in P-7 is seen to be lower than in P-8, the higher absolute humidity in P-7 was responsible for increasing the O<sub>2</sub> generation for canister A, P-7 over that of A, P-8. Canisters A, P-7, and A, P-8 operated under similar conditions of temperature and flowrate. The absolute humidity for the P-7 series was about 100 grains/lb of dry air, and about 10-35 grains/lb of dry air for the P-8 series. At 420 minutes the total O<sub>2</sub> generation for the A Canister in P-7 series was ~ 31% greater than that of the A canister in P-8 series.

The graph below compares the O<sub>2</sub> generation curves from the man-rate support level graphs for the effect of PCO<sub>2</sub> on a particular canister configuration.



Canisters A, P-1, and A, P-9 operated under similar conditions of absolute humidity, temperature, and flowrate. The PCO2 in the P-1 series was  $\sim 1.0\%$  and  $\sim 0.2\%$  in the P-9 series. At 210 minutes the total O2 generation for the A canister in the P-1 series (1.0% CO2) was  $\sim 66\%$  greater than that of the A canister in the P-9 series (0.2% CO2).

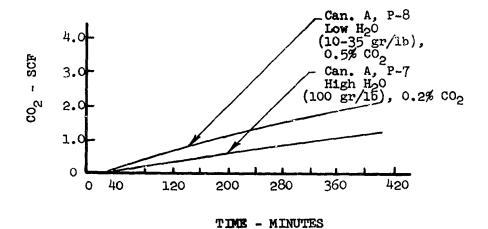
The graph below compares the  $\rm O_2$  generation curves from the man-rate support level graphs for the effect of temperature on a particular canister configuration.



Canisters A, P-3 and A, P-1 operated under similar conditions of absolute humidity, PCO<sub>2</sub>, and flowrate. The temperature was  $\sim 95 - 100\,^{\circ}\text{F}$  in the P-3 series, and  $\sim 76\,^{\circ}\text{F}$  in the P-1 series. At 250 minutes the total O<sub>2</sub> generation for the A canister in the P-3 series was  $\sim 63.5\%$  greater than that of the A canister in the P-1 series.

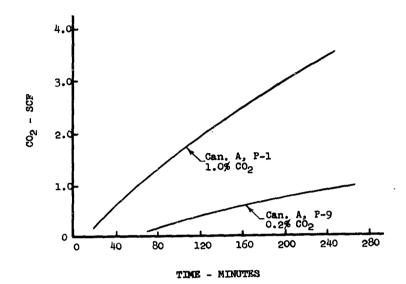
ABSOLUTE HUMIDITY,  $PCO_2$ , AND TEMPERATURE EFFECTS ON  $CO_2$  ADSORPTION RATES

We can compare the  $\rm CO_2$  adsorption curves of the same canisters whose man-rate support level graphs were useful for comparing various effects on the  $\rm O_2$  generation rates. The graph below compares the  $\rm CO_2$  adsorption curves from the man-rate support level graphs for the effect of both absolute humidity and  $\rm PCO_2$  on a particular canister configuration.



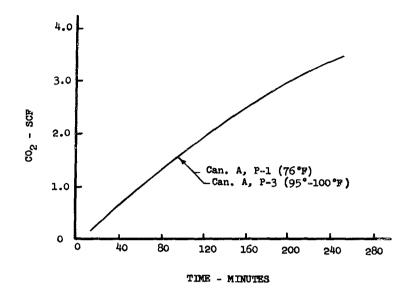
Canisters A, P-7, and A, P-8 operated under similar conditions of temperature and flowrate. The absolute humidity for the P-7 series was about 100 grains/lb of dry air, and only about 10-35 grains/lb of dry air for the P-8 series. At 420 minutes, however, the total  $\rm CO_2$  adsorption for the A canister in the P-8 series was  $\sim 73\%$  greater than that of the A canister in the P-7 series. Since the absolute humidity was much lower in P-8 than in P-7, this difference is clearly due to the greater  $\rm PCO_2$  (0.5%) in P-8 than in P-7 (0.2%), and shows the dominance of  $\rm PCO_2$  as a controlling variable in  $\rm CO_2$  adsorption (of the A canister) over that of absolute humidity.

The graph below compares the  $\rm CO_2$  adsorption curves from the man-rate support level graphs for the effect of  $\rm PCO_2$  on a particular canister configuration.



Canister A, P-1 and A, P-9 operated under similar conditions of absolute humidity, temperature, and flowrate. The PCO<sub>2</sub> in the P-1 series was  $\sim 1.0\%$  and  $\sim 0.2\%$  in the P-9 series. At 210 minutes the total CO<sub>2</sub> adsorption for the A canister in the P-1 series (1.0% CO<sub>2</sub>) was  $\sim 447\%$  (4-1/2 times) greater than that of the A canister in the P-9 series (0.2%) CO<sub>2</sub>, showing the very large effect of PCO<sub>2</sub> on CO<sub>2</sub> adsorption rates, and the similarity to the five fold increase in PCO<sub>2</sub>.

The graph below compares the  $\rm CO_2$  adsorption curves from the man-rate support level graphs for the effect of temperature on a particular canister configuration.



It is immediately seen that the two  $CO_2$  adsorption curves above coincide in this case. Canisters A, P-1 and A, P-3 operated under similar conditions of absolute humidity,  $PCO_2$ , and flowrate. Even though the temperature in the P-3 series was  $\sim 95-100\,^{\circ}\text{F}$ , and  $\sim 76\,^{\circ}\text{F}$  in the P-1 series, the  $CO_2$  adsorption rate (for the A canister) was not affected by the increased temperature.

#### THE DUAL CANISTER SYSTEM

The author feels strongly that the dual-canister system will prove superior to a single canister containing one solid chemical (or even a mixture of two solid chemicals) for the purpose of O<sub>2</sub> generation and CO<sub>2</sub> adsorption. The best way to match man's R.Q. with a solid chemical system is to control the airflow through each one of two canisters (in the dual-canister system) by oxygen and carbon dioxide sensors which will modulate the airflows to each canister upon demand signals.

## ADVANTAGES OF LITHIUM OXIDES OVER KO2

Lithium peroxide,  $\text{Li}_2\text{O}_2$ , appears to offer some promise as an alternate to  $\text{KO}_2$  for solid chemical atmosphere composition control.  $\text{Li}_2\text{O}_2$  has a slight theoretical weight advantage ( $\sim 34.6\%$  available  $\text{O}_2$ ) over  $\text{KO}_2$  ( $\sim 33.7\%$  available  $\text{O}_2$ ). Of more importance than the slight weight advantage is the advantage of  $\text{Li}_2\text{O}_2$  in providing a less hygroscopic intermediate product (LiOH instead of KOH) which is less subject to liquid slugging. No published data is currently available on the performance of this promising material for use in an atmospheric composition control system, although lithium peroxide of 90 to 95 percent purity and good stability is available from three known sources. (Reference 3).

If lithium superoxide,  $\text{LiO}_2$ , can be synthesized and isolated in a stable form at normal room temperature (Reference 3), it would have a greater theoretical weight advantage ( $\sim$  61.7% available  $\text{O}_2$ ) over  $\text{KO}_2$  ( $\sim$  33.7% available  $\text{O}_2$ ).

Use of a stable form of sodium superoxide, NaO<sub>2</sub>, ( $\sim$  37% available O<sub>2</sub>), might also prove advantageous weightwise, especially with the NAA annular screen device used to prevent plugging of the canister.

# DISPROVING A THEORETICAL EQUATION FOR MATCHING MAN'S R.Q.

During the test program, it was observed that the necessary rate of reaction in the KO<sub>2</sub> canisters required a considerably higher absolute humidity (grains H<sub>2</sub>O per pound dry air) than had been calculated on the basis of the much-publicized theoretical equation for KO<sub>2</sub> purporting to match man's R.Q. (used in Appendix A). This lead to the conclusion that the theoretical equation does not occur. The theoretical reaction is written as follows:

$$2 \text{ KO}_2 + 1.23 \text{ CO}_2 + 0.23 \text{ H}_2\text{O} = 0.77 \text{ K}_2\text{CO}_3 + 0.46 \text{ KHCO}_3 + 1.5 \text{ O}_2$$

The matching of man's R.Q., assuming it does occur in this equation, is calculated as follows:

$$\frac{1.23 \text{ pound-mo?. } \text{CO}_2}{1.5 \text{ pound-mol } \text{O}_2} = \frac{1.23 \text{ x } \frac{1:4}{1.5 \text{ x } 32}}{1.5 \text{ x } 32} = \frac{54.12 \text{ lbs } \text{CO}_2}{48 \text{ lbs } \text{O}_2}$$

$$\frac{0.123 \text{ lb } \text{CO}_2}{1 \text{ ft}^3} = \frac{54.12 \text{ lbs } \text{CO}_2}{\text{ x } \text{ ft}^3}, \text{ x } = 440 \text{ ft}^3 \text{ CO}_2$$

$$\frac{0.089 \text{ lb } \text{O}_2}{1 \text{ ft}^3} = \frac{48 \text{ lbs } \text{O}_2}{\text{ x } \text{ ft}^3}, \text{ x } = 540 \text{ ft}^3 \text{ O}_2$$

$$R. Q. = \frac{\text{Vol } \text{CO}_2}{\text{Vol } \text{O}_2} = \frac{440 \text{ ft}^3}{540 \text{ ft}^3} = 0.82$$

The calculation of the theoretical weight of H<sub>2</sub>O for a pound of dry air (absolute humidity) required for this reaction to occur stoichiometrically is as follows:

Assuming 2.25 lb. CO<sub>2</sub> is produced per man-day, then the equation is simplified to the proportion,

$$\frac{\text{1.23 pound-mol CO}_2}{\text{2.25 lb CO}_2} = \frac{\text{0.23 pound-mol H}_2\text{0}}{\text{X lb H}_2\text{0}}$$

or

$$\frac{1.23 \times (44)}{2.25} = \frac{0.23 \times (18)}{X}$$

$$X = 0.173 \frac{1bs. H_20}{man-day}$$

and

$$\frac{0.173}{(24)(60)} = 0.00012 \frac{1b \text{ H}_20}{\text{man-min.}} \times 7000 = 0.84 \frac{\text{grains H}_20}{\text{man-min.}}$$

Assuming an airflow of 12.5  $\frac{\text{lbs}}{\text{man-hour}}$ , an inlet air temperature of 70°F (530°R), a gas pressure of 1 atmosphere, and the weight of one pound-mol of air to be 29 pounds, the inlet absolute humidity required per minute

(on a one man basis) to satisfy the theoretical equation is then calculated as follows:

Absolute Humidity = 
$$\frac{0.84 \frac{\text{grains}}{\text{min}}}{2.78 \frac{\text{ft}^3}{\text{min}}} = 0.302 \frac{\text{grains H}_20}{\text{ft}^3 \text{ air}}$$

or Absolute Humidity = 
$$\frac{0.84 \frac{\text{grains}}{\text{min}}}{12.5 \frac{\text{lbs}}{\text{hr}}} \times 60 \frac{\text{min}}{\text{hr}} = 4.04 \frac{\text{grains H}_20}{\text{lb air}}$$

Since no apparent chemical reaction occurred during the tests at an absolute humidity as low as 8 grains  $\rm H_2O$  per pound of dry air, this was understood (by empirical induction) to mean that the attempt to prove the theoretical equation correct had failed, and therefore the equation was not true. This was further substantiated by the fact that the use of  $\rm KO_2$  (without a separate canister of  $\rm CO_2$  adsorbent) could not provide for matching the respiratory quotient of man.

#### Section X

#### REFERENCES AND BIBLIOGRAPHY

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#### APPENDIX A

#### A CANISTER DESIGN METHOD

The following method was used to determine the basic  $KO_2$  canister sizes for the atmospheric composition control system test:

1. Assume the following:

$$O_2$$
 consumption/man-day = 2.0 lbs.  $O_2$ 

2. Assume that both the absolute and relative humidities (through dew point and temperature control respectively), temperature and % CO<sub>2</sub> input to the test canisters can be controlled to match the overall chemical reaction with man's R.Q. (0.82) as follows:

$$2 \text{ KO}_2 + 1.23 \text{ CO}_2 + 0.23 \text{ H}_2\text{O} = 0.77 \text{ K}_2\text{CO}_3 + 0.46 \text{ KHCO}_3 + 1.5 \text{ O}_2$$

3. Simplify this equation to its essentials for the canister design, and calculate the stoichiometric weights, as follows:

$$(2)(71) \times 2.2 \times 10^{-3} = 0.312 \text{ lb. } \text{KO}_2$$

$$(1.23)(44) \times 2.2 \times 10^{-3} = 0.119 \text{ lb. } \text{CO}_2$$

$$(1.5)(32) \times 2.2 \times 10^{-3} = 0.105 \text{ lb. } 0_2$$

and determine that this matches man's R.Q.;

R.Q. = 
$$\frac{\text{co}_2\text{ft}^3}{\text{o}_2\text{ ft}^3} = \frac{\text{0.119/0.123}}{\text{0.105/0.089}} = \frac{\text{0.97 ft}^3}{\text{1.18 ft}^3} = \text{0.82}$$

4. Determine the KO2 required to absorb 1.0 lb.

CO2 at 100% and 80% efficiencies;

$$\frac{0.312 \text{ lb. } \text{KO}_2}{0.119 \text{ lb. } \text{CO}_2} = \frac{\text{Xlb. } \text{KO}_2}{1.0 \text{ lb. } \text{CO}_2}$$

 $X = 2.6 \text{ lb. } KO_2/1.0 \text{ lb. } CO_2, 100\% \text{ efficiency}$ 

 $X = \frac{2.6}{0.8} = 3.2 \text{ lb. } KO_2/1.0 \text{ lb. } CO_2, 80\% \text{ efficiency}$ 

5. Determine the KO<sub>2</sub> required to produce 1.0 lb. C<sub>2</sub> at 100% and 80% efficiencies;

$$\frac{0.312 \text{ lb. } \text{KO}_2}{0.105 \text{ lb. } \text{O}_2} = \frac{\text{X lb } \text{KO}_2}{1.0 \text{ lb.O}_2}$$

 $X = 2.9 \text{ lb. } KO_2/1.0 \text{ lb. } O_2$ , 100% efficiency

 $X = \frac{2.9}{0.8} = 3.6 \text{ lb. } KO_2/1.0 \text{ lb. } O_2, 80\% \text{ efficiency}$ 

6. Assume the CO, production per man per 5 hours is;

$$\frac{2.3 \text{ lbs. x 5}}{24 \text{ hrs.}} = 0.5 \text{ lb. } \text{CO}_2$$

and the O2 consumption per man per 5 hours is;

$$\frac{2.0 \text{ lbs. x 5}}{24 \text{ hrs.}} = 0.42 \text{ lb. } 0_2$$

7. Determine the weight of KO<sub>2</sub> required for CO<sub>2</sub> absorption and O<sub>2</sub> production per man per 5 hours (at 80% efficiencies);

$$\frac{1 \text{ lb. } CO_2}{3.2 \text{ lb. } KO_2} = \frac{0.5 \text{ lb. } CO_2}{\text{X lb. } KO_2}$$

X = 1.6 lb.  $KO_2$  per man per 5 hours

and 
$$\frac{1 \text{ lb. } 0_2}{3.6 \text{ lb. } \text{ KO}_2} = \frac{0.42 \text{ lb. } 0_2}{\text{X lb. } \text{KO}_2}$$

X = 1.5 lb. KO<sub>2</sub> per man per 5 hours

Take the weight of KO, required as being the greater of these two values, or 1.6 lb., since a greater weight of  $\rm CO_2$  is produced by man than is  $\rm O_2$  consumed.

8. Since a requirement of the test program is that the preliminary test canisters operate at near steady-state conditions during at least one hour of the 5 hour test, and also considering that the weight of the KO2 in some canister packing configurations may account for only 70% of the total packing weight, we should calculate a canister volume good for double the hours of the actual test, and for a maximum of 30% inert packing weight, as follows:

 $1.6 \times 2 \times 1.3 = 4.16 \text{ lbs. } \text{KO}_2$ 

40 lbs/ft<sup>3</sup> = density of 2 to 4 mesh  $KO_2$ 

therefore,  $\frac{40 \text{ lbs.}}{1 \text{ ft}^3} = \frac{4.16 \text{ lbs. } \text{KO}_2}{\text{X ft}^3}$ 

 $X = canister volume required = \frac{4.16}{40} = 0.10 ft^3$ 

9. If we assume a satisfactory residence time of the gases in the canister to be 1 second, and a face velocity of 1 ft/sec, we need only solve for the cross-sectional area of a 12 inch long cylindrical canister as follows;

 $\pi r^2 = \frac{\text{volume}}{\text{length}}$ 

 $r^2 = \frac{0.10 \text{ ft}^3}{1 \text{ ft}}$ 

 $r = \sqrt{0.0318}$ 

 $r = 0.178 \text{ ft } \times 12 = 2.136 \text{ inches}$ 

Inside diameter of required cylinder =

D = 4.3 inches and  $D^2$  of cylinder = 18.5 inches<sup>2</sup>.

The volume of a 12 inch long cylinder is then;

Volume = 
$$\frac{(18.5)(12)}{4}$$
 = 174 inches<sup>3</sup>

or

Volume = 
$$174 \text{ inches}^3 \times 5.78 \times 10^{-4}$$
  
= 0.10 ft<sup>3</sup>

The weight of 100%  $K0_2$  that would fill up this size cylinder is;

$$W = (0.10 \text{ ft}^3)(40 \text{ lbs/ft}^3) = 4 \text{ lbs}$$

and this corresponds closely to the weight of the  $KO_2$  in step 8 which will make allowance for up to 30% inert material in the canister packing contents, and provide for about 10 hours of operation before exhaustion of the chemical.

- 10. The volume of gas which would flow through a canister of this size at a face velocity of 1 ft/sec is 6 CFM.
- 11. The engineering laboratory had on hand three plexiglass canisters (from the test in Reference 1) with dimensions of 5" (I.D.) x 10" long and it was decided to use these cylinders, and fabricate the others needed with the same inside diameter. The volume of these basic canisters is 196.35 inches or 0.11 ft3. With the same density of KO2 as in step 8, this canister could hold 4.4 lbs. of 100% KO2, or allowing for up to 30% inert material, the canister will provide about 9.6 man-hours of operation before exhaustion of the chemical. The volume of gas which would flow through a canister of this size at a face velocity of 1 ft/sec is 8.16 CFM, 5.5 CFM at 0.67 ft/sec, and 2.7 CFM at 0.33 ft/sec.
- 12. Two other sized canisters were needed for the test program where the face area/length ratio is varied. The experimental test plan indicated that with a constant face area, the length of these other two canisters should be respectively one-half and twice the length of the basic canister (in step 11). The three canister face area/length dimensions are therefore 5" x 5", 5" x 10", and 5" x 20".
- 13. The dimensions and packing configuration of the extended-duration test canister depended on data gathered during the preliminary tests.

14. If we assume the volume of air flowing through the canister to be 5.5 CFM (step 11), we can check the feasibility of 1.0 inch ducts (0.0054 ft<sup>2</sup>) to and from the transition pieces at each end of the canister, as follows;

$$\frac{5.5 \text{ CFM}}{0.0054 \text{ ft}^2}$$
 = 1018.5 FPM = 16.9 FPS

This duct velocity corresponds to recommended duct velocities for taking wet-bulb temperatures.

15. The transition pieces at each end of the canisters were designed with a half-angle of 15° to assure uniform superficial flow over the inlet face of each canister. Number 10 mesh stainless steel screens at each canister face served further as airflow straighteners and diffusers.

## Appendix B

#### TEST DATA

#### PICTORIAL REPRESENTATION

The test data is presented here in five distinct groups of tables and graphs (described on pages 89-91) accompanied by a series of photographs which serve in some measure to illustrate the degree of fusion of the solid chemical granules in each canister after each test. Inasmuch as the photography is not in color, the identification of the materials, and their appearance in the black and white photos is given as follows:

In canister A, the reacted (used)  $\mathrm{KO}_2$  granules appear as white  $\mathrm{K}_2\mathrm{CO}_3$ , as expected. However, in the E and F canisters, (Figures 39, 41) the unused  $\mathrm{KO}_2$  appears white in the photo whereas its actual color was yellow. Also easily seen (especially in canisters E and F) are the grey unreacted granules of  $\mathrm{KO}_2$  impregnated with the catalyst which is close to their actual "color". It is noticeable (especially in canister A-8, and the reacted portion of F), that the grey granules turn black when reacted, which is true to the actual "color". In canister C-2 the small white granules of LiOH are easily distinguished from the larger white reacted  $\mathrm{KO}_2$  ( $\mathrm{K}_2\mathrm{CO}_3$ ) granules, and of course the black granules are reacted  $\mathrm{KO}_2$  ( $\mathrm{K}_2\mathrm{CO}_3$ ) impregnated with the catalyst. The small white granules of LiOH are easily apparent in the C, D, A-5, and A-7 canisters. The black lava rock appears as black granules in canisters B, C and A-3, and of course are difficult to distinguish from the reacted  $\mathrm{KO}_2$  ( $\mathrm{K}_2\mathrm{CO}_3$ ) granules impregnated with the catalyst.

The activated alumina in canister B appears in its true "color" as very small grey round pellets.

The small grey rod-shaped pieces of molecular sieve in canister A-3, A-4, and A-6, are very difficult to identify. The photographs, if carefully observed, show the varying degree of fusion, and/or bonding, of the chemical granules in each canister. It is apparent that the least fusion occurred in the various A canisters. Canister C and D, both of which contained LiOH, show a greater fusion than A.

Canister B, containing activated alumina also shows greater fusion than the various A canisters. The photograph of canister E clearly shows the solid plug formation (about 1-1/2" thick) which caused complete airflow stoppage of the canister after 425 minutes. Canister F shows this same plug, with the same results. The photo of canister C-2, even though it contained LiOH, illustrates how the entire contents were turned into a soggy mass which offered greatly increased impedance to the flow of air through the canister. In the last photo, the white reacted  $\mathrm{KO}_2$  ( $\mathrm{K}_2\mathrm{CO}_3$ ) granules, and the black reacted granules of  $\mathrm{KO}_2$  ( $\mathrm{K}_2\mathrm{CO}_3$ ) with catalyst, are seen throughout the entire A-8 canister, which contrasts strongly with the F canister with its 2" plug formation at the inlet end, and its 18" long unused bed of  $\mathrm{KO}_2$  + catalyst.

## PARTIALLY REDUCED DATA

Tables 5 through 14 are the partially reduced data from the test instrumentation measurements. These are self-explanatory with some exceptions. In the test series P-1 through P-5, the grains/1b H<sub>2</sub>O columns for outlet air are not filled. This is because the data obtained from the wet bulb thermocouples was not considered correct when it was discovered that the wicking material around the thermocouple junctions did not remain properly wetted by capillary action from the wick-wetting water jars. (Figure 17). The readings from these particular thermocouples were ignored after ducts were installed for taking wet bulb readings with the Hygrophil "gun". (Figures 13 and 14). Also, after the P-5 test series, the exhaust manifold was removed (Figure 13) which lowered the pressure upstream of the canisters (P<sub>n</sub> columns). This allowed a more nearly constant flow-rate through the canisters as well as providing for a greater capacity from the blowers.

## IBM REDUCED DATA

Tables 15 through 24 consist of the data reduced by the IBM 7090 computer from the first group of tables. These tables are self-explanatory. Standard conditions in these tables are 70°F and 1 atmosphere pressure.

#### SC-4020 GRAPHS

The curves in Figures 42 through 144 were drawn by the Stromberg-Carlson SC-4020 computer from the IBM 7090 reduced data. The graphs are self-explanatory with the possible exception of Figures 108, 114, and 117. The portion of the curves below the abscissa in these figures represent negative

values, or desorption of water from the  $KO_2$  bed due to the low inlet absolute humidities. The horizontal lines below the abscissa represent the limit of the computer scale, or a truncation of the curves. The requirements for  $O_2$  consumption per man are  $\sim 0.0168$  SCFM, and the man-rate of  $CO_2$  production is  $\sim 0.0138$  SCFM, where  $70^{\circ}F$  and atmosphere pressure are taken as standard conditions. (R.Q. = 0.82).

The heavy black horizontal line across each graph of O2 produced and CO2 adsorbed represents the rate of O2 consumption and CO2 production, respectively, of one man. Of course, in a sealed cabin with the dual-canister system described on pages 16, 73 these rates would be adjusted by demand controls, and, provided the absolute humidity and partial pressure of the CO2 in the cabin air are both sufficient to assure a good rate of O2 evolution (Figure 42), the curves would approach, and tend to parallel, the man-rate line.

On some graphs, the man-rate line for CO2 appears above the grid. In these cases, the line is actually beyond the computer scale limit, but has been inserted as a reminder that the true location is somewhere above the grid. These cases in particular occurred when low absolute inlet humidites and low PCO2 were used, or when molecular sieve was used. The % inlet CO2 is repeated on each graph of O2 generation and CO2 adsorption for convenient reference because of the importance of the CO2 partial pressure in the evolution and adsorption rates of oxygen and carbon dioxide respectively. The inlet absolute humidity range is also indicated on the graphs. Because of the relatively wide range of inlet absolute humidities, reference to the exact values of absolute humidities can be made from the appropriate partially reduced data sheet. The same may be done in regard to the HoO adsorption graphs. It is seen from the curves that the H2O adsorption is, in general, proportional to the inlet absolute humidity. Also, the reader will notice that a low (0.2%) CO2 partial pressure, or a moderate (0.5%) CO2 partial pressure, is not the sole influence for a low generation-adsorption rate inasmuch as a low absolute inlet humidity may have the same effect on the curve. For example, Figures 94, 95, and Figures 106, 107 show low generation-adsorption rates, but for different reasons. Figures 106, 107 represent a moderate CO2 partial pressure with a low absolute humidity. (Table 12).

## INTEGRATED DATA

#### Cumulative Data

Tables 25 through 34 are cumulative and comparative, and were compiled (partly) by geometric integration of the areas

under the SC-4020 curves. The columns show the parameters becoming (usually) larger by successive additions as functions of time. The oxygen generated and carbon dioxide absorbed columns together represent the chemical respiratory quotient. Their ratio is presented in the CUM. R.Q. column which does not show an increase in every case (such as when the cumulation rate of oxygen is greater than the cumulation rate of carbon dioxide). The oxygen consumed and carbon dioxide produced columns together represent the normal physiological respiratory quotient of man, which is taken as 0.82.

# Man-Rate Support Level Graphs

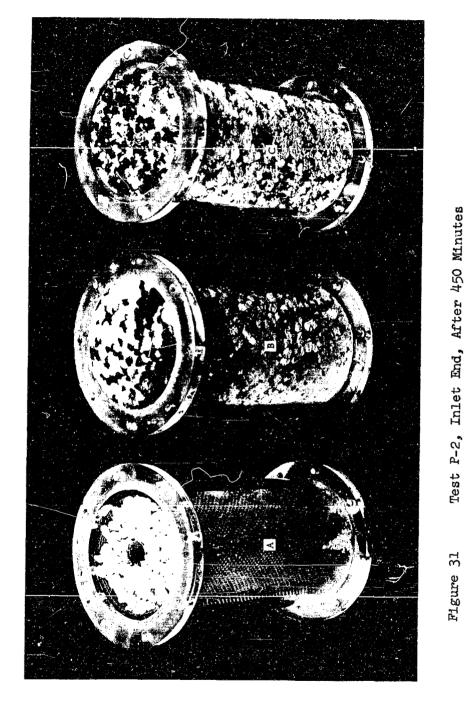
The areas under the SC-4020 curves for  $O_2$  generation and  $CO_2$  adsorption were geometrically integrated, and the results were plotted in Figures 145 through 164 with the man-rate production, or adsorption, as a function of time. Curves, or portions of curves, above the man's support level (dashed line) indicate the ability to support the respiratory requirements of one man. This ability is mainly dependent on bed packing configuration, inlet  $CO_2$  partial pressure, and inlet absolute humidity.

## FURTHER ANALYSIS OF TEST DATA

All test data of any importance has been included in this report to encourage the individual reader towards further study and analysis of the figures, tables, and graphs, inasmuch as it was beyond the capabilities of the author to extract all possible information in the time allotted to do so.

Test P-1, Inlet End, After 250 Minutes

Figure 30



Test P-2, Inlet End, After 450 Minutes

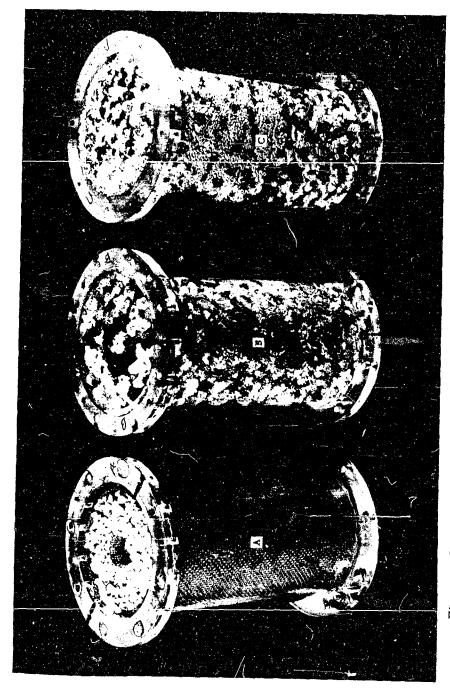


Figure 32 Test P-3, Inlet End, After 450 Minutes

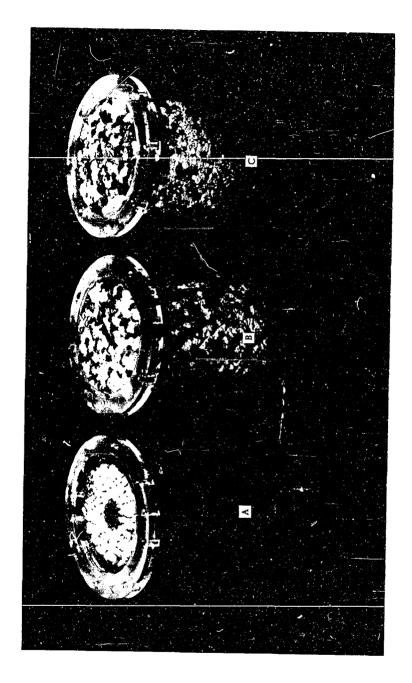


Figure 33 Test P-4, Inlet End, After 255 Minutes

Figure 34 Test P-5, After 337 Minutes

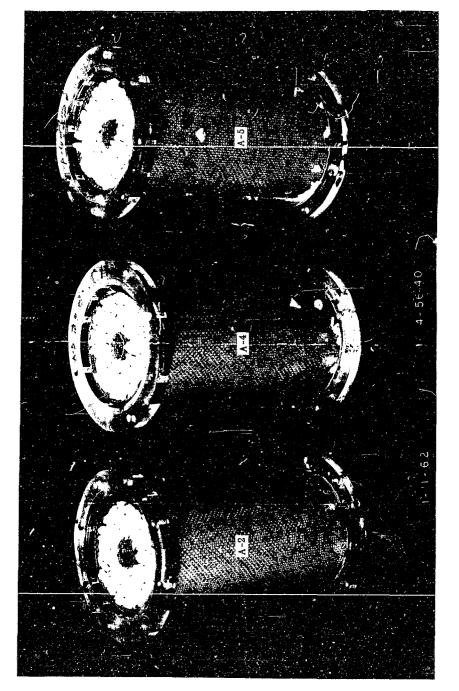


Figure 35 Test P-6, Inlet End, After 461 Minutes

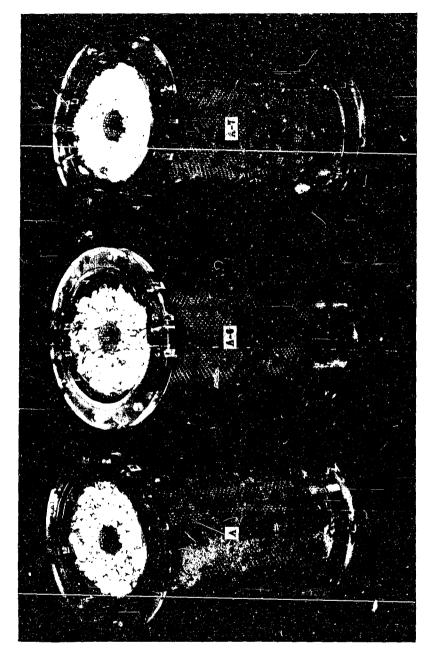
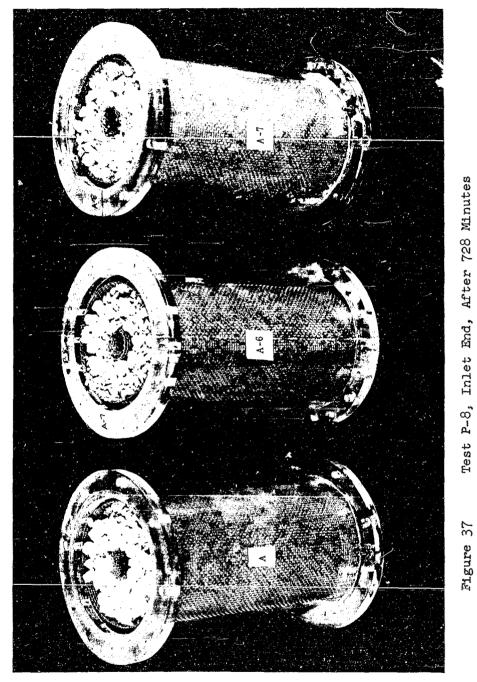


Figure 36 Test P-7, Inlet End, After 435 Minutes



Test P-8, Inlet End, After 728 Minutes

Test P-9, Inlet End, After 650 Minutes

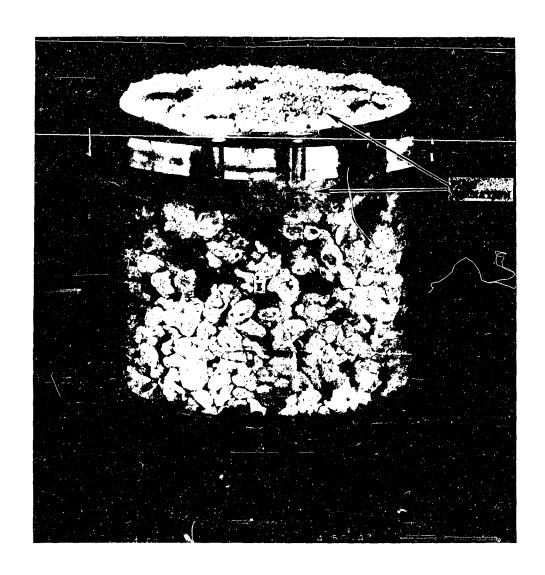


Figure 39 Test P-10, Inlet End, Canister E, After 425 Minutes

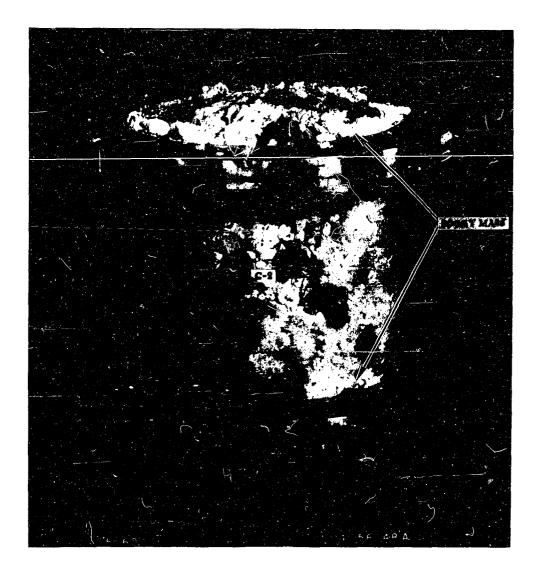


Figure 40 Test P-10, Inlet End, Canister C-2, After 680 Minutes



Figure 41 Test P-10, Inlet End, Canisters A-8 After 1765 Minutes, and F After 425 Minutes

Table 5 Partially Reduced Data, P-1

Date: 12 - 18 - 61

Test

95.0 8. 9.79 1.56 0.73 .0 2.00 20.12 8 8 16.02 ₹ 8 21.54 0.75 22.12 SV VR OUTLET AIR 0.83 28.0 0.55 0.83 08.0 98.0 8,\* 9810 79.4 #:# 79.1 5.56 95.4 35.1 35.5 MOLE E E E CANGESTER B
1/3 KO + Cat. + 1/3 Lava Rock
+ 1/3 Activated Alumina 0 P 2 æ æ 4 % E 覆 ন 0.30 77 , † A. 7.70 0.25 76 H 0.30 92 1 0.30 77 H In.Eco In.Eco 08.9 25.0 0.25 9 7.80 7.80 3.80 7.88 7.80 7.85 0.38 21.54 0.65 21.39 0.85 21.30 0.69 0.69 23.27 0.67 0.73 20.07 1.60 0.75 21.12 0.61 0.72 21.21 0.75 21.15 0.90 ö S VR OUTLET AIR 55 22 8 8 × EZ. 4.47 FLOW C Pa 24.4 4.48 4.47 4.52 4.63 4,66 CANISTER A KO2 + Cat. + Annular Screen A. ř Ħ 9 8 듉 8 Ħ 35 92 92 0.97 20.76 8.15 0.32 77 Eio. H Pro APu In.H20 In.H20 g q 1.01 20.85 8.05 0.30 0.30 0.99 20.85 8.05 0.30 20.79 8.05 0.33 61.4 0.97 20.85 8.05 0.30 20.82 8.10 0.32 8,00 Shut D 61.8 1.01 20.85 8.05 1.05 21.12 8 % 0.97 1001 6.73 INLET AIR g, se H20 Gr/Lb. 47.1 68.5 61.4 60.5 68.1 100 M.n. 8 र्च क 125 죓 245 S 215 22

 $P_{\rm B}$  = Static pressure just upstream of canister  $T_{\rm B}$  = Temperature just upstream of canister

 $\Delta P_{\rm R}$  . Pressure drop across canister  $\Delta T_{\rm R}$  . Temperature rise across canister

Table 5 (cont.) Partially Reduced Data, P-1

		ж .о		0.56	0.71	0.65	8.	8	1.13	5	0.51		Ì					
	٥	ું,≉		93.0	21 54 0.71	21.36 0.65	21.18 0.90	00.1	21.00 1.13	20.94	20.04 0.51							
	OUTLET AIR	8,14		6	62	179.0	0.70	37.0	8	a c	98.0							
	8	H20 Gr/Lb																
MSA		FLOW		4.15	4.15	4.15	4.15	4.15	4.15	4.15	4.15							
CANESTER		ΔT.		65	69	59	177	38	#	27	25							
CAN		€1.0 ©\$4		28	78	78	78	7.1	78	78	78							
		APrt In.H20		1.10	1,10	1.10	1,15	1.10	1.15	1,15	1.20							
		Pie APir In.H2O		7.40	7.60	7.60	7.60	7.60	7.50	7.60	7.70							
		R.Q.		0,81	1.02	1.17	1.74	2.47	3,56	3.00	2.00					1		
	Fi	02 82		21.57	21,36	71.1 12.12	31.06	21.00	20.94	20.91	20,88				-			
	OUTLET AIR	Z <sub>0</sub> 05 ★		0.27	0.49	0.55	0.58	0.62	0.65	0.70	0.73	_						
	Ю	97/25 H				_												
υ υ		MOTA		4.75	4.75	4.76	4.76	4.76	4.76	4.76	4.76							
CANISTER	HO.	ΔT3		19	345	47	715	24	33	27	24							
١.	1/3 E	는*		76	76	77	77	77	77	77	77							
1/3 KO <sub>2</sub> + Cat.	+ Lava Rock + 1/3 L10H	<b>AP</b> 15 In.H <sub>2</sub> 0		0.30	0.35	0,40	24.0	0.45	0.45	0.45	0.50	DOWN						
1/3 KC	+ Lave	P.4 AP.15 In.H20 In.H20		7.45	7.55	7.65	7.60	7.60	7.60	7.60	7.70	SHUT						
_		8,8	21.12	21.00	20.85	20.85	20.79	20.87	20.85	20.82	20.76							
INLET AIR		602 %	1.05	0.73	1.01	0.97	1.05	0.99	0.97	0.97	76.0							
NI		H20 Gr/Lb.		53.1	47.1	48.3	68.5	61.4	61.4	60.5	68.1							
	TURE	Min.	0	5	앜	8	115	145	175	205	235	250						

 $P_{\Omega}$  . Static pressure just upstream of canister  $\gamma_{\Omega}$  . Temperature just upstream of canister

 $\Delta P_{\rm R}$  = Pressure drop across canister  $\Delta \Pi_{\rm R}$  = Temperature rise across canister

Table 6 Partially Reduced Data, P-2

		, ,		9,46				0.67	0.58	0.53	0.49	₹. 0	0.62	94.0	19.0	0.73	95.0	0.75		
	ر ه	84.00		21.93	21.42	21.24	21.09	20.91	20.82	20.88	20.79	20.73	20.67	20.70	20.6	20.64	20.61	20.61		
	OUTLET AIR	8,*		0.21	0,55	19.0	69.0	0.75	0.73	8.0	0.83	98.0	98.0	0.88	0.88	0.88	0.88	0.90		П
	B	47/19																		
B	¥	MOTAL MOTAL		4.15	3.75	3.86	3.81	3,75	3.48	3.30	2.01	3.18	3.18	9.8	3.12	3.18	3,16	3.18		П
CANESTER	+ 1/3 Lava Rock utna	ro F		27	201	69	&	7	62	55	6#	94	44	41	37	34	32	30		
CAN	+ 1/3	†Bi H°°		7.4	69	2	77	7	69	69	20	S S	70	8	20	20	70	70		
	1/3 KO2 + Cat. + 1, + 1/3 Act. Alumina	APis In.H20		0,75	0.85	0.95	1.10	1.30	1.35	2905	81	1.85	1.95	20.2	2.08	2.10	2,03	5.06		
	1/3 Kg	Fr. Dela In. Ego		8,60	9.10	9.00	9.00	9.15	9.15	9.15	9.25	9.25	9.30	9.35	9.35	9.35	9.24	9.30		
		73. C.		0.48	29.0	0.48	4.0	29.0	0.48	0.45	24.0	£.0	84.0	94.0	0.46	5,5	0.73	0.50		
	[ _	S, A		21.42	21.27	21.12	21.12	20.02	20.94	20.88	20.76	20.73	20.73	20.70	20.70	20.70	20,64	20.64		П
	OUTLET AIR	8 ×		0.43	64.0	69.0	0.73	9Z-0	0.76	8.0	0.85	98.0	98.0	88.0	98.0	0.88	88.0	0.90		
	б	H20 Gr/Lb																		
<b> </b>		FLOW		4.22	4.10	4.22	4.23	4.36	49.4	4.50	4.88	4.90	6.07	6.07	6.07	6.07	50.0	6.07		
CANISTER	+ Annular Screen	ΔT,		å	65	62	27	25	7.17	3	070	37	ř	F	30	83	92	96		
3	Annular	E-i-e		92	69	69	72	r P	22	89	69	69	69	8	69	89	89	88		
	Cat. +	A.n. In.		8.	0.80	0.80	0.82	0.85	0.88	0.95	0.90	1.02	1.07	1.15	1:15	1.15	1.20	1.21	IOWN	
	KO2 + Cat.	Fo In.H20 In		8.65	9.10	8.95	8.95	8.92	8.75	8.75	8.75	8.65	8.70	8.65	8.75	8.75	8.73	8.72	SHUT	
		જ ૪	20.52	20.55	20.52	20.58	20,49	20.58	20.46	20.46	20.46	20,49	20.46	20.46	20.46	20.49	20.49	20.49		
INLET AIR		S, №	0.97	0.85	65.0	0.95	65.0	75.0	8.0	0.99	0.99	0.39	0.99	96.0	0.99	99	0.90	0.99		
		H20 Gr/Lb.	99	66.5	69.3	0.69	73.0	67.7	70.8	0.99	66.1	68.1	65.5	70.3	5.99	68.8	8,99	6.69		
	TOTA	M.n.	٥	2	45	12	205	133	165	195	225	255	285	315	345	385	405	435	450	

 $P_{\rm R}$  = Static pressure just upstream of canister  $T_{\rm R}$  = Temperature just upstream of canister

 $\Delta P_{\rm R}$  = Pressure drop across canister  $\Delta \Gamma_{\rm R}$  = Temperature rise across canister

Table 6 (cont.) Partially Reduced Data, P-2

	_	R. 6.		0.49	21.24 0.51	0.55	0.67	1.78	20.70 0.58	0.78	0.50	0.47	0.47	20.61 0.40	9	20.58 0.45	20.58 0.45	0.45			
	E E	જ જ	Γ	2,5	21.24	21.00	20.73	20.67	20.70	20.64	20.64	20.6	20.61	20.61	19.02	20.58	20.58	20.58		Π	
	OUTLET AIR	8 *		0.38	0.62	0.72	0.83	9.83	0.85	0.85	%.0	8	26.0	0.93	8	9.95	0.95	90			
	8	H20 Gr/Lb																			
MSA	_	FLOW		3.8	3.88	3.96	3,98	4.06	4,11	3.84	3.84	3.84	3.84	3.84	3.84	3.84	4.11	18			
CANTSTER		TQ.		105	g	61	4.7	45	31	27	23	22	21	20	17	16	15	15			
Š		<u>Θ</u> β.,	Γ	75	72	72	75	73	7.	71	7.1	71	77	7.3	77	71	n	7.7			
		APA In.E20		1.32	30	1.30	1.30	1,35	1.40	1.45	1.50	1,55	1.55	1.70	1,75	1.85	1.90	1.95			
		In.H20 In.H20		8.20	8.75	8.05	7,90	B,00	7.85	7.85	7.80	7.80	7.85	7.30	7,92	7.90	8.05	8.05			
		ы; Су.		0.62	92.0	0.83	8	1.55	96.0	0.78	0.79	98.0	6.93	0.87	0.73	0.75	82.0	0.78			
	ĸ	°, 8		18,13	नट रट	30,15	20.82	20,76	20.73	20.73	20.70	20.70	19.02	20.61	20.61	20.61	20.58	20.58			
	OUTLET AIR	° 8 8 ° 8 ° 8		20.05	44.0	0.55	99'0	69.0	0.73	0.78	80	0.81	0.85	98.0	0.88	8.	0.92	0.92			
	б	чт/жо ж																			
0		FLOW		51.5	4.86	4,69	4.65	4.47	4.46	4.86	4.86	4.45	4.45	4.46	4.46	4.33	4.50	4.51			
1/3 KO <sub>2</sub> + Cat. + Lava Rock		ξŢό.		29	211	8	99	55	48	44	9	37	34	31	28	25	23	22			
t. +		13. 14.		75	69	20	72	7	69	9	69	Ş	69	69	69	69	69	20			
80 + co	12 HO	APis In-H <sub>2</sub> 0		0.75	0.85	0.1	1.10	1.20	1.23	1.35	1.43	1.50	1.58	1.70	1.80	1.8	1.82	1.80	DOWN		
1/3 K	+ 1/3 140H	Fl4 In.H20 In		8.50	8.70	8.60	8.60	8.70	8.60	8.60	8.60	8.60	8.68	8.70	8.80	8.75	8.58	8.70	SHUT		
		8, 8	20.52	20.55	20.52	20.58	20.49	20.58	20.46	20.46	20,46	20.49	20.46	20.46	20.46	20.49	20.49	20.49			
INLET ATR		S &	0.97	0.85	0.99	0.95	0.99	76.0	0.99	0.99	0.99	0.99	0.99	0.99	0.99	0.99	88.0	0.99			
INI		H2O Gr/Lb.	9	66.5	69.3	69.0	73.0	67.7	70.8	0.99	66.1	68.1	65.5	70.3	66.2	68.8	8.99	66.69			
	TORE	M.n.	٥	12	15	&	115	140	170	8	230	262	290	350	350	377	408	애	Ç,		

Pn = Static pressure just upstream of canister In = Temperature just upstream of canister

 $\Delta P_{\rm R}$  = Fressure drop across canister  $\Delta T_{\rm R}$  = Temperature rise across canister

Table 7 Partially Reduced Data, P-3

12-20-61	
Date	
P-3	
Test	

1 400	CANISTE	CANISTE	CANISTER	TSTE Company	2 (	_					1/3 KG	1/3 KO <sub>2</sub> + Cat. + 1/3 Lava Rock	CAN:	CANISTER 1/3 Lava R	zk Zk				
		¥02 +	Car. +	Annutar	Screen	1	g	OUTLET ALR	6	7	+ 13	Act. Al	umina			8	OUTLET AIR		
જ ૪૨		Flo In	AR, In.H20	Eig.	ΔT,	FLOW	H20 Gr/Lb	20°2 *	0, PS	R.Q.	P. A. L.	AP. In.H20	†Br Ero	27. F.	FLOW	R20 Gr/Lb	8 ×	જ,*	R.Q.
ŏ	20.61																		
_~	20.64	8.75	0.27	8	65	5.26		15.0	21.36	0 50	9.25	1.05	88	20	3.72		0.35	27.12	0.48
17	20°6	8.75	0.27	8	98	5.31		0.64	21.33	0.45	9.20	1.07	%	87	3.75		0.55	23	5
- 1	20.64	8.60	0.30	ā	62	5.31		29.0	21.27	0.48	9.20	1.10	ま	69	3.76		0.63	23.38	0.47
- 1	20.6μ	9.60	0.30	32	8	5.31		0.72	21.27	0.40	9.20	1.10	95	62	3.77		79.0	2.2	0.48
	20.59	8.50	0:30	88	25	5.40		0.77	21.15	0.43	9.10	1.15	98	52	3.77		47.0	2.16	0.47
	20.59	8.40	0.30	26	45	5.44		0.77	21.06	0.51	9.00	1.20	%	94	3.77		7.0	21.00	0.59
	20.61	8.35	0.30	700	37	5.47		080	21.03	0.50	00.0	1.20	98	39	3.77		85	8	5.58
	20.61	8.30	0.30	18	34	5.50		0.81	20.97	0.56	9.00	1.20	66	32	3.72		98.0		0.62
	20.61	8.20	0.29	100	8	5.49		0.81	20.88	0.74	8.93	1.23	99	28	3.60		0.85	20.89	0.67
	20.64	8.15	0.30	8	98	5.50		0.85	20.88	0.63	8.30	1,30	66	25	3.60		0.87	20.76 1.08	90.
	20.61	8.17	0.29	99	23	5.52		0.85	20.85	0.67	8.93	1.30	98	21	3.56		98.0	20.79	0.72
_	20.61	8.15	0.30	2.5	25	5.57		0.88	20.76	0.87	8.95	1,33	96	21	3.56		8.	20.76 0.73	7
_	20.49	8.07	0.29	33	17	5.48		0.88	20.70	0.62	8.92	1.39	8	18	3.56		8.	20.670.61	9
-+	20.49	8.03	0.30	8	15	5.48		0.32	20.61	0.75	8.95	1.40	8	15	3.56		6.93	20.62 0.67	.67
_	20.49	8.07	0.29	8	7.	5.49		0.92	20.58	3.0	8.96	1.42	æ	#T	3.56		8	20 61 0.67	3
	20.52	SHUT	DOWN																
BOILER	DRY																I		Ī
-				Γ	T		Ī	T	T	Γ		Ī	ľ	T	Ţ	Ī	Ť	1	T

 $P_{\rm n}$  = Static pressure just upstream of canister  $\gamma_{\rm n}$  = Temperature just upstream of canister

 $\Delta P_{\rm R}$  = Pressure drop across canister  $\Delta T_{\rm R}$  = Temperature rise across canister

Table 7 (cont.) Partially Reduced Data, P-3

12-20-61

5

		R.Q.	ĺ	5.0.44	21.30 0.53	0.40	9	0.62		0.72	95.0	20.75 0.92	1.22	20.73 1.08	20.73 1.08	5	0 78	8			
	×	°,*		2.15	21.30	2. 2	51.0	20.88	20.82	20.74	20.73	20.7	20.73	20.73	20.73	20.58	20.58	20.5=1.00		Γ	
	DUTTET AIR	205		5.5	9.0	0.70	0.78	0.83	98.	0.88	9.0	3.	6.0	0.88	0.88	80.0	9.	5.05			
	ρo	47/19 Gr/79									Γ									T	
MSA		FLOW		10. 17	4.12	90.4	4.09	4.06	4,02	4.06	4.09	4.15	4.11	4.11	4.09	4.13	4.11	4.13			
CANISTER		ΔT4 P		5	99	5	17	31	25	20	16	13	12	10	o	7	9	ư			
CAN		ET. OPT		%	8	83	95	8	98	97	26	98	86	76	96	22	26	40			Γ
		Ln.H20		1.25	1.20	1.20	1.22	1.20	1.20	1.25	1.25	1.27	1.30	1.29	1.32	1.35	1.40	1.37			
		oz∺.nī		8.25	8.20	8.15	8.10	8.00	7.85	7.85	7.80	7.81	7.80	7.81	7.80	7.76	7.78	7.78			Γ
		R.Q.		0.77	0.76	1.15	1.33	1.48	1.40	1.39	1,50	1.20	3.8	1.25	1.55	1.25	1.25	1.22			
	es.	O, 84		21,69	42, 12	21.03	20.91	20.82	20.79	20.79	20,73	20.76	20.70	20.73	20.Zc	20.61	20.58	20.58			
	OUTLET AIR	8 ×		0.00	0.34	0.52	0.61	0.67	0.73	0.76	0.83	0.83	0.83	98.0	0.87	98.0	0.90	%			
	δ	H20 Gr/Lb																			
Sk Sk		FLOW		4,66	4.62	4.63	4.65	4.65	4.65	4.64	49.4	4.64	494	4.63	4.67	4.67	4.71	4.71			
1/3 KO <sub>2</sub> + Cat. + 1/3 Lava Rock		ors		103	77	59	44	33	83	23	18	14	7	9	10	7	2	3			
57		44 44		98	35	16	97	97	26	100	86	100	100	93	97	88	8	99			
)2 + Cat	LAOH	Pie APis In.H20 In.H20		0.60	0.77	8.	0.82	0.85	0.85	0.85	0.85	0.87	0.90	0.89	0.90	ъ. О	0.95	946.0	DOWN		
1/3 K	+ 1/3 EAOH	Pi4 In.H20		8.45	8.50	8.45	8.40	8.35	8.25	8.20	8.15	8.10	8.10	8.10	8.10	8.10	8.07	8.05	SHUT		
		8, A	20.61	20.64	20,64	20.64	20.64	20.59	20.59	20.61	20,61	20.61	20,64	19.02	20.61	20.49	20.49	20.49	20.52	DRY	
INLET AIR		00 %	1.01	0.90	0.95	0.97	0.97	10.1	1.01	1.01	1.01	1.01	1.01	10.1	1.01	1.01	10-1	1.01	10.1	BOILER	
NI SE		r20 Gr/lb.	65.8	63.6	54.8	56.2	61.1	59.2	75.9	58.1	57.8	58.4	2.95	56.5	90.09	58.1	55.7	55.7	27.6		
	TIME	Min.	0	gg	65	83	110	140	170	200	230	560	230	320	350	385	410	왉	450		

 $P_{\rm h}$  = Static pressure just upstream of canister  $r_{\rm h}$  = Temperature just upstream of canister

 $\Delta P_n$  " Pressure drop across canlater  $\Delta T_n$  " Temperature rise across canister

Table 8 Partially Reduced Data, P-4

		H.O.		9.63	0.53	ە ئ	0.53	99.0	92.0	0.65								
		9,40		21.45	21.42	21.36	21.27	42.12	81.19	21.18								
	OUTLET AIR	S &		0.17	16.0	0.28	0.32	0.32	0.33	0.36								
	οo	H20 Gr/Lb																
M		PLOW Cfm		4.39	GE 17	4.32	4.19	61.4	4,19	4.19								
CANISTER	v	ST2		31	917	94	75	Off.	36	45				_				
CAN	umina	€1.0. 4 br		9/	2.2	92	92	92	92	92								
CANISTER 1/3 KOn + Cat. + Lava Book	+ 1/3 Act. Alumina	APIS In.H20		04.0	94.0	0.55	0.65	0.80	0.85	06.0								
1/3 KO	+ 1/3	Fig AP's In.H20 In.H20		04.8	24.8	8.50	8.55	9.80	8.62	8.60								!
		я. 9.		69.0	24.0	0.42	94.0	95.0	0.55	15.0								
	ATR	o 80		27.36	21.36	21.34	21.30	22, 27	21.24	12.12								
	OUTLET A	205 *		0.27	0.30	0.34	0.33	0.33	0.35	0.37								
	°	H20 Gr/Lb																
4		FLOW		49.4	49.4	4.89	4.89	4.99	5.21	5.32				_				
CANTSTER	KO2 + Cat. + Annular Screen	AT.		32	4.7	35	39	36	33	31	L							
19	Annula	£40.		75_	92	92	76	92	92	92	L				 			
	Cat. +	Δ <b>Ρ</b> ι Σπ.πΣ		0.30	0.36	0.35	0,40	0.40	0.42	0,40	NMOG							
	K02 +	Pio In.H2O In.		8,50	8.50	8.45	8.50	8.40	8.40	8.40	SHUT							
		8 0	20.88	20.88	20.91	20.91	20.91	20.95	20.95	20.95	20.95							
INLET AIR		چ 205	0,52	0.51	0.51	0.52	0.51	0.51	0.51	0.51	0.51							
Ä		H20 Gr/Lb.	91.4	91.4	61.6	62.0	62.2	63.6	61.6	61.2	61.2							
	TIME	Man.	٥	80	65	95	140	180	218	242	255		j					

 $P_{\rm B}$  = Static pressure just upstream of canister  $T_{\rm B}$  = Temperature just upstream of canister

 $\Delta P_{\rm B}$  = Fressure drop across canister  $\Delta T_{\rm B}$  = Temperature rise across canister

Table 8 (cont.) Partially Reduced Data, P-4

			, e	Ì	4.18	21 42 0.53	977	21.01 0.43	93 - 0 - FC	0.00	1 1									
		Я	on,*e	Ī	5	2 10	5	2 2	מר נפ	1	2, 0 0, 10	9				Ī				
		OUTLET AIR	8,8		5	77.0	6	9,38	0 30	5	2	7								
		no	H20 Gr/Lb										Ī	Ī	Ī					П
	ISA.		PLOW cfu	T	17.77	, ,	4.70	4.03	90.4	88	-8								Γ	П
	CANTSTER MSA		4E4.		77	S	14	35	õ	92	25						T			
	CAN		Q <sub>D</sub> ,		20	, a	77	ŧ	1	2.2	0,2	-								П
			APr In.R20		1.20	1 27	1.36	2	1.55	1.85	8.0									
			Pie APir In.H20 In.H20		8.5	7.60	8.09	00,00	8.30	8.40	č.									
			R.Q.		0.46	0.73	0.76	0.92	1.10	1.30		Ī								
		ps.	0" k2		21.00	23.51	21.45	05.19	21.24	91.16	21.15						1			П
		OUTLET AIR	202		٥	0.07	0.11	0.15	0.19	16.0	42.0									
		ŏ	H20 Gr/Lb																	
	U		FLOW of in		4.74	†9°†	4.64	49.4	4-53	4.53	4.53									
i	CANISTER		oF.		18	75	69	56	6†	74	1,1									
	;		F+0		77	92	76	75	76	92	92									
2-19-61	+ cg	+ 1/3 глон	<b>ΔΡισ</b> Ιπ.Η <sub>2</sub> ο		02.0	0.81	0.94	1.10	1.21	1.30	1.37	NMOQ								
Date: 12-19-61	1/3 KK	÷	P. APIS In.H20 In.H20		7.90	8.40	8.36	20.91 8.14	8.55	8.62	8.55	SHUT								
4 			8,18	20.88	20.88	20.91	20.91	20.91	20.95	20.95 8.62	20.95	20.95								
P-4	INLET AIR		S &	0.52	0.51	0.51	0.52	_	0.51	0.51	0.51	0.51								
	Ä	{ _[	H20 Gr/Lb.	4.16	4.16	61.6	62.0	62.2	63.6	61.6	61.2	61.2								
fest		TIME	Min.	0	5	ß	8	135	175	213	237	552								

 $P_{\rm R}$  = Static pressure just upstream of canister  $\gamma_{\rm R}$  = Temperature just upstream of canister

 $\Delta P_{n}$  = Pressure drop across canister  $\Delta T_{n}$  = Temperature rise across canister

Partially Reduced Data, P-5 Table 9

	_	1	T	Τ.	Τ.,	1	Τ.,	T	Τ.	Τ.	Ι.	Ι.,	1	Τ.	Т	Т	т	1	Т	Т	Т
	_	B. 6.	L	6			j 0	0					-								
	E P	જજ		9	2 6	20.12	8	8	20 SF	6	8	20 62	8	8	2						
	OUTLET A			0.24	95.38	0.38	0.30	14	c#.0	0.40	c I	3	0.42	-	,						T
	B	H20 Gr/Lb									Γ									T	T
A-3		PLOW		4.37	4.37	4.37	4.37	4.37	67.4	4.40	9 4	6 <del>1</del> .	4.64	11 64						Ī	T
CANTSTER	7: E +	ΔT2 P	I	8	30	36	33	30	23	8	8	8	61	یو							T
8	KA KOG	Ho.	T	2	8	79	79	88	62	5	5	6	85	ę							T
	creen	ΔP,3 n.H20	T	0,42	0.42	0.42	0.42	0.45	0.45	0.45	0.47	24.0	24.0	0.47						T	t
1	Ave + vat. + Lava Rock + M.S. + Ann. Screen	In.H20 In.H20	T	8.25	8.30	8.30	8.35	8.30	8.25	8,20	8.25	8.25	8.25	8.20	├	-			T	l	t
	4 7	 	t	0.47	0.35	0.37	0.33	0.40	0.44	0.54	0.41	0.42	0.37	95.0		-	-		l	T	t
		o" ve	T	21.45	21.42	22	21.21	21.12	21.03	20.91	20.88	20.82	20.79	20.67	_				-	r	+
	OUTLET AIR	8 %		0.30	0.29	- 65.0	0.33	9:34	35.0	0.38	0.38	0.39	0.39	0.41		-	-	-			T
	8	H2O Gr/720			1	1	1	7	+	-	1	1						-	-	-	H
A-2		PLOW Cfm		4.58	61.4	4.58	4.58	011	85.4	4.58	4.67	4.69	4.69	4.69				_	۰	$\vdash$	<del> </del>
CANISTER	arse re	Ω. ¥.		38	E.	9	7	6	e e	34	8	F	30	20	_					-	
CAN	- ALIE	βt, €10		2	8	78	38	æ	79	22	æ	78	38	78							-
1 48	2 (in car.) + Annuar Screen	4. 1. 1. 1. 1. 1. 1. 1. 1. 1. 1. 1. 1. 1.		o.40	9.40	040	24.0	24.0	04.0	0.45	24.0	24.0	0.42	0.42	NAMO			-			
٤	7 2	Pio APH In.H20 In.H20		8.35	8.45	8.35	8.30	8.45	8.35	8.30	8.35	8.35	8.35	8.30	SHUT		7				
		3" ×	20.76	8	20.70	20.70	20.67	20.70	20.67	20.67	20.61	20.58	20.52	20.49	20.49		7				
INTERT AIR		S ×	0.51	15	50	15-0	650	15.0	0.51	Ľ.	64.0	0.49	64.0	٠ د	0.51	1	1	1		_	
ă		H20 Gr/T.d.	79.9	20.0	76.6	9 2	8.6	70.6	61.4	61.3	4.69	70.1	_	72.0	62.6	1	1	1	1		
		M.n.	°	7.7	7.1	88	20	132	167	202	122	222	287	31,	337		7	1			

 $P_{\rm D}$  =. Static pressure just upstream of canister  $T_{\rm D}$  = Temperature just upstream of canister

 $\Delta P_{\rm R}$  . Pressure drop across canister  $\Delta \Gamma_{\rm R}$  = Temperature rise across canister

Table 9 (cont.) Partially Reduced Data, P-5

Date: 12-29-61

Test P-5

		 	1	9,46	11	1	0.40	± ₹	0.50	1.33	79.0	0.53	5.53	4	1	T	Τ	Τ	Τ	Τ	1
	<u></u>		╀	21.54 0	0 647.12	0 6.19		20.9	20.91	20.76	0.76	_		8	-	+	+-	+	╀	+	-
	ş	જજ	L	ដ	2	5	12	8	å	8	8	20.73	20.67	8		L	L				
	OUTLET A	% <sup>20</sup>		0.17	0.23	0.30	0.34	0.38	0.39	0.30	0.39	0.41	0.41	07/0							
	8	H20 Gr/Lb																			
MSA		FLOW	Γ	3.65	3.65	3.65	3.65	3.65	3.65	3.61	3.61	3.61	3.44	07,							
CANESTER		4±4°		98	29	45	45	41	27	25	25	54	23	٥					Ī		
CAN		Et e	Γ	8	&	80	80	8	8	80	80	80	62	8			Γ				
		APr. In.H20		1.35	1.35	1,35	1.35	1.45	1.50	1.55	1.68	1.80	2.05	2.35							1
		Pis APr In.H20		7.85	7.90	7.85	7.85	8.00	7.90	7.95	7.90	8.15	8.35	8.75							
		oj.	T	0,40	0.53	0.63	0.56	0.75	17.0	1.14	0.92	8	0.90	1.56		Γ				T	Ī
		0, A		21.87	21.54	21.39	21.24	21.06	21.00	20,88	20.85	20.79	20.73	20.67	-	$\vdash$		-		-	
	OUTLET AIR	Z 205 ¥		90.0	11.0	0.13	0.19	0.24	72.0	0.27	0.27	05.0	05.0	0.33	<b>-</b>	-		<del> </del>			t
	ωo	H20 Gr/Lb	-				7	1	1		1	1	7		-	_	$\vdash$		-	$\dagger$	f
A I		FLOW of the		4.66	19-8	4.58	4.62	4.62	99*#	99.4	4.66	99-1	4.70	4.66	_		-		$\vdash$	-	-
CSTIER (40H)		oF 1	H	75	88	E	65	8	7	47	SE SE	37	34	30	_		-		$\vdash$	-	
+ 1/3 I		T	Н	79	8	8	22	8	22	22	22	82	- P	SZ ZZ		_		-	H	-	-
2/3 KO <sub>2</sub> + Cat. + 1/3 L40H	Probes	APie n.H20	Н	0.45	0.48	0.52	0.55	0.60	80.0	0.62	0.65	0.68	02.0	0.75	DOWN				-	-	-
2/3 KO <sub>2</sub>	+ Brass Probes	Pi4 APis In.H20 In.H20		7.85	3.90	7.85	2°8	7.95	2.80	7.85	7.80	7.90	8.8	8.10	SHUT I				_	<del> -</del>	L
		Sh Ar	20.76	20.73	20.79	20.79	20.67	20.70	20.67	20.67	20.61	20.58	20.52	8 64.02	20.49			_		-	-
INLET AIR		දි <sup>1</sup> /ඳ	0.51	₹ 0	15.0	0.51	0.51	0.51	16.0	0.51	64.0	64.0	64.0	0.51	0.51				_	_	ŀ
INTE		H20 Gr/Tb.	6.62	6.67	9.92	70.0	9.08	9.02	61.4	61.3	4.69	70.1	77.12	72.0	62.6						L
		Man.	٥	22	22	43	701	137	2	502	232	257	265	322	337						-

 $P_{\rm B}$  = Static pressure just upstream of canister  $T_{\rm B}$  = Temperature just upstream of canister

 $\Delta P_n$  = Pressure drop across canister  $\Delta \Gamma_n$  = Temperature rise across canister

Table 10 Partially Reduced Data, P-6

THE HEAD COP COP COP TO THE PARTY SOURCE ALTHOUGH THE TOTAL CONTINE ALTHOUG	Test	94		e i	Deter	1-10-05	N N	1								{	- 1			}		Γ
Hgo cog cog sin Hgo lin.Hgo fig. Trow correct ATC correct ATC cog	ſ						ČÝŠ	ISTER A	52							CANT	4-4 A-4				1	
HgO         GO2         O2         Pro         ABIN         Tr         ATIN         Tr         ATIN         ATIN <th></th> <th>H</th> <th>THE ALTH</th> <th></th> <th>KO<sub>2</sub> (n</th> <th>o catal</th> <th>yst) + A</th> <th>nn. Son</th> <th>e E</th> <th>į</th> <th></th> <th>٢</th> <th><u>01.</u></th> <th>/3 K02</th> <th>(no cata Serem</th> <th>lyst)</th> <th>. 1/3 1/3</th> <th>, ,</th> <th>TOO</th> <th>OUTLES AIR</th> <th></th> <th></th>		H	THE ALTH		KO <sub>2</sub> (n	o catal	yst) + A	nn. Son	e E	į		٢	<u>01.</u>	/3 K02	(no cata Serem	lyst)	. 1/3 1/3	, ,	TOO	OUTLES AIR		
Harden   Code	Ħ						- 1	┝	7-	4			<del> </del>		2	Ė	Ę	7	100	క్ట		
18.1 0.10 21.00 0.20 0.10 73 2.78 15.3 0.10 21.12 834 0.20 0.20 73  18.1 0.20 21.00 0.20 0.10 71 1.5 17.5 2.78 12.9 0.13 21.18 446 0.20 0.20 71.5  18.4 0.20 21.00 0.20 0.10 71.5 17.5 2.78 12.9 0.13 21.18 446 0.20 0.20 72  18.3 0.20 21.00 0.20 0.10 72.5 16 2.78 12.9 0.13 21.18 446 0.20 0.20 72  18.3 0.20 21.00 0.20 0.10 72.5 16 2.78 12.9 0.13 21.15 451 0.20 0.20 72  18.3 0.20 21.00 0.20 0.10 73 14 2.78 12.9 0.13 21.15 451 0.20 0.20 73  18.4 0.20 21.00 0.20 0.10 73 14 2.78 12.9 0.13 21.15 451 0.20 0.20 74  28.0 0.20 21.00 0.20 0.10 73 18.5 2.78 12.5 0.11 21.12 432 0.20 0.20 74  11.8 0.20 21.00 0.20 0.10 73.5 18.5 2.78 18.5 0.11 21.12 432 0.20 0.20 74  11.8 0.20 21.00 8407 12.4 2.78 2.78 18.5 0.11 21.12 432 0.20 0.20 74  11.8 0.20 21.00 8407 12.4 2.78 2.78 0.11 21.12 432 0.20 0.20 74  11.8 0.20 21.00 8407 12.4 2.4 2.78 0.11 21.12 432 0.20 0.20 74  11.8 0.20 21.00 8407 12.4 2.4 2.78 0.11 21.12 432 0.20 0.20 74  11.8 0.20 21.00 8407 12.4 2.4 2.4 2.4 0.11 21.12 432 0.20 0.20 74  11.8 0.20 21.00 8407 12.4 2.4 2.4 0.11 21.12 432 0.20 0.20 74  11.8 0.20 21.00 8407 12.4 2.4 2.4 0.11 21.12 432 0.20 0.20 74  11.8 0.20 21.00 8407 12.4 2.4 2.4 0.20 0.20 74  11.8 0.20 21.00 8407 12.4 2.4 2.4 2.4 0.11 21.12 779 0.20 0.20 74  11.8 0.20 21.00 8407 12.4 2.4 2.4 2.4 0.11 21.12 779 0.20 0.20 74  11.8 0.20 21.00 8407 12.4 2.4 2.4 2.4 0.20 0.20 74  11.8 0.20 21.00 8407 12.4 2.4 2.4 2.4 0.20 0.20 74  11.8 0.20 21.00 8407 12.4 2.4 2.4 2.4 0.20 0.20 74  11.8 0.20 21.00 8407 12.4 2.4 2.4 2.4 2.4 0.20 0.20 74  11.8 0.20 21.00 8407 12.4 2.4 2.4 2.4 2.4 2.4 2.4 2.4 2.4 2.4	ė	dr/rb	8, *		In.R20	In.H20					gi 💉			n.H20	15 15 15 15 15 15 15 15 15 15 15 15 15 1	; <u>,</u>		Ę	41/10	V.	Va	o i
18.1 0.20 21.00 0.20 0.10 73 2.78 15.3 0.10 21.12 894 0.20 0.20 71.5  18.4 0.20 21.00 0.20 0.10 71 19.5 2.78 17.0 0.13 21.18 446 0.20 0.20 71.5  18.4 0.20 21.00 0.20 0.10 72.5 17.5 2.78 12.9 0.13 21.18 320 0.20 72  18.3 0.20 21.00 0.20 0.10 72.5 16 2.78 12.9 0.13 21.15 451 0.20 0.20 73  18.3 0.20 21.00 0.20 0.10 72.5 16 2.78 14.5 0.12 21.15 451 0.20 0.20 73  18.3 0.20 21.00 0.20 0.10 73 14 2.78 10.5 0.12 21.15 441 0.20 0.20 73  18.4 0.20 21.00 0.20 0.10 73 14 2.78 10.5 0.12 21.18 442 0.20 0.20 74  28.0 0.20 21.00 0.20 0.10 73 18.5 2.78 10.5 0.11 21.12 442 0.20 0.20 74  11.8 0.20 21.00 0.20 0.10 73.5 18.5 2.78 10.5 0.11 21.12 442 0.20 0.20 74  11.8 0.20 21.00 8407 1008	1		[					T				Н	H		1	1	1	7	1	+	1	
18.4 0.80 21.00 0.80 0.10 71.5 17.5 2.78 12.9 0.13 21.18 446 0.80 0.80 71.5 18.3 0.80 0.10 71.5 17.5 2.78 12.9 0.13 21.18 320 0.20 0.20 72 13.3 0.20 21.00 0.20 0.10 71.5 17.5 2.78 12.9 0.13 21.18 320 0.20 0.20 72 13.3 0.20 21.00 0.20 0.10 72.5 16 2.78 12.9 0.13 21.15 355 0.20 0.20 72 13.3 0.20 21.00 0.20 0.10 73 14 2.78 14.5 0.12 21.15 355 0.20 0.20 74 13.3 0.20 21.00 0.20 0.10 73 14 2.78 12.9 0.03 21.15 355 0.20 0.20 74 13.3 0.20 21.00 0.20 0.10 73 18.5 2.78 12.5 0.13 21.21 449 0.20 0.20 74 12.5 0.20 0.20 0.20 74 12.5 0.20 0.20 0.10 73 18.5 2.78 18.5 0.11 21.21 442 0.20 0.20 74 11.8 0.20 21.00 0.20 0.10 73 18.5 2.78 18.5 0.11 21.21 442 0.20 0.20 74 11.8 0.20 21.00 0.20 0.10 73 18.5 2.78 18.5 0.11 21.12 442 0.20 0.20 74 11.8 0.20 21.00 0.20 0.10 74 2.78 18.5 0.11 21.12 442 0.20 0.20 74 11.8 0.20 21.00 0.20 0.10 74 2.78 18.5 0.11 21.12 442 0.20 0.20 74 11.8 0.20 21.00 0.20 0.10 74 2.78 18.5 0.11 21.12 42 0.20 0.20 74 11.8 0.20 21.00 0.20 0.10 74 2.78 18.5 0.11 21.12 42 0.20 0.20 74 11.8 0.20 21.00 0.20 0.10 74 2.78 18.5 0.11 21.12 42 0.20 0.20 74 11.8 0.20 21.00 0.20 0.20 74 2.78 18.5 0.20 0.20 0.20 0.20 74 2.78 18.5 0.20 0.20 0.20 0.20 74 2.78 18.5 0.20 0.20 0.20 74 2.78 18.5 0.20 0.20 0.20 0.20 0.20 0.20 0.20 0.	1	9 9		1.		65.0	Ę		2.78		0.10	21.12	-	0.30	8.0	13	1	2.72	17	110	5 5	8
13.4 0.20 21.00 0.20 0.10 72.5 17.5 2.76 12.9 0.13 21.18 350 0.20 0.20 72  13.4 0.20 21.00 0.20 0.10 72.5 16 2.76 12.9 0.13 21.15 451 0.20 0.20 72  13.3 0.20 21.00 0.20 0.10 72.5 16 2.78 14.5 0.12 21.15 451 0.20 0.20 73  13.3 0.20 21.00 0.20 0.10 73 14 2.78 12.2 0.11 21.15 653 0.20 0.20 74  28.0 0.20 21.00 0.20 0.10 73 14 2.78 12.2 0.11 21.15 653 0.20 0.20 74  28.0 0.20 21.00 0.20 0.10 73.5 18.5 2.78 18.5 0.10 21.2 0.20 0.20 74  11.8 0.20 21.00 0.20 0.10 74 2.78 18.5 0.11 21.12 779 0.20 0.20 74  11.8 0.20 21.00 0.20 0.10 74 2.78 18.5 0.11 21.12 779 0.20 0.20 74  11.8 0.20 21.00 0.20 0.10 74 2.78 18.5 0.11 21.12 779 0.20 0.20 74  11.8 0.20 21.00 0.20 0.10 74		707	1_	30.12			,	٩	78	<u> </u>	_	21.18	1	0.50	0.20	73.5	13.5	2.72	7.7	0.13	ह	117
18.4   0.20   21.00   0.20   0.10   72.5   16   2.16   21.5   2	ا	7777		200			i i	1	8	1_	-	21.18	Į	0.30	0.20	3.5	21	2.72	5:5	9.16	21.03	FF-
13.3   0.20   21.00   0.20   0.10   72.5   16   2.78   14.5   0.12   21.15   535   0.20   0.20   73     17.1   0.20   21.00   0.20   0.10   73.5   17.5   10.6   0.12   21.18   449   0.20   0.20   74     13.3   0.20   21.00   0.20   0.10   73   14   2.78   12.2   0.11   21.15   603   0.20   74     28.0   0.20   21.00   0.20   0.10   73.5   18.5   2.78   18.5   0.11   21.15   432   0.20   74     28.4   0.20   21.00   0.20   0.10   73.5   18.5   2.78   18.5   0.11   21.12   432   0.20   0.20   74     11.8   0.20   21.00   0.20   0.10   74   2.78   18.5   0.11   21.12   432   0.20   0.20   74     11.8   0.20   21.00   0.20   0.10   74   2.78   18.5   0.11   21.12   432   0.20   0.20   74     11.8   0.20   21.00   0.20   0.10   74   2.78   0.11   21.12   432   0.20   0.20   74     11.8   0.20   21.00   0.20   0.10   74   2.78   0.11   21.12   432   0.20   0.20   74     11.8   0.20   21.00   0.20   0.30	ال	18.4	_	2 5	Q G	0:0	1	1	45	10.01	₩	21.15	-	0.20	0.50	22		2.72	1:2	91.0	27:08	0.67
15.1   0.20   21.00   0.20   0.10   73   15   2.78   10.6   0.12   21.18   448   0.20   0.20   714   13.5   0.20   21.00   0.20   0.20   714   28.0   0.20   0.10   73   14   2.78   12.2   0.11   21.15   603   0.20   714   28.0   0.20   0.10   73   18.5   2.78   17.9   0.08   21.27   447   0.20   0.20   74   22.4   0.20   0.10   0.20   0.10   73.5   18.5   2.78   18.5   0.11   21.12   432   0.20   0.20   74   22.4   0.20   21.00   0.20   0.10   74   2.78   18.5   0.11   21.12   432   0.20   0.20   74   22.4   0.20   0.20   0.20   74   2.78   0.20   0.20   0.20   74   2.78   0.20   0.20   0.20   74   2.78   0.20   0.20   0.20   74   2.78   0.20   0.20   0.20   74   2.78   0.20   0.20   0.20   0.20   74   2.78   0.20   0.20   0.20   0.20   74   2.78   0.20	, ای		8 6	8 6	8	01.0	ــــــــــــــــــــــــــــــــــــــ	١	1 %	24.5		21.15		8.0	0.50	F	2	2.72	8.7	9770	8.8	0.67
15.5   0.20   21.00   0.20   0.10   73   14   2.76   12.2   0.11   21.15   603   0.20   74     28.0   0.20   21.00   0.20   0.10   73   18.5   2.76   17.9   0.08   21.27   447   0.20   0.20   74     22.4   0.20   21.00   0.20   0.10   73.5   18.5   2.76   18.5   0.11   21.21   432   0.20   0.20   74     11.8   0.20   21.00   0.20   0.10   74   2.76   2.76   0.11   21.12   759   0.20   0.20   74     11.8   0.20   21.00   0.20   0.10   74   2.76   0.11   21.12   759   0.20   0.20   74     11.8   0.20   21.00   0.20   0.20   0.20   0.20   74     11.8   0.20   21.00   0.20   0.20   0.20   0.20   0.20   0.20     11.8   0.20   21.00   0.20   0.20   0.20   0.20   0.20   0.20     11.8   0.20   21.00   0.20   0.20   0.20   0.20   0.20   0.20     11.8   0.20   21.00   0.20   0.20   0.20   0.20   0.20     11.8   0.20   21.00   0.20   0.20   0.20   0.20   0.20   0.20     11.8   0.20   21.00   0.20   0.20   0.20   0.20   0.20   0.20     11.8   0.20   21.00   0.20   0.20   0.20   0.20   0.20   0.20     11.8   0.20   21.00   0.20   0.20   0.20   0.20   0.20   0.20     11.8   0.20   21.00   0.20   0.20   0.20   0.20   0.20   0.20     11.8   0.20   21.00   0.20   0.20   0.20   0.20   0.20   0.20     11.8   0.20   21.00   0.20   0.20   0.20   0.20   0.20   0.20     11.8   0.20   21.00   0.20   0.20   0.20   0.20   0.20   0.20     11.8   0.20   21.00   0.20   0.20   0.20   0.20   0.20   0.20     11.8   0.20   21.00   0.20   0.20   0.20   0.20   0.20   0.20     11.8   0.20   21.00   0.20   0.20   0.20   0.20   0.20   0.20     11.8   0.20   21.00   0.20   0.20   0.20   0.20   0.20   0.20     11.8   0.20   21.00   0.20   0.20   0.20   0.20   0.20   0.20     11.8   0.20   21.00   0.20   0.20   0.20   0.20   0.20   0.20   0.20   0.20     11.8   0.20   21.00   0.20   0.20   0.20   0.20   0.20   0.20   0.20   0.20   0.20   0.20     11.8   0.20   21.00   0.20   0.	١,	1 1	4-	_ـــــ		9 0	1_	٢	2.78			21.18	-	02.0	R o	73.5	38	27.2	777	27.0	21.00 0.50	92.5
28.0 0.20 21.00 0.20 0.10 73 18.5 2.76 17.9 0.08 21.27 44.7 0.20 0.20 74 22.4 0.20 21.00 0.20 0.10 73-5 18.5 2.78 18.5 0.11 21.21 472 0.20 0.20 74 11.8 0.20 21.00 0.20 0.10 74 2.78 18.5 0.11 21.12 779 0.20 0.20 74 11.8 0.20 21.00 880 0.10 74 2.78 2.78 0.11 21.12 779 0.20 0.20 74.5 11.8 0.20 21.00 880 0.10 74 2.78 7.78 0.11 21.12 779 0.20 0.20 74.5	١,	7	4_	4-	8	9.50	1	7.	2.78			21.15	-	8.0	0.3	72	۵	2.72	8.7	0.15	2	200
25.0 0.00 21.00 0.20 0.10 73.5 18.5 2.78 18.5 0.11 21.2 4.20 0.20 0.20 74. 22.4 0.20 21.00 0.20 0.10 74. 2.78 0.20 0.20 74.5 11.8 0.20 21.00 0.20 74.5 0.10 74.5 0.20 0.20 74.5 11.8 0.20 21.00 880 0.20 0.20 74.5 11.8 0.20 21.00 880 0.20 74.5 11.8 0.20 21.00 880 0.20 74.5 11.8 0.20 21.00 880 0.20 74.5 11.8 0.20 21.00 880 0.20 74.5 11.8 0.20 21.00 880 0.20 74.5 11.8 0.20 21.00 880 0.20 74.5 11.8 0.20 21.00 880 0.20 74.5 11.8 0.20 21.00 880 0.20 74.5 11.8 0.20 21.00 880 0.20 74.5 11.8 0.20 21.00 880 0.20 74.5 11.8 0.20 21.00 880 0.20 74.5 11.8 0.20 21.00 880 0.20 74.5 11.8 0.20 21.00 880 0.20 21.00 880 0.20 21.00 880 0.20 21.00 880 0.20 21.00 880 0.20 21.00 880 0.20 21.00 880 0.20 21.00 880 0.20 21.00 880 0.20 21.00 880 0.20 21.00 880 0.20 21.00 880 0.20 21.00 2	۱,		4-	+-	8	6	1_	18.5	2.78	17.9	_	22.27	7,44.	0.20	0.30	#	2	2.72	27.7	0.13	2	2
11.8 0.20 21.00 03.0 43.0 73. 2.78 0.11 21.12 .759 0.20 0.20 744.5	1	3	٠.	+-	1	1	┺-	Į.	87.6		_	21.21	-	02.0	0.20	7	9	2.72	17.0	21.0	27.79	0:45
11.8 0.20 21.00 0.30 74 2.75 11.8 0.20 21.00 8407 12.8		22		-	3		ᅩ	L		-	┺-	91 10		06.0	0.30	74.5	7	2.72		0.15	21.00	83
11.8 0.20 21.00 SHITT	g	11.8	_	-	8	9	_L		2	T	1	3	+-				Γ					
		11.8				NONCE				1	T											
		_	_							1	1	1					T					
		_	L	L	_												1				1	T
			L	L																	T	
						_																
	1	-	L			_	L															
	١	1	1	1		1																
	1	$\downarrow$	$\downarrow$	1		1	-															
			_																		! 	İ

 $P_{\rm h}$  = Static pressure just upstress of canister  $r_{\rm h}$  = Temperature just upstress of earlster

 $\Delta P_{\rm B}$  = Pressure drop across canlater  $\Delta r_{\rm B}$  = Temperature rise across cantater

Table 10 (cont.) Partially Reduced Data, P-6

		.o.		0.63	87	87.	0.57	19.0	19.0	0.61	o.53	74.0	29.0		Γ	Γ		Γ		Ī	Γ
	Г		$\vdash$	21.27 0	<del> </del>	-	⊢	-		<b>-</b>	⊢	-	}—	$\vdash$	H	├	+	╀	+	╀	$\vdash$
	AT A	o°v.	L	Z	21.27	21.24	12.23	21.18	21.18	21.18	21.33	21.30	21.12	L	L		L	L	L	L	L
	OUTLET A	8 ×		0.03	90.0	90.0	90.08	60.0	90.0	60.0	0.03	90.0	0.12								ļ
	8	R20 Gr/Lb		12.9	16.4	15.4	12.9	14.5	12.2	9.6	18.5	18.5	8.3					Γ			
MSA		MO H		1.93	1.93	1-93	1.93	1.93	1.93	1.93	1.93	1.93	1.93								
CANESTER		ΔT. ₹			54	21		16	13	11	22.5	21	וו								
CAN		Et e	Γ	77	76.5	92	76.5	77	77.5	77.5	78	78	77.5								
		In.H20		0.70	0.75	0.75	0.75	0.75	0.73	0.73	0.75	0.75	0.75		-		-				
		rate ogh.ni		0.75	0.75	0.75	0.75	0.77	0.78	22.0	8.80	98.0	0.78								
		7. Q.		1.00	29.0	0.50	0.60	9.0	0.75	0.92	29.0	0.61	0.67								
		o, 8		21.19	21.18	21.18	21.19	21.15	21.12	21.12	21.18	21.18	21.12								
	OUTLET AIR	8° <b>x</b>		50.0	90.0	11.0	0.11	11.0	0.11	0.0	0.08	60.0	0.12		-						
	B	HO Gr/LD		13.7	15.4	12.9	12.9	14.5	14.5	12.2	23.3	19.5	30.6								
A-5	HOF	FLOW		2.74	₽.2.	2.74	2.74	2.74	2.74	2.74	2.74	2.74	2.74	_							
CANISTER	(no catalyst) + 1/3 Screen	•F			21.5	23		នុ	3.7	16.5	18.5	19	97								
8	talyst, 1	Ę,		4	22	22	72.5	2	73.5	7,	7	74.5	75.0								
	4/3 KO2 (no ca + Ann. Screen	48.60 Tn.H20		0.25	0.25	0.25	0.25	0.25	0.28	0.28	0.28	0.28	0.28	DOWN							
	2/3 KO2 + Ann.	F <sub>I+</sub> In.H <sub>2</sub> 0 In.		0.35	0.35	0.36	0.36	0.35	0.35	0.38	0.38	0,0	0.38	SHOT							
		St PS	21.03	21.00	21.00	21.00	21.00	21.00	21.00	21.00	21.00	21.00	21.00	21.00							$\neg$
DALET ATR		202	0.19	0.50	0.20	0.20	0.20	02.0	0.20	0.20	0.20	0.20	0.50	02.0							٦
I KI		H20 Gr/Lb.	18.1	18.1	7.71	18.4	13.3	17.0	13.3	15.5	28.0	22.4	11.8	11.8							
	TO S	M.n.	٥	13	6	93	121	156	216	276	336	386	456	461							

 $P_{\rm D}$  = Static pressure just upstream of canister  $T_{\rm D}$  = Temperature just upstream of canister

 $\Delta P_{\rm R}$  = Pressure drop across canister  $\Delta T_{\rm R}$  = Temperature rise across canister

Partially Reduced Data, P-7 Table 11

Date: 1-19-62

		T .	T	Т	Τ	Т	1	T	Т	7	Т	Т	ĭ-	7	7	7	T	$\overline{}$	Т	Т	7
	_	G; G;		8	0.17	0	8	0.33	0.26	6.3	0.3	72.0	5	3	L		L		L	$\perp$	1
	AIR	8,4		23.27	21.27	8.	200	21.27	21.27	6	15	21.15	5	2							
	OUTLET A	1 4	Π	90.0	$\mathbf{T}$	7	T			$T^{T}$			1	1						Ī	T
	B	42/26 Gr/Lb	T	34.8	40.2	23.3	66.2	63.8	55.4	67.0	1.4	69.2	12.2					Ī			T
¥-6		FLOW		2.69	2.71	2.2	2.71	2.71	2.71	2.71	2.71	2.81	<del> </del>	•				T			1
1_		ong.		21.5	38	10.5			37.5	37.0	35.5	33.0							Ī	T	
	· 0. ™ +	<b>♦</b> 8.		73.5	ř.	92	76.5	76.5	14	77	16	75	92					T			<del> </del>
	Sereen	Zu.Rgo		0.12	0.12	0.12		0.12	0.12	0.12	0.12	0.12	0.12						T		T
3	4 Ann. Screen	In.H20	1	8.0	0.20	0.30	0.20	0.20	0.20	0.00	0.20	0.20	0.20							T	T
Γ		3.C	T	0.28	0.22	0.22	0.24	0.23	0.22	0,23	0,33	0.22	0.22		_	-		-		t	T
		0, 8	T	21.24	21,39	21.39	21,36	21.36	21.36	21.24	21.18	21.18	21.18	_	-	-	-	-	r	T	╁
	OUTLET AIR	8 ×		6.0	80.0	60.0	80.0	80	60.0	60.0	60.0	51.0	0.12	_		-		T			T
	b	41/10 H20		42.9	51.5	54.7	59.0	9.09	63.8	62.0	61.9	9. 29	73.4		-	-	-	-	-		T
		MOLA		2.77	2.79	2.81	2.81	2.81	2.81	2.81	2.81	2,88	2.81					-	r		T
CANTSTER	u e	AT,		24.5	48.5	45.5	0.44	0.54	41.5	38.5	32.0	36.0	37.5				-	-			
S.	KO2 + Cat. + Ann. Screen	ĕĭ•		7.	7.14	75	75.5	92	76.5	76.5	26	74.5	75.5						l	r	
	at. + A	AP.		0.15	0.15	21.0	0.15	0.14	0.15	0:15	0.15	310	0.15	DOSM							
	x02 + c	Po APil In.H20 In.H20		0.22	0.22	0.25	0.25	0.26	%	9.50	0.26	92.0	9.50	SHITT		-			-		
		8,8	20.79	8 8.8	20.85	88.88	20.85	88.02	20.85	20.85	20.85 0.26	38.05	20.82	28.82					-	-	
INLET AIR		હું *	0.34	0.20	0.30	9.18	0.50	81.0	0.20	0.20	0.30	02.0	0.50	QZ O	7	1				-	
INT		H20 Gr/Lb.		#. #.	93.4	2.38	1.65	1.8	99.1	33:1	78.8	108.1	108.1	108.1	1	1			_	-	
	andr Sport	MEn.	٥	1,5	45	22	128	527	8 8	윉	30	360	02.7	435							

 $P_{\rm R}$  = Static pressure dust upstream of canistar  $\gamma_{\rm R}$  = Temperature just upstream of canister

 $\Delta P_{\rm B}$  . Pressure drop across canister  $\Delta T_{\rm B}$  . Temperature rise across canister

Table 11 (cont.) Partially Reduced Data, P-7

		9.		46.0	0.22	02.0	0.25	0.24	0.25	0.33	44.0	0.37	0.33					
	Ĺ	o°v.	Γ	21.48	21.57	21.57	21.48	24-12	21.36	21.24	21.12	21.12	21.09	Γ				
	OUTLET AIR	8,*		0.05	0.05	40.0	0.04	0.05	2 20.0	0.07	0.08	60.0	11.0	 Γ				
	ğ	H20 Gr/Lb		29.5	39.2	7.94	53.9	57.0	63.4	68.3	64.3	2.77	27.77					
4		MOTA GUI	Ī	2.13	2,13	2.13	2.16	2.16	2,14	2.13	$\overline{}$	39.1	2,05					
CANISTER MSA		4T.	Γ	34.0	54.0	54.0	50.5	46.5	41.0	33.5	27.0	23.0	20.05			_		
CAN		O.B.		76.5	77.5	78.0	78.5	78.5	79	79.5	62	78.5	78.5					
		APIT In.H20		0.46	94.0	0.46	0.50	0.51	0.52	0.68	0.70	0.90	1.70			_		
		P. APH In.H20 In.H20	Γ	0.53	0.55	0.55	0.56	0.59	19.0	0.75	0.75	0.95	27.1					
		я. ф.	Γ	0.22	0.25	71.0	92.0	0.23	0.28	0.27	94.0	0.27	0,33					
	[	9, x		21.45	21.33	21.30	21.27	23.27	23.24	21,18	21,119	21.12	23.09				<u> </u>	
	OUTLET AIR	8,1%		0.07	80.0	11.0	0.0	60.0	60.0	لنده	60.0	0.12	11.0					
	8	dr/rb		8. 6.	₽.09	65.0	67.5	5.69	70.8	72.4	57.2	71.5	75.6			_		
A-7		MTON cfm		2.71	2.73	2.74	2.85	2,82	2.82	2.74	2.88	2.88	88.5		_			
CANISTER A-7		ΔT3 *P		20.5	31.5	35.0	35.5	35.0	33.0	31.5	30.5	28.5	29.5					
- CO		t- fk, E10		73.5	75.5	76.5	77	76.5	77	77	76	22	76.5					
2/3 KO. + Cat. +	+ Ann. Screen	O In.H20		0.14	6,14	41.0	0.14	9.74	0.14	0.14	91.0	0.15	0.15					
2/3 KG	+ Ann	P;4 In.H20		0.20	0.23	0.23	0.23	0.23	0.23	0.25	0.25	0.25	92.0					
		8,8	20.79	20.85	20.85	20.88	20.85	20.88	20.85	20.85	20.85	20.82	20.82					
INLET AIR		S) 86	7.34	0.20	0.20	81.0	0.20	0.18	0.20	0.20	02.0	02.0	0.20					
		<sup>Н20</sup> Gr∕Ľb.		4.08	93.4	2,98	99.1	1.66	99.1	99.1	78.8	108.1	108.1					
	TOGE	Man.	٥	R	9	&	125	155	185	245	305	365	425					

 $P_{\rm R}$  = Static pressure just up, tream of canister  $T_{\rm D}$  = Temperature just upstream of canister

 $\Delta P_{\rm B}$  = Pressure drop across canister  $\Delta r_{\rm B}$  = Temperature rise across canister

Table 12 Partially Reduced Data, P-8

-11-62
Dete:
P-8
Jest T

			Т	T	т-	Т	Τ-		Т	1		_		т	Т	1	1	T	Т	Т	T
ļ	_	R. 0.		8	8	0.67	9	84.0	5.5	0.47	0.53	8	0.47	0.67	84.0	0.75			L		
	E E	8,*		20.76	20.76	02.02	8.8	60	5	90.13	90.19	70.09	20.97	20.85	20.91	80.82					
	OUTLET ALR	g,∗v		0.30	₹	0.45	1	,	1	Ι	0.32	Г	į	14.0	0.41	0.42					Ī
	g	H20 Gr/Lb		7:7	6.6	6.6	Г	_	21.3	Ι	23.2	18.	13.2	12.5	12.2	6.6					T
4-6		WILCW of F		2.76	2.76	2.76	2.76	_		2.76	2.76	_	_	2.76	2.76	2.76					T
		ΔT2		14.5	15.0	14.0	15	22.5	-		33	34.5	27	19.5	16.5	14.5			T	T	T
CANCE +	een -	Ft.	T	73.5	72.5	73.5	47	7.4	74	17.2	74	74	74	74	7.4	7.5				T	
CANCESTER 2/3 KO. + Cat. + 1/3 M. S.	+ Annular Screen	APIS In.H20	Γ	0.8	0.20	0.20	0.20	0.20	0.20	0.20	0.20	0.20	0.20	0.20	0.20	0.20					T
2/3 KO	Tage of	Pig OFI. n. II.		0.22	0.50	0.20	0.20	0.20	0.20	0.20	02.2	0.20	0.20	0.50	0.20	0.20					
		R.Q.		0.78	59.0	0.51	94.0	0.47	0.53	0.57	0.50	0.47	0.63	79.0	0.62	79.0		ĺ			T
		o <sup>™</sup> ×		20.88	20.97	21.09	21.18	22.12	21,15	أعدره	21.12	130,15	20.97	20.9¢	20.91	20.88					-
	OUTLET AIR	8,*		0.37	0.34	0.31	0.28	9.54	0.27	0.27	0.30	0.34	0.34	0.35	0.38	0.39	-	-		T	
	8	HO Gr/720		8.1	20.5	21.3	22.4	25.5	27.0	27.0	27.0	24.0	16.9	17.4	14.8	13.8					
'		HO'LA		2.93	2.73	2.73	2.73	2.57	2.57	2.57	2.57	2.57	2.57	2.57	2.57	2.57					
CANESTER	reen	A.		23	27	32	35	39.5	39.5	35.5	34.0	32.5	28.5	25.5	23.0	20.5					
8	Arn. Se	Eigh		73.5	72.5	73	7,7	73.5	5	73.5	73.5	73.5	73.5	73.5	7.0	7.4					
	KO <sub>2</sub> + Cat. + Ann. Soreen	A.H.		0.13	0.13	6.13	0.13	0.13	0.13	0.13	0.13	0.13	0.13	0.13	0.13	0.13	NAOG				
	+ 20x	Po ARI In.H20 In.H20		0.23	0.23	0.23	0.23	0.23	0.23	0.23	0.23	0.23	0.23	0.23	0.23	0.23	SHUT				
		& જ	20.67	20.70	20.70	20.70	8.70	20.70	20.70	8 2.2	20.70 0.23	20.70	20.70	20.70	20.70	20.70	20.70				П
INCET AIR	Ì	8 <b>*</b>	64.0	6.51	0.53	0.51	S.	0.51	6.51	0.51	6.51	0.51	٠. تر	0.52	0.51	0.51	0.53				
E		H20 Gr/Lb.	10.8	10.8	15.5	19.3	19.3	35.2	35.2	35.3	32.1	26.4	20.0	51.6	19.3	13.9	13.9				П
	108	M.n.	٥	F	£13	23	ខ្ម	148	183	208	238	888	298	328	358	388	408				

 $P_{\rm h}$  = Static pressure just upstream of canister  $\eta_{\rm h}$  = Temperature just upstream of canister

 $\Delta P_{\rm R}$  = Pressure drop across canister  $\Delta T_{\rm R}$  = Temperature rise across canister

Table 12 (cont.) Partially Reduced Data, P-8

1		α. .0.		0.60	3.0	0.30	0.45	9	6			l		L		ļ	
	Ħ	o,*		8	20.79	8	21.15										
	OUTLET AIR	°, ×		9		1	Т										
	g	R20 Gr∕∴b		9	6.6	27.0	7.8	5. 8.	55.6				T				
A-6		FLOW afte		3	2.87	2.87	2.87	2.8	2,87		Γ						
CANESTER	3.5	ST2		7.	16	21.5	40.5	#3	30								
AS.	} + • ₽	₽¢ Fi		٤	2	72	72.5	73.5	74.5								
1	Soree	AP13 In.H20		8	0.20	0.20	0.20	0.20	0.20								
2/3	+ Am	Pas O In. H20		0.20	0.25	0.22	0.22	0.25	0.23								
Г		R.Q.		0.30		79.0	0.41	0.61	0,69								
	AIR	SN NS		2,00	20.85	20,88	21.15	20.97	20.04								
	OUTLET A	8, 86 8		0.38	0.34	0.35	0.30	0.31	0.33								
	ō	H20 Gr/12b		17.71	17.0	29.5	51.5	50.3	58.8								
A A		FLOW of m		2,89	2.89	2.89	2.89	2.89	2.89								
CANESTER	reen	Ą.		18	g	23.5	35	92	32.5								
5	XO2 + Cat. + Ann. Screen	E-i €		70.5	69.5	71.5	72	22	74								
	cat.	AP, In.H20		0.13	0.15	613	0.13	0.15	0.15								
	+ 0x	Fig AFii		0.23	0.25	20.64 0.23	20.64 0.23	0.23	0.23								
		જ ૪૯	20.73	20.67	20.67	20.64	20.64	20.64	20.64								
INLET AIR		g №	0.51	0.51	0.51	0.51	۰ ت	0.51	0.51	NIMOCI							
Ä		H2O Gr/Lb.	23.2	23.2	21.6	29.6	68 8	73.6	73.9	SHUT							
	TOGE	K.n.	408	423	468	533	593	653	713	728							

 $P_{\rm R}$  = Static pressure just upstream of canister  $T_{\rm R}$  = Temperature just upstream of canister

 $\Delta P_{\rm B}$  = Pressure drop across canister  $\Delta T_{\rm B}$  = Temperature rise across canister

Table 12 (cont.) Partially Reduced Data, P-8

	_	8.0		92.0	0.72	69.63	0.52	99	53	82.0	9.0	0.67	0.67	8	8	6	3				
	Fi	8°4		21.06	21.12	21.18	21.45	21.30	21.21	21.18	21.00	\$. \$.	88	85.05	20.82	8	3				
	OUTLET AIR			6.3	0.21	0.21	0.12	0.16		_			1 -	1	ł	ı	1	T	T	T	
	g	H20 Gr/Lb		15.3	13.2	15.3	28.6	27.1	28.6	27.1	27.1	21.3	15.4	14.5	12.2	3	7.7		Ī		
MSA		FLON FLON		2.14	2.14	2.14	2.14	2.14	2.14	2.14	2.14	2.14	2.24	2.14	2.14	- 0			Ī		
CANESTER		ÅT4.	Г	34	%	30.5	41.5	45	39.5	32.5	32.5	26.5	17.5	13.5	13.0	1.0					T
CAN		ŠP!		78	22	7.8	78.5	78	78	27.5	77.5	78	77.5	78	77.5	77. 5					
		O In Bo		0.55	0.55	0.55	0.63	0.63	0.64	0.64	0.64	0.64	0,64	0.64	0.64	0.64					
		Pie In.H20		0.63	0.63	0.65	0.68	0.70	0.70	0.70	0.70	02.0	02.0	0.70	0.70	0.70					
Γ		R.Q.		0.39	0.77	0.63	0.57	0.64	0.64	0.64	0.67	79.0	0.52	0.87	0.80	0.83					Γ
		0, N		21.24	21.00	21.00	21.00	21.03	21.00	21.00	21.00	20.02	20.97	20.85	20.85	20.82					
	OUTLET AIR	8 % %		0.30	0.28	0.32	0.34	0.30	0.32	0.32	0.31	0.35	0.37	0.38	0.39	0.41				]	
	õ	41/20 о'н		19.4	24.4	22.9	25.7	28.7	33.4	32.0	32.0	27.3	21.6	20.5	17.9	17.0					
7-4		MOTAL		2.74	2.74	2.74	2.74	2.86	2.86	2,86	2.86	2.86	2.86	2,86	2.86	2.86					
84		era P		†Z	30	30	29.5	31	30.5	28.5	28	26.5	23	23	ୡ	18.5					
\$\\\\\\\\\\\\\\\\\\\\\\\\\\\\\\\\\\\\\		Æ.		73	73.5	7.4	74.5	74	7.9	7.4	74	7.4	74	7,7	4.7	74					
) + Cat	+ Ann. Soreen	F <sub>14</sub> AP <sub>18</sub> In.H <sub>2</sub> 0 In.H <sub>2</sub> 0		0.10	0.10	01.0	0.10	0.11	0.11	0.11	110	17.0	0.11	11.0	0.11	נניס	DOWN				
2/3 KG	+ Vuu	F14 In.H20		8.0	0.20	0.20	8.0	0.20	0.20	0.20	0.20	0.30	0.20	0.20	8.0	0.20	SHUT				
		8,8	20.67	20.70	20.70	20.70	20.70	20.70	20.70	20.70	20.70	20.70	20.70	20.70	20.70	20.70	20.70				
DILET AIR		% №	64.0	0.51	0.51	0.51	0°.50	0.51	0.51	0.51	0.51	6.53	0.51	0.51	0.51	15.0	0.51				
Ä		H2O Gr/Lb.	10.8	10.8	15.5	19.3	19.3	35.2	35.2	35.3	32.1	26.4	20.0	21.6	19.3	13.9	13.9				
	TOCK	M.n.	0	23	23	g	113	153	980	213	243	E S	333	E	193	88	80				

 $P_{\rm B}$  = Static pressure just upstream of canister  $T_{\rm B}$  = Temperature just upstream of canister

 $\Delta P_{\rm B}$  = Pressure drop across canister  $\Delta \Gamma_{\rm B}$  = Temperature rise across canister

Table 12 (cont.) Partially Reduced Data, P-8

Date: 1-12-62

Test P-8 (cont.)

0.75 0.70 0.69 0.30 20.94 2.15 16.1 0.39 20.82 2.15 16.1 0.38 20.79 2.15 23.7 0.33 20.88 37.5 2.15 48.7 0.26 21.00 AIR 8, 8 OUTLET 53.1 59.4 2.15 16.5 2,15 29.0 2.15 FLOW CANISTER MSA 0.61 20.5 80.5 410 77.5 74.5 Fig. 15 0.60 78 0.60 79 0.58 0.62 In.H20 In.H20 0.58 99.0 99.0 99.0 58.3 0.32 20.94 0.63 0.66 29.5 2.75 65.2 0.34 20.88 0.71 0.66 20.0 2.75 21.7 0.37 20.79 1.17 21.5 2.75 33.9 0.35 20.82 0.89 30.0 2.75 59.5 0.31 21.00 0.56 16.5 2.75 18.4 0.41 20.91 0.42 9 9, K OUTLET AIR 8, × Gr/Lb 2.75 FLOW 2/3 KO<sub>2</sub> + Cat. + 1/3 L4OH + Ann. Screen 31.5 AT3 5 75 72 8 12 F Ľ, 0.11 0.11 0.11 P.4 OP.8 0.11 0.11 DOWN 20.67 0.15 20.67 0.15 0.51 20.64 0.15 0.51 20.64 0.15 20.64 0.15 SHUT 20.73 g, re . 약 약 0.51 0.51 INCET AIR 0.51 8, 4 H20 Gr/Lb. 23.5 3.6 5473 543 598 658 718 433 TIME 802

R.Q.

 $P_{\rm R}$  = Static pressure just upstream of canister  $T_{\rm R}$  = Temperature just upstream of canister

 $\Delta T_{\rm n}$  = Temperature rise across canister  $\Delta P_{\rm B}$  = Pressure drop across canister

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Table 13	

				L		CA	ANISTER								CANT	CANISTER A-6	و				
TDE	1	LALET ALR		KO2 +	KO <sub>2</sub> + Cat. + Ann. Screen	Ann. Se	reen		8	OUTLET AIR			2/3 KO + Ann.	2/3 KO <sub>2</sub> + Cat. + Ann. Screen	+ 1/3 M.S.	M.S.	·	9	OUTLET ALR		
Æn.	H20 Gr/Lb.	% 202	8,0	Fio In.H20 In.	Tn.R20	F10 -Br	7.4.	MOTE SUP	H20 Gr/Tb	8 %	0, 18	o,	In.H20	In.H20	46	P.	FLOW of the	H20 Gr/Lb	8,46	S.A	9.
٥	15.3	0.23	20,85								-	Γ									Ì
15	75.3	0.23	20.85	0.55	0.33	72	24.5	4.50	1.94	60.0	21.12	0.52	0.60	0.33	72	1,1	4.65	38.4	80.0	21.09	0.02
45	57.3	0.20	20.88	0.55	0.35	7.1	29.5	4.60	0.84	91.0	21.12	0.17	0.50	0.33	74.5	38	4,62	26.5	0.12	21,15	0.30
22	55.4	0.20	20.88	9.60	0.35	92	F	4.60	51.5	51.0	21.15	0.26	0.50	0.33	92	38.5	4.66	34.0	0.12	21.18	0.27
115	0.79	0.20	20.88	0.60	0.35	76.5	30.5	4.60	49.9	51.0	21.15	0.26	0.50	0,33	1	35.5	4,66	35.6	0.09	21.15	0.41
145	57.3	0.20	20.88	0.60	0.35	92	28.5	4.60	51.5	6.13	21.15	0.26	95.0	0.33	76.5	33.5	4,66	43.9	0.11	135	0.33
175	1.22	0.20	20.88	0.60	0.35	76.5	29.5	9.4	51.5	0.13	21.12	0.29	0.50	0.33	77	31	4.64	45.5	0.11	21.12	0.37
802	95.58	0.20	20.88	9.60	0.35	76.5	28.5	4.60	62.0	0.16	21.15	0.15	0.50	0.33	11	28.5	4.64	50.09	91.0	21.19	0.15
240	87.9	0.20	20.88	9.0	0.35	78	82	4.61	21.5	13	21.15	0.26	0.50	0.33	78.5	30.5	4.65	65.9	0.14	21.15	0.22
592	54.1	0.20	20.88	0.60	0.35	77.5	70	4.78	51.2	51.0	901.2	0.39	9.0	0.33	77.5	22.5	4.85	46.4	0.17	21.06	0.17
310	77.1	0.21	20.88	0.60	0.35	76.5	23.5	4.66	57.9	91.0	30.15	22.0	05.0	0.33	77	23	4.76	56.7	0.16	60.12	61.0
335		0.21	20,88	TOHS	NMOC						1	٦									
355	72.1	0.20	20.88	0.55	0.38	81.5	9.5	4.62	5.54	21.0	20.97	0.33	84.0	0.33	18	12.5	4.63	42.9	91.0	90.1	0.22
385	67.0	0.24	20.82	0.58	0.38	72	18.5	4.61	59.0	0.17	20.97	0.78	0.50	0.35	72.5	18.5	4.64	54.2	0.17	46.05	8.5
445	73.6	0.20	28.85	0.57	0.38	13	18.5	4.61	4.29	0.16	20.97	0.33	0.50	0.35	73.5	17.5	4.59	59.1	0.16	g. g.	44.0
505	77.1	02.0	20.85	0.57	0.38	2	&	4.61	65.8	0.14	21.00	0.40	0.52	0.35	73.5	18.5	4.64	62.7	0.14	21.00	0,40
565	82.5	0.20	20.85	0.57	0.38	7.	83	4.65	67.9	1	21.00	٦	25.5	9.36	74.5	21.5	4.69	8.99		00.19	
525	77.1	0.20	20.85	0.57	0.38	73	19.5	4.65	65.9		21,00		0.52	0.38	73.5	18.5	4.64	62.7		00-14	
650		02.0	20,85	SHUT	DOWN																

 $P_{\rm B}$  = Static pressure just upstream of canistor  $r_{\rm B}$  = Temperature just upstream of carister

Table 13 (cont.) Partially Reduced Data, P-9

1-16-62
Date:
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7-			1	_	_	_	•	_	~	1	Ţ	7	T	_	1	7	_	T	Τ~	$\overline{}$	_
		R.O.		o.51	0.33	0.27	0.43	0.41	0.37	0.37	0.43	000	96.0		0.33	92.0	8	2 0			
		%		21.24	21.21	21.21	21.18	21.15	21.12	21.18	21.09	90.19	2	Γ	5	500	Γ	$I^{-}$			
	OUTLET AIR	8, ×		0.03	60.0	0.11	0.07		Т	1		Г	т-	Г	0.10		1	1			
	б	41/тр Gr/Лъ		32.4	45.5	41.9	46.7	49.9	-	<del>  -</del>	73.1	_	1	1-	47.8	-	+-	-			T
	_	FLOW cfn		3.59	3.60	3.60	3.61	8.5		┢	8.9	┿	<b>├</b> ─	┪	3.50	-	┰	Т	1	T	
CANISTER MSA		47. 4		杰	39	35.5	30.5		$\vdash$		59	T	Г	Γ	6.5	1-	⇈		Γ		
CANES		64.0 FF		76.5	78	79	79.5	_	L	2	8	8.5			82.5	44	Г		ŝ		
		Ogu.uI		0.92	0.95	1.00	1.10	1.20	1.40	1.50	1.80	2.20	-		5.10	5.57	┼	Η-		1	
		I ogu-uz		71.1	1.17	1.22	1.30	1.40	Н	1.65	1.98	2.35	3.60		5.35	5.75	┝	$\vdash$	۳	1	
		R.Q.		0.35	0.37	0.24	0.43	0.41	0.37	n.33	0.52	0.44	0.39		0.25	0.53	0.78	_	_	_	
	_	o'w		21.33	21.18	12.12	21.18	21.15	21.12	21.15	21,09	21.06	21.06		21.12	20.97			20,97	20.97	
	OUTLET AIR	202 ×		90.0	60.0	0.12	0.07	0.0	0.11	0.11	60.0	0.12	0.13		41.0	91.0	Н	0.13	-		П
	Ф	H20 Gr/Lb		8, 74	50.3	43.5	45.1	51.5	55.6	66.7	8.02	49.9	62.0		43.0	55.9	61.1	61.1	75.5	69.5	
A-7		FLOW cfm		4.59	4.59	4.64	4.64	4.64	179.1	19.4	4.64	4.83	4.83		4.60	4.56	4.69	4.69	4.57	4.57	
<b> </b>		ΔT3 °F		27.5	31	32	30.5	59	30	30	31.5	27	%		77	54	22.5	22.0	23.5	19.5	
+ 1,5g	2	Fig.		22	75.5	76.5	12	77	44	77	78.5	7.2	77.5		8	72.5	73.5	75. L	74.5	74	
CANISTE 2/3 KO, + Cat. + 1/3 L10H	Screen	APis In.R20		0.25	0.25	0.25	0.25	0.25	0.25	0.25	0.25	0.26	0.27	DOWN	0.28	0.28	0.30	0.30	0.28	0.28	DOWN
2/3 KO,	+ Ann.	Fi4 In.H20 In		0.41	0.43	0.43	6.43	0.43	0.43	0.43	0.43	0.45	0.45	SHUT	24.0	0.46	94.0	84.0	94.0	94.0	SHUT
		8,8	20.85	20.85	20.88	20.88	20.88	20.88	20.88	20.88	20.88	20.88	20.88	20.88	20.88	20.82	20.85	20.85	20.85	20.85	20.85
INLET ALR		S, №	0.23	0.23	0.30	0.20	0.30	0.20	0.50	0.20	0.20	0.20	0.21	0.21	0.20	0.24	0.20	0.20	0.20	0.20	0.20
INI		H20 Gr/Lb.	75.3	75.3	57.3	65.4	67.0	57.3	62.1	82.6	87.9	54.1	77.1		54.7	67.0	73.6	77.1	82.5	77.1	77.1
	ECT.	M.n.	٥	g	55	8	115	150	180	210	240	270	315	335	360	395	455	515	575	625	650

 $P_{\rm n}$  = Static pressure just upstream of canister  $\tau_{\rm n}$  = Temperature just upstream of canister

 $\Delta P_{\rm R_1}$  = Pressure drop across canister  $\Delta R_{\rm R_2}$  = Temperature rise across canister

Partially Reduced Data, P-10 Table 14

	_	o, ne	1	] 3	3 8	3 k	8	3	8	8	3 8	ν κ 2 ς	3 8				8	핅	잃	0.28	33	1
	R	જ,≉		1	7 2	21.54	21.48	24 16	2	21 45	1	1		4	9	- 1	-1	- }	27.27	21.21	81.19	1
	OUTLET AIR	ઈ ₩	T	[	1	1	Т	Т	$\top$	5	Т	Т	7-	7-	Т	Т	Т	Т	6.0	0.07	20.0	١
	g	H20 97/19	T	ā	30.5	22.2	31.8	36.5		38.1	, 4	9	1 6	2 6	ا ا	┿	╅	12:1	32.0	43.2	43.0	1
8-8		MOTAL CLB	T	á	8.8	2.8	8.8	2.78		2.78	92.0	2.78	92.0	9 0	3 6	┰	+	_	2:75	2.75	2.75	т
STER		ÅT2		1 1	ŧ	51.5	23	5	1	18.5	16.5	5.5	11	1	,	7		200	× ×	S	35.5	ľ
CAN		44	T	12	t,	73.5	73.5	Ę		ř.	1,5	£	Ł	٤	٤	9 ;	2 1	2	8	2	70	1
KO <sub>2</sub> + Cat., in 20" Canister	+ Ann. Screen	In.H20 In.H20		8.5	0.18	0.18	0.18	81.0	9,18	0.18	0.18	0.18	81.0	80.0	9		9 9	9	819	87.0	0.20	ľ
KO <sub>2</sub> +	+ Ann.	Fig.	T	8	0.26	0.26	0.26	0.26	%	0.26	0.26	9.50	0.26	9.0	y,	3	8	8 3	92.0	9:58	92.0	
		ж. Э.	T	2,3	_	0.32	0.39	0.43		0.32	_	0.30	_				T		T	1		
	E	on re	T	2.5	21.51	21.48	23,39	21.30		21,45	21.45	21.45			r		T	T	1	1	1	
	OUTLET AIR	8,*	T	8.0	100	10.0	10.0	10.0		0.01	0.03	0.03					T	T	T	1	1	_
	8	H20 Gr/Lb	Ī	28.1	23.0	23.0	37.2	39.2		43.9	53.1	38.2					T	-	T	1		_
EQ.		MOTA CUR		2,32	2.88	2.88	2.79	2.74		1.62	5.93	0.0					T	Ī		1	1	_
CANISTER Canister		e.		82	66.5	67.0	59.5	51.5	valve open	45.0	0.44	39.0	MEN.				ſ			T		_
5.		E-T-BL		47	72.5	F	73.5	7	5 8 8	74.5	76.5	78	AT 425						Ī	Ť	1	_
KO2 + Cat., in	reen)	48.11 In. 1820		0,12	21.0	42.0	89.0	3.20	5.00	6.10	6,95	7.75	DOWN									_
+ 02	OI BO	Pio APii In.H20 In.H20		0270	0.30	0.35	92.0	3.25	5.05	6.10	6.95	7.75	SHUT									_
		% %	20.82	20.85	88,02	20.88	28.02	20.85		20.85	20.85	20.85	20.85	20.82	20.82	20.85	88.02	20.85	20.85	00	20.02	
INLET AIR		<sup>20</sup> №	0.23	0.19	0.20	020	0.23	0.50		0.50	0.21	0.21	0.21	0.23	0.21	0.20	0.30	8.0	8.0	9.		_
IN		H20 Gr/Lb.		87.3	72.0	25.5	77.7	53.7	1	75.3	20.2	72.0	% %	65.3	6.99	71.8	68.5	£.3	73.4	2, 2,		_
	DET.	Man.	٥	35	5		135	195	315	255	315		435	495	555	615	675	735	795			_

 $P_{\rm B}$  = Static pressure just upstream of canister  $T_{\rm B}$  = Temperature just upstream of canister

 $\Delta P_{\Pi}$  = Pressure drop across canister  $\Delta r_{\Pi}$  = Temperature rise across canister

Table 14 (cont.) Partially Reduced Data, P-10

Test P-10

			_	T-	_	_	_		_		_	_	$\boldsymbol{T}$	_	_	_	1	~	_	7	7
	_	ж.о.	L	9.54	0.27	0.27	9.30	0.36		0.31	0.30	0.42	0.33	6.43	0.38	0.39	82.0	82	85	\ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \	<u> </u>
	H	%		12.12	21.48	21.48	21.39	21.21		12.12	21.18	22.09	90.12	21.03	21.03	21.03	76.05	70.07	76.00	78	2
	OUTLET AIR	2°2		0.03	о. 9	40.0	90.0	20.0	ı	60.0	111.0	11.0	47.0	0.14	0.13	0.13	0.13	0.13	l	1	т-
	Б	47/10 02H		35.1	38.7	38.7	48.7		Ι	57.9	62.7	72.27	72.7	9.09	63.8	60.5	$\vdash$	52.7	-	1-	✝
ا ـــا	<b>-</b>	FLOW of m		2.21	2.15	2.16	2.31	_		2,15	2.15	2.15	2.15	2.14	2.12	2.12	2.13	2.13	1	1-	+
CANLSTER MSA		AT.		39.5	47.0	45.5	37.5	$\overline{}$	_	25.5	22.0	19.0	17.0	15.0	11.5	11.5	9.5	7.5	1	Т	Т
CANTE		et.o Ger	l	76.5	92	76.5	_	_		76	76	П	75.5	75.5	75.5	74	74	72.5	ι-	-	Γ
		APr.	T	0.43	0.45	0.45	84.0	0.53	0.57	Н	0.72	$\vdash$	1.00	1,15	1.64	1.78	5.60	3.25	-	-	╀
		In.H2 In.H20	l	8.5	0.53	0.55	Н	$\vdash$	89.0	Н	0.83	1.03	1,12	1.30	88.1	-82	2.84	3.40	┢	1	╁╴
		R.Q.	┢	0.24	0.25	0.26	0.27	0.25	ľ	0.22	0.23	0.23	-	1,	1	-	2	•	3	7	-
	Γ	°, ×	-	21.63	21.69	21.66	21.66	21.66		21.78	21.75	21.75	_							-	ŀ
	Er Afr	20° ≯	H	0.0	0.0	0.0	0.0	0.0	Н	0.0	0.0	0.0			$\dashv$	-	_			-	H
	OUTLER	H <sub>2</sub> O (4		1.9	9.7	5.7	2.6	7.7	-	13.0	35.0	9.1	_		7						-
,	Ч	efm Gr		2.84	2.79	2.76	2,68	2.75	Н	1.30	16.9	0.0		$\dashv$	+	+	1		_	_	
Eter P		o F		11.5	60.0	65.5	59.5	_	valve	Н	┪	30.5	425 MIN.	$\dashv$	$\dashv$	+	$\dashv$				_
CANISTER KO. + Cat. in 20" Canister				Н	73.5 6	┪	-	_		ᅥ	38	78.5 30		$\dashv$	$\dashv$	$\dashv$	+		_		$\vdash$
7.		20 97		27 73		11 74	33 74		13 130W	щ	35 77	-4	-	$\dashv$	$\dashv$	$\dashv$	$\dashv$	$\dashv$	-		
4 Cat.	screen	Pi4 APIG In.H20	Ц	0.27	0.36	0.41	0.83	3.8	4.45	5.97	6.85	7.30	NAMOC	_	4	$\downarrow$	4	_	_		_
Ş	ž Š	Pl4 In.H2		0.38	0.43	8	86.0	3.25	4.70	6.20	7:15	7.45	SHUT								
		8,8	20.82	20.85	20.88	20.88	28.05	20.85		20.85	20.85	20.85	20.85	20.82	20,82	20.85	88.88	20.85	20.85	20,82	
INLET AIR	   	8g №	0.23	0.19	0.20	0.20	0,23	0.20		0.50	0.21	0.21	0,21	0.23	0.21	02.0	0.20	0.20	0.50	0.19	
Ë		H2O Gr/lb.		87.3	72.0	75.5	77.7	63.7		75.3	70.2	72.0	4.08	65.3	6.99	71.8	68.5	54.3	73.4	75.0	
	200	M.n.	٥	R	ξ	8	140	200	220	360	320	380	435	495	560		88	740	8	960	

 $P_{\rm D}$  = Static pressure just upstream of canister  $T_{\rm B}$  = Temperature just upstream of earlister

A.T. a Transcrature

Table 14 (cont.) Partially Reduced Data, P-10

_			_	_		_	_	_	_	_	_	_	_	_	_	_	_	_	_	_	_
		o. C	9.56	°.5	0.50									Ĺ							
	F	%%	20.91	20.88	20.88	20.82															Ī
	OUTLET AIR	8 %	77.0	0.16		0.20															Ī
	50	н <sub>2</sub> о Gr/ло	68.5	6.99	68.5	4.59															T
MSA		MCTA.	2.13	2.13	_	2.12				-										_	T
CANESTER N		ΑŢ.	3.5	3.0	3.5	2,5	AT 1120 MIN.	_				-						-			
CAN	!	Et.	73.0	73.0	72.5	72.0	NMOC									-				_	ľ
		ART In B20	4.35	4.67	5,12	5,35	SHUT														Ī
		In. Age Car. II.	4.45	4.80	5.20	5.40															
		R.Q.	92.0	0.36	0110	0.43	0.37	01/10	0.45	0.40	0.61	0.52	0.33	0.50	0.53	0.50	0.50				
	Fi	Q. 85	21.15	21.15	21.12	22.12	21.12	21.06	21.03	21.06	21.00	21,00	20.02	20.04	20,94	20.94	20.94				Γ
	OUTLET AIR	c <sub>0</sub> 5	20.0	20.0	0.07	20.0	80.0	0.08	60.0	0.09	60.0	0.09	0.11	11.0	11.0	0.11	0.11				
	õ	нго Ст/тр	8.44	49.9	# 9 <u>1</u>	46.2	51.5	53.1	49.6	54.7	51.2	57.9	57.9	54.3	55.9	62.7	62.7				
8-1		FLOW	2.75	2.75	2.75	2.75	2.75	2,75	2.79	2.79	2.79	2.79	2.79	2.79	2.70	2.79	2.70				
CANISTER		ÅF.	34	34	33.5	32	33	30.5	28.5	27.5	Я	25.5	24.5	22.5	22.5	21.5	20.5				
20. 20.		154 104	70	70	69.5	89	77	70	72	72.5	72	73.5	72	71.5	73.5	7.5	2	AT 1765 MIN			
at., in	Screen	AP.13 In.H20	0.20	0.20	0.20	0.50	0.20	0.20	0.20	0.20	0.20	0.20	0.20	0.20	0.20	22.0	22.0	DOWN			
, cox	+ Ann. Screen	Pia APia In.H20 In.H20	0.28	0.28	0.28	0.28	0.28	0.28	0.30	0.32	0.32	0.32	0.32	0.32	0.32	48.0	0.34	SHUT			
		જ જ	20.82	20.82	20.82 0.28	20.82	20.82 0.28	20.79	20.76	20.79	20.82	20.79	20.76	20.76	20.73	20,76	20.76				
INLET AIR		ું ×	0.19	0.19	61.0	°.3	0.19	0.19	0.20	0.20	0.20	02.0	0.20	0.20	0.20	02.0	02.0				
Ä		H20 Gr/lb.	75.0	75.0	76.7	81.7	70.1	70.1	68.5	71.8	70,3	71.8	70.3	73.4	73.4	73.4	75.0				
	TOCK	Man.	915	275	1035	1095	1155	1215	1275	1335	1395	1455	1515	1575	1635	1695	1755				

 $P_{\rm B}$  = Static pressure just upstream of canister  $\mathbb{T}_{\rm B}$  = Temperature just upstream of canister

 $\Delta P_{\rm R}$  = Pressure drop across canister  $\Delta T_{\rm R}$  = Temperature rise across canister

Table 14 (cont.) Partially Reduced Data, P-10

	_	я.э.	L					L					L	L	L			
	95	Q, A							L									
	OUTLET AIR	S 80																
	8	47/19 61/19																
		FLOW																
CANISTER		4.0°																
CAN		Et u																
		AP In.H2C																
		In.H20 In.H20	Γ															
		, .		0.16	0.30	0.38	0.45	0.48	0.46	0.53	6.67	1.00	1.83	2.30	1.20			
	PS.	ું હ		22.05	21.42	21.33	21.21	21.18	21.15	21.12	23.06	20.97	20,88	20,85	20.94	_		
	OUTLET AIR	8, 8		0.01	0.03	0.03	2.04	40.0	40.0	0.03	0.03	0.04	0.09	0.12	0.13			
	10	H2O Gr/Lb		45.1	32.4	32.4	32.9	37.2	40.3	43.5	42,7	53.1	52.4	9.09	60.5			
1-2 1n 5"		PLOW of a		2.76	2.74	2.72	2.72	2.72	2.72	2.72	2.72	2.72	2.72	2.72	2.72			_
MISTER 3		• F.		18.5	54.0	51.0	45.5	48.5	46.5	43.5	43.0	37.0	25.5	21.5	17.0	MEN.		
. + CA	screen	(+10 -]g-i		76	g	2	98	68.5	2	69	69	69	63	69.5	69.0	AT 680		
2/3 KO, + Cat. + 1/3 L10H, 1n 5"	ter (no	Ln.H20		0.22	0.25	9.44°C	9.65	1.28	2.72	4.00	4.05	3.90	4.05	4.05	3.85	NMOC		
2/3 K	Canis	Fio APH In.H20 In.H20		0.34	0.36	24.0	08.0	04.1	2.80	4.10	4.10	3.95	4.20	4.10	3.95	SHUT		
		0, Pr	20.92	20.82	20.85	20.83 0.42	20.85 0.80	20.85 1.40	20.82	20.82 4.10	20.82	20.82 3.95	25.62	20.82	20.79 3.95			-
INLET AIR		% % %	0.21	0.21	0.20	02.0	0.20	0.20	0.19	0.19	0.19	0.39	0.50	0.19	91.0			
INI		H20 Gr/Lb.	6.99	6.99	27.8	68.5	64.3	73.4	75.0	75.0	75.0	76.7	81.7	70.1	70.1			
	TIME	Min.	٥	5	65	125	185	245	305	365	425	485	545	605	599	680		

 $P_{\rm n}$  = Static pressure just upstream of canister:  $T_{\rm n}$  = Temperature just upstream of canister

 $\Delta P_{\rm B}$  . Pressure drop across canister  $\Delta T_{\rm B}$  = Temperature rise across canister

Table 15 IBM Reduced Data, P-1

TES# P-1	K02 A	TMOSPHERIC COMP CANISTER A	KØ2 ATMØSPHERIC COMPØSITION CONTROL EVALUATION CANISTER A	EVALUATION	DATE 12-18-1961	12-18-1961
	(KOZ IN SCREEN			(1/3KG2+1/	(1/3KG2+1/3LAVA RGCK+1/3ACT.ALUMINA)	ALUMINA?
TIME (MINUTES)	G2 GENERATED	CG2 ABSGRBED	H2G ABSORBED	02 GENERATED	CGZ ABSGRBED	HZG ABSGRBED
15.	0.02480	0.01607	•	0.02556	0.00852	0.
45.	0,02480	0.02113	•	0.01307	0.01259	•
• 06	Ó. 32067	0.01424	•	0.01038	0.00890	•
125.	0.02229	0.01486	•	0.01018	0.00776	.0
150.	0.01672	0.01254	•0	0.00710	00600	•0
180	0.01409	0.01268	•	0.00425	0.00662	
215.	0.00720	0.01153	•	0.00567	0.00520	•
245.	0.01746	0.01063	•	0.00709	0.00520	
			}			<b>;</b>
TEST P-1	K62 AT	MØSPHERIC COMPO	KG2 ATMOSPHERIC COMPOSITION CONTROL EVALUATION		DATE 12-18-1961	12-18-1961
	(1/3K62+1/3	CANISTER C	LIOH )		CANISTER MSA	
TIME	32 GENERATED	CG2 ABSGRBED	H2G ABSORBED	G2 GENERATED	CA2 ARCERED	HOG ARCGORED
(MINUTES)	(SCFM)	(SCFM)	(GR/MIN)	( VCEM)		CONTRINI
s,	0.02809	0.02267	•	0.03989	0.02231	0.
•0 <del>4</del>	0.02514	0.02563	•	0.02961	0.02103	
- 00	0.01776	0.02071	•	0.02189	0.01416	•
115.	0.01331	0.02318	•	0.01674	0.01502	•
145	0.00740	0.01825	•	0.01032	0.01032	•
175.	###00 ° 0	0.01578	•	14900°0	0.00730	•
205	###OO * O	0.01331	•	0.00515	0.00515	ೆ
235.	0.00592	0.01184	•	0.00773	0.00472	ဝီ

Table 16 IBM Reduced Data, P-2

DATE 12-21-1961 12-21-1961 CANISTER B [1/3KG2+1/3LAVARGCK+1/3ACT.AL)	3ED H2G ABSGRRED (GR/MIN) 0. 0. 0. 0. 0. 0. 0. 0. 0.	ED H2G ABSØRBED (GR/MIN)  ED H2G ABSØRBED  (GR/MIN)  0.0000
DATE 12-21: Canister B KØ2+1/3LAVARG	CG2 ABSGRBED (SCFM) (SCFM) 0.01754 0.01754 0.01210 0.00874 0.00341 0.00439 0.00439 0.00439 0.00439 0.00439	CG2 ABSGRBED (SCFM) (SC
-	92 GENERATED (SCFM) 0.06023 0.03587 0.03587 0.02420 0.01310 0.01332 0.01342 0.01342 0.00774 0.00774 0.00506 0.00506	EVALUATION  02 GENERATED  12 GENERATED  12 GENERATED  12 GENERATED  13 GENERATED  14 GENERATED  15 GENERATED  16 GENERATED  17 GENERATED  16 G
GSITION CONTROL	H20 ABSGRBED (GR/MIN) 0.00000000000000000000000000000000000	173110N CONTROL 173110H } H20 ABSGRBED (GR/MIN) 0. 0. 0. 0. 0. 0. 0. 0. 0. 0. 0. 0. 0.
KG2 ATMGSPHERIC COMPOSITION CONTROL EVALUATION CANISTER A ( KG2 IN SCREEN )	CG2 ABSGRBED (SCFM) 0.01857 0.02179 0.01166 0.01162 0.00910 0.00726 0.00678 0.00638 0.00678 0.00719 0.00719 0.00719	KGZ ATMGSPHERIC COMPGSITION CONTROL CANISTER C (1/3KG2+1/3LAVARGCK+1/3LIGH ) ATED CG2 ABSGRRED H2G ABSGRRED (SCFM) 0.002838 0.00000.0001988 0.0001988 0.0001988 0.0001989 0.0001934 0.0001934 0.00000000000000000000000000000000000
K 02 A T	02 GENERATED (SCFM) 0.03847 0.03268 0.02421 0.02421 0.02861 0.01361 0.01549 0.01547 0.01547 0.01547 0.01547 0.01547	KGZ A1 (1/3k 02 GENERATED (SCFM) 0.06999 0.01620 0.01620 0.01631 0.01238 0.01238 0.01238 0.01238 0.01238 0.01238 0.01238 0.01238 0.01238 0.01238
TEST P2	TIME (MINUTES) 5. 45. 75. 105. 105. 105. 2255. 285. 285. 315. 345. 345. 405.	TEST P2  TIME (MINUTES) 12- 51- 60- 170- 200- 230- 250- 250- 350- 350- 350- 408-

Table 17 IBM Reduced Data, P-3

TEST P3	KG2 A	MOSPHERIC COMPC	KG2 ATMOSPHERIC COMPOSITION CONTROL EVALUATION	EVALUATION	DATE 12-20-196	12-20-1961 12-20-1961
	( K02 IN S	CANISTER A SCREEN	-	(1/3K02+1/3	CANISTER B (1/3KG2+1/3LAVARGCK+1/3 ACT.ALUMINA)	, ALUMINA)
W. Z.	02 GENERATED	CO2 ABSORBED	H2G ABSØRBED	62 GENERATED	CO2 ABSORBED	H20 ABSORBED
(MINUTES)	(SCFM)	(SCFM)			(SCFM)	(GR/MIN)
75.	0.03882	0.0194;	.0	0.04344	0.02096	•
55.	0.03681	0.01654	•	0.02829	0.01509	•
75•	0.03348	0.01594	•	0.02713	0.01281	•
107.	0.03342	0.01326	•	0.02376	0.01131	•
135.	0.03004	0.01287	•	0.02137	0.01012	<b>.</b>
165.	0.02544	0.01299	•	0.01543	0.00903	•
195.	0.02273	0.01137	0.	0.01237	0.00712	•
225.	0.01959	0.01088	•0	0.00886	0.00554	•
255.	0.01566	0.01086	0	0.00858	0.00572	ď
280	0.01305	0.00816	0	0.00429	19 to 0 0	Č
315.	0.01313	0.00875	•	0.00637	0,00400	
345	0.00831	0.00720	•0	0.00533	0.00391	ic
375.	0.01140	0.00706		0.00639	0.00391	
405	0.00551	0.00489		0.00425	0.00283	
435	68400*0	0.00489	•	0.00425	0.00283	0
TEST P3	KG2 A	FMOSPHERIC COMP	KOZ ATMOSPHERIC COMPOSITION CONTROL EVALUATION	EVALUATION	DATE 12-20-1961	1 12-20-1961
		CANISTER C			CANISTER MSA	
	(1/3K02+1/	[1/3K62+1/3LAVARGCK+1/3LI0H	- H	_		^
FIME	02 GENERATED	CG2 ABSGRBED	H2G ABSORBED	02 GENERATED	CA2 ABSGRBED	H2G ABSORBED
(MINUTES)		(SCFM)	(GR/MIN)		(SCFM)	(GR/MIN)
30.	996400	0.03831	0	0.03546	0.01549	•
65.	0.02783	0.02830	•	0.02728	0.03447	•
83.	0.01806	0.02084	•	0.02089	0.01024	•0
110.	0.01249	0.01665	•	0.01591	0.00775	•
140.	0.01064	0.01573	•	0.01172	0.00728	•
170.	0.00925	0.01295	•	0.00920	009000	•
200.	0.00826	0.01147	•	0.00726	0.00524	•
230.	0.00552	0.00827	•	0.00488	0.00447	•
260.	0.00688	0.00826	•	46400*0	0.00453	•
290.	0.00275	0.00826	•	0.00367	0°00##8	•
320.	0.00551	0.00689	ċ	06200.0	0.00531	•
350	0.00418	0.00650	•0	0.00488	0.00529	<b>.</b>
385.	0.00556	0.00695	•	0.00369	0.00369	•
# 10 <b>.</b>	0.00420	0.00513	°	0.00367	0.00286	•
*011	0.00420	0.00513	•	0.00246	0.00246	•0

Table 18 IBM Reduced Data, P-4

12-19-1961 12-19-1961 TER B IVATED ALUMINA 1	HZG ABSGRBED Co. O. O. O. O. O.	12-19-1961 1 12-19-1961 120 ABSGRBED 130 0.0000000000000000000000000000000000
ATION DATE 12-19-1961 12- CANISTER B (KO2-LAVA ROCK-ACTIVATED ALUMINA	C 02 ABS 3R BED (SCFM) 0.015 49 0.01208 0.00826 0.00826 0.00783	DATE 12-19-1961 12-19-1961 CANISTER MSA ) 1 ( CG2 ABSGRBED H2G ABSGRBED (SCFM) (GR/MIN) 0.02272 0.0001156 0.00000000000000000000000000000000000
EVALUATION D	02 GENERATED (SCFM) 0.02459 0.0282 0.02017 0.01565 0.01261 0.01000	24 TED
KOZ ATMOSPHERIC COMPOSITION CONTROL EVALUATION CANISTER A 1 (KOZ-	H2G ABSGRBED (GR/MIN) 0. 0. 0. 0. 0.	K02 ATMGSPHERIC CGMPGSITIGN CGNTRGL EVALUATION LIGH-LAVA ROCK  ATED C02 ABSGRBED H2G ABSGRBED G2 GENEI (SCFM) (GR/MIN) 0.00590 0.02118 0.00594 0.01973 0.00783 0.01504 0.01504 0.01504 0.0066
TMØSPHERIC CØMPI Canister a Reen	CGZ ABSGRBED (SCFM) 0.01447 0.00913 0.00913 0.00932 0.00932 0.00955	FMOSPHERIC COMPO CANISTER C LAVA ROCK COZ ABSGRBED 0.02500 0.02118 0.01973 0.01732 0.01504 0.01411
KO2 ATMOS CA (KO2 IN SCREEN	02 GENERATED (SCFM) 0.02316 0.02167 0.01979 0.01657 0.01568 0.01435	K02 ATM0SPHERI CANISTE (K02-LIGH-LAVA ROCK 02 GENERATED C02 ABS (SCFM) 0.0250 0.01976 0.01976 0.0150 0.01081 0.00940 0.0126
TES¶ P.4	TIME 20. 20. 20. 95. 95. 95. 242.	TEST P4  TIME  TIME  (MINUTES)  50  50  135  175  213

Table 19 IBM Reduced Data, P-5

TEST P5	KG2 A	KG2 ATMOSPHERIC COMPOSITION CONTROL EVALUATION CANISTER A-2 (KG2(NG CATALYST)IN SCREEN )	ØSITIØN CØNTRØL N	EVALUATION (KG2-LAVA	ATION DATE 12-29-196 CANISTER A-3 KKG2-LAVA RGCK-MS X-13 IN S	12-29-1961 12-29-1961 TER A-3 X-13 IN SCREEN )
TIME (MINUTES) 17. 47. 68. 102. 132. 132. 207. 227. 2527. 257. 257. 257. 317.	92 GENERATED (SCFM) 0.03402 0.0272 0.0272 0.01701 0.01701 0.01303 0.01303 0.01309	CG2 ABSGRBED (SCFN) 0.01606 0.01017 0.00852 0.00786 0.00786 0.00531 0.00485 0.00485	H20 ABSGRBED (GR/MIN) 0. 0. 0. 0. 0. 0.	62 GENERATED { SCFM} 0.01352 0.01080 0.01082 0.01082 0.00945 0.00278 0.00278 0.00417 0.00431	CG2 ABSGRBED (SCFM)- 0.00901 0.00585 0.00586 0.00541 0.00417 0.00334 0.00335	HZG ABSGRBED (GR/MIN) 0. 0. 0. 0. 0. 0.
TEST P5	KUZ ATA	KOZ ATMOSPHERIC COMPOSITION CONTROL EVALUATION CANISTER D (2/3KO2+1/3LIOH WITH PROBES ) (	ITION CONTROL E		DATE 12-29-1961 CANISTER MSA	12-29-1961
TIME (MINUTES) 22. 22. 73. 107. 137. 137. 209. 232. 292.	02 GENERATED (SCFM) 0.05474 0.03556 0.02826 0.02714 0.01711 0.01585 0.01008 0.01152 0.01010	CG2 ABS GRBED (SCFM) 0.02209 0.01897 0.01790 0.01523 0.01152 0.01152 0.00914 0.00922	H23 ABSGRBED (GR/MIN) 0.00000000000000000000000000000000000	02 GENERATED (SCFM) 0.03041 0.0345 0.01802 0.01577 0.00901 0.00334 0.00557 0.00557	CG2 ABSGRBED F (SCFM) 0.01389 0.0151 0.0058 0.00488 0.00488 0.00485 0.00571 0.00297 0.00297 0.00297 0.00324	H20 ABSGRBED (GR/MIN) 0. 0. 0. 0. 0.

Table 19 IBM Reduced Data, P-5

TEST P5	KOZ (NG C	KOZ ATMOSPHERIC COMPO CANISTER A-2 (KOZ(NO CATALYST)IN SCREEN	KG2 ATMGSPHERIC CGMPGSITIGN CGNTRGL EVALUATION CANISTER A-2 NG CATALYST)IN SCREEN ) (KG2-	EVALUATION (KG2-LAVA	ATION DATE 12-29-196 CANISTER A-3 (KG2-LAVA RGCK-MS X-13 IN S	12-29-1961 12-29-1961 STER A-3 X-13 IN SCREEN )
TIME (MINUTES) 17. 47. 68. 102. 132. 207. 227. 227. 252.	02 GENERATED (SCFM) 0.03402 0.02913 0.02972 0.02556 0.01942 0.01134 0.01134 0.01163 0.01163	CG2 ABSGRBED (SCFM) 0.01606 0.01017 0.00852 0.00852 0.00756 0.00614 0.00631 0.00485 0.00485	H20 ABSGRBED (GR/MIN) 00.00.00.00.00.00.00.00.00.00.00.00.00.	02 GENERATED 8 SCFM) 0.01352 0.01080 0.01082 0.01082 0.00945 0.00278 0.00278 0.00431	D CG2 ABSGRBED (SCFM). 0.00901 0.00586 0.00586 0.00586 0.00417 0.00417 0.00324 0.00336	H2G ABSGRBED (GR/MIN) 0. 0. 0. 0. 0. 0.
TEST P5	KG2 AT (2/3KG2+1/3	KOZ ATMOSPHERIC COMPOS Canister D (2/3k02+1/3lioh With Probes	KO2 ATMGSPHERIC COMPOSITION CONTROL EVALUATION CANISTER D (22-1/3LIOH WITH PROBES ) (	VALUATI GN	DATE 12-29-1961 CANISTER MSA	12-29-1961
TIME 22. 22. 52. 73. 107. 137. 172. 209. 232. 257. 322.	02 GENERATED (SCFM) 0.05474 0.03556 0.02826 0.02714 0.01711 0.01585 0.01008 0.01152	CG2 ABSGRBED (SCFM) 0.02209 0.01790 0.01790 0.01523 0.01152 0.01152 0.01056 0.00914 0.00922	HZG ABSGRBED (GR/MIN) 0. 0. 0. 0. 0. 0. 0.	92 GENERATED (SCFM) 0.03041 0.02365 0.01802 0.01577 0.00901 0.00557 0.00557 0.00552	CG2 ABSGRBED (SCFM) 0.01389 0.01051 0.00788 0.00488 0.00451 0.00451 0.0027	H2G ABSGRBED (GR/MIN) 0. 0. 0. 0. 0.

Table 18 IBM Reduced Data, P-4

TEST P.4	KG2 ATMGS CAI (KG2 IN SCREEN	KGZ ATMOSPHERIC COMPOSITION CONTROL EVALUATION CANISTER A IN SCREEN	OSITION CONTRO	L EVALUATION {KG2-LAVA	ATION DATE 12-19-1961 12- CANISTER B. CANISTER A. HINESER	12-19-1961 12-19-1961 FER B
TIME (MINUTES) 20. 55. 95. 140. 180. 218. 242.	02 GENERATED (SCFM) 0.02316 0.02167 0.02182 0.01979 0.01657 0.01658	CGZ ABSGRBED (SCFM) 0.01447 0.01011 0.00913 0.00913 0.00952 0.00865	H2G ABSGRBED (GR/MIN) 0. 0. 0. 0.	92 GENERATED (SCFM) 0.02459 0.02282 0.02282 0.01261 0.01261 0.01000 0.01000	C 02. ABSGR8ED (SSFM) (SSFM) (SOFM) 0.01549 0.01549 0.01076 0.00826 0.00826 0.00782	H2G ABSGRBED (GR/MIN) 0. 0. 0. 0.
TEST P4	KG2 ATMGSPHERI Caniste (KG2-Ligh-lava Rock	KG2 ATMOSPHERIC COMPOSITION CONTROL EVALUATION CANISTER C LIGH-LAVA ROCK )	SITION CONTROL		DATE 12-19-1961 CANISTER MSA	12-19-1961
MIME (MINUTES) 5. 5. 50. 90. 135. 175. 237.	92 GENERATED (SCFM) 0.05CF42 0.02888 0.02599 0.01876 0.01363 0.01081	CG2 ABSGRBED   SCFM    SCFM    0.0250   0.01973   0.01732   0.01504   0.01411	H2G ABSGRBED (GR/MIN) 0. 0. 0. 0.	02 GENERATED (SCFM) 0.00545 0.02184 0.01663 0.01251 0.00967 0.00683	CG2 ABSGRBED (SCFM) 0.02272 0.00767 0.00542 0.00542 0.00504	H20 ABSGRBED (GR/MIN) 0. 0. 0. 0.

Table 19 IBM Reduced Data, P-5

TEST P5	KG2 A	KG2 ATMOSPHERIC COMPO CANISTER A-2 (KG2(NG CATALYST)IN SCREEN	KG2 ATMOSPHERIC COMPOSITION CONTROL EVALUATION CANISTER A-2 NO CATALYST)IN SCREEN ) (KG2-	EVALUATION (KG2-LAVA	ATION DATE 12-29-1961 12- CANISTER A-3 IKG2-LAVA RGCK-MS X-13 IN SCREEN	.1 12-29-1961 CREEN )
11ME (MINUTES) 17. 47. 68. 102. 132. 167. 207. 227. 252. 287.	92 GENERATED (SCFM) (SC	CG2 ABSGRBED [ SCFM ) 0.01606 0.01017 0.00852 0.00856 0.00786 0.00786 0.006314 0.00485 0.00485	H20 ABSGRBED (GR/MIN) 0. 0. 0. 0. 0. 0. 0.	92 GENERATED (SCFN) 0.01352 0.01080 0.01082 0.01217 0.0034 0.00278 0.00278 0.00431 0.00431	0 CG2 ABSGRBED (SCFM)** 0.00901** 0.00585 0.00586 0.005841 0.00417 0.00324 0.00336	HZG ABSGRBED (GR/MIN) 0. 0. 0. 0. 0.
TEST P5	KG2 ATM	KOZ ATMOSPHERIC COMPOS CANISTER D (2/3KO2+1/3LIOH WITH PROBES	KOZ ATMOSPHERIC COMPOSITION CONTROL EVALUATION COLTROL CANISTER D (22+1/31.10H WITH PROBES ) (	V A <u>l</u> UATI GN: [	DATE 12-29-1961 CANISTER MSA	12-29-1961
TIME 22. 22. 73. 73. 107. 107. 137. 232. 232.	02 GENERATED (SCFM) 0.03474 0.03556 0.02826 0.02714 0.01585 0.01088 0.01152 0.01010	CG2 ABSGRBED (SCFM) 0.01209 0.01897 0.01790 0.01523 0.01152 0.01152 0.01152 0.01056 0.00914 0.00922	H20 ABSORBED (GR/MIN) 0.00.00.00.00.00.00.00.00.00.00.00.00.0	02 GENERATED [SCFM] 0.03041 0.02365 0.01802 0.01577 0.00901 0.00901 0.00557 0.00557 0.00557	CGZ ABSCRBED 1 (SCFM) 0.01389 0.0158 0.00788 0.00488 0.00446 0.00297 0.00294	H20 ABSORBED (GR/MIN) 0. 0. 0. 0. 0.

Table 20 IBM Reduced Data, P-6

TEST P-6	K02 A1	KOZ ATMOSPHERIC COMPOSITION CONTROL EVALUATION CANISTER A-2 (KOZ (NO CATALYST) IN SCREEN	ISITIAN CONTROL :en	EVALUATION (2/3 KG2(NO	ATION DA:E 1-10-1962 1-10- CANISTER A-4 (2/3 KG2(NO CAT.)+1/3 MSX13 IN SCRN)	1-10-1962 IN SCRN)
TIME 114. 56. 66. 116. 1266. 2266. 326. 326. 446.	02 GENERATED (SCFM) 0.00342 0.00515 0.00515 0.00428 0.00428 0.00770 0.00598	C02 ABS GRBED (SCFM) 0.00285 0.00229 0.00200 0.00228 0.00228 0.00257 0.00256	H20 ABSGRBED (GR/MIN) 0.580 0.146 1.142 0.083 0.559 0.559 0.683 2.092 0.807	02 GENERATED (SCFM) 0.00084 0.00084 0.00168 0.00167 0.00167 0.0084 0.0084 0.00501	C62 ARSGRBED (SCFM) 0.00251 0.00176 0.00112 0.00112 0.00084 0.00139 0.00223	H2G ABSGRBED (GR/MIN) 2-10727 2-03152 2-61870 1-25860 1-25860 1-35486 1-37517 2-08297 1-09204
TEST F-6	KØ2 AT (2/3 KØ2INO	KOZ ATMOSPHERIC COMPOSITION CONTROL EVALUATION CANISTER A-5 (2/3 KOZINO CAT)+1/3 LIOH IN SCREEN)	SITTON CONTROL IN SCREEN)		DATE 1-10-1962 Canister Msa	1-10-1962
TIME (MINUTES) 19. 61. 91. 91. 121. 121. 216. 276. 336. 356. 456.	G2 GENERATED (SCFM) 0.00423 0.00507 0.00507 0.00422 0.00337 0.00505 0.00505	CG2 ABS GRBED { SCFM } { SCFM } 0.00423 0.00254 0.00253 0.00253 0.00253 0.00253 0.00253 0.00309	H20 ABSORBED (GR/MIN) (GR/MIN) 0.401 0.470 1.125 0.682 0.531 -0.245 0.672 3.239 0.590	02 GENERATED (SCFM) C. 00531 0.00532 0.00473 0.00414 0.00354 0.00354 0.00589	CG2 ABSGRBED F 1 SCFM) 0.00534 0.00276 0.00276 0.00276 0.00216 0.00216 0.00216 0.00275 0.00275 0.00157	H2G ABSGRRED (GR/MIN) 0.7 H2O3 0.18565 0.42889 0.05712 0.37102 0.15684 0.79846 1.35304 0.49904

Table 21 IBM Reduced Data, P-7

rest P-7	K02 AT	KGZ ATMGSPHERIC COMPOSITION CONTROL EVALUATION CANISTER A	SITION CONTROL	EVALUATION	DATE 1-19-1962 CANISTER A-6	2 1-19-1962
	KGZ IN SCREEN	EEN	-	1 Z 0 X C / Z 1	12/3 AUZ + 1/3 Mese AIS IN SCACEN	SCREEN J
7	AS GENERATED	CO2 ABSORBED	H2G ABSGRBED	62 GENERATED	CO2 ABSGRBED	H20 ABSGRBED
(NINITES)	(SCFM)	(SCFM)	(GR/MIN)	(SCFM)	(SCFM)	(GR/MIN)
15.	0.01117	0.00315	7.723	0.01169	0.00334	9.12746
	0.01557	0.00346	8.692	0.01174	0.00335	8.88920
75.	0.01479	0.00261	8.758	0.01172	0.00195	8.91289
120.	0.01477	0,00348	8.353	0.01171	0.00335	7.39911
150	0.01389	0-00260	8.014	0.01087	0.00251	7.07828
180.	0.01474	0.00318	7 - 339	0.01170	0.00306	6,75246
200	0-01128	0.00318	4.674	0.01003	0.00306	6.43187
2002	00000	0.00318	3.518	0.00837	0.00307	4.83785
	0.01071	0.00238	8-662	0.00957	0.00232	8.11198
• • • • • • • • • • • • • • • • • • • •			1 2 2 2 2 2 2 2 2 2 2 2 2 2 2 2 2 2 2 2	83000	19000	7.24844
¥20.	0.01043	0.00252	977.	00,000	107000	****
TECT P.7	KG2 A	KG2 ATMGSPHERIC CGMPGSITION CONTRGL EVALUATION	GSITION CONTROL	EVALUATION	BATE 1-19-1962	52 1-19-1962
		CANISTER A-7		•	CANISTEN ASA	_
	(2/3 KG2 +	_	REEN )	-		
		CAS ABCADAGO	MOR ARSGRAFD	02 GENERATED	CO2 ABSORBED	H20 ABSGRBED
HE H	02 GENERALED	COZ ABSURBEU	NIN STATE	(MHC)	( SCFM)	(CR/MIN)
(MINUTES)	(SCFM)	( SC + M )	010 3	0.01382	0.00329	8.06926
20.	0.01682	400000	4 4 48	0.01576	0.00328	8.52881
50•	0.01351	0.00558	0.0.0	0.01509	0.00306	7.85920
80.	0.01184	0.000	- 77.0	0.01396	0.00355	7.19682
125.	0.01217	0.00319	0.004	79110-0	0,00288	6.70323
155.	0.01131	0.00261	0.234	01110	0.00285	5.62743
185.	0.01130	0.00319	200	0.00851	0.00284	4.82702
245	0.00929	0.00253	20404	0000	0.00231	2.00796
305	0.00712	0.00326	4.000	0.00511	0.00187	3.78107
365.	0.800	0.00257	900	0.00569	0.00190	4.66939
425.	0.00800	0.00267	0.420			

Table 22 IBM Reduced Data, P-8

TEST P-8	KG2 AT	KOZ ATMOSPHERIC COMPO	COMPOSITION CONTROL	EVALUATION	DATE 1-11-1962	1-12-1962
	IKOZ IN SCREEN	EEN	-	12/3 K02 +	12/3 KO2 + 1/3 M.S. X13 IN SCREEN	CREEN )
TIME	82 GENERATED	CO2 ABSORBED	H2G ABSGRBED	02 GENERATED	C 02 ABS GRBED	H20 ABSGRBED
(MINOTES)	(SCFM)	( SCFM)	(GR/MIN)	(SCFM)		(GR/MIN)
13.	0.00631	0.00420	0.589	0.00255	0.00339	0.63665
, E. A.	0.00757	0.00477	-1.018	0.00170	0.00170	1,15223
73.	0.01092	0,00560	-0.407	0.00255	0.00170	1.93050
103.	0.01342	0.00615	-0.629	0.00424	0.00141	1.45695
148.	0.01501	0.00711	1.855	0.00763	0.00367	3.85784
183.	0.01186	0.00633	1.570	0.00848	0.00480	2,85234
208.	0.01106	0.00632	1.587	0.00848	0.00396	2.87286
238.	0.01106	0.00553	0.975	0.01017	0.00537	1.82632
268.	84600*0	0 <b>•</b> 00448	0.459	0.00763	0.00396	1.70320
298.	0.00711	0.00448	0.593	0.00763	0.00367	1.39539
328.	0.00632	0.00421	0.803	0.00424	0.00283	1.78528
358.	0.00553	0.00342	0.860	0.00593	0.00283	1.45695
388.	42 400 ° 0	0.00316	0.019	0.00339	0.00254	0.82082
423.	0.00971	0.00383	1.190	0.00439	0.00263	2-66617
# <b>68</b> •	0.00531	0.00501	166.0	0.00351	0.00351	2,51566
533.	0.00705	0.00470	0.086	0.01049	0.00408	0.36413
593.	0.01497	0.00616	3.731	0.01485	0.00670	4.30051
653.	0,00967	0.00586	5.016	0.01046	0,00668	4.86911
713.	0.00877	0.00526	3.245	0.00870	0.00522	3.90056

Table 22 (cont.) IBM Reduced Data, P-8

TEST P-8	KG2 A	KGZ ATMOSPHERIC COMPOSITION CONTROL CANISTER A-7	ISITION CONTROL	EVALUATION	DATE 1-11-1962 CANTSTER MSAT	1-12-1962
	(2/3 K02 +	(2/3 KG2 + 1/3 LIGH IN SCREEN	(EEN )	_		^
TIME	32 GENERATED	CG2 ABSGRBED	H20 ABSGRBED	32 GENERATED	CG2 ABSGRBED	H20 ABSORBED
(MINUTES)	(SCFM)	(SCFM)	(GR/MIN)	( SCFM)	(SCFM)	(GR/MIN)
23.	0.01518	0.00590	-1.755	0.00784	0.00435	-0.71065
53.	0.00842	0.00646	-1.815	0.00916	0.00654	0.36392
83.	0.00842	0.00533	-0.733	0.01045	0.00653	0.63169
113.	0.00841	0.00448	-1.302	0.01631	0.00827	-1.46705
153.	9960000	0.00615	1.382	0.01306	0.00762	1.27917
186.	0.00878	0.00556	0.383	0.01110	0.00588	1.04228
213.	0.00878	0.00556	0.702	0.01046	0.00610	1.29639
243.	0.00878	0.00586	0.021	0.00850	0.00523	0.79048
273.	0.00703	0.00468	-0-191	0.00523	0.00348	0.80540
303.	0.00791	0.00410	-0.340	0.00392	0.00262	0.72725
333.	0.00439	0.00381	0.234	0.00327	0.00218	1.12125
363.	0.00439	0.00351	0.298	0.00262	0.00218	1.12249
393.	0.00351	0.00293	-0.659	0.00262	0.00218	0.63239
433.	0.00674	0.00281	0.991	0.00326	0.00261	1.13368
473.	0.00336	0.00392	-0.021	0.00261	0.00283	0.87755
543.	0.00503	0.00448	-0.884	0.00521	0.00390	0.93963
598.	0,01003	0.00557	1.905	0.00779	0.00541	3.19260
658.	0.00834	0.00528	3.128	0.00648	0.00454	3.25253
718.	0.00666	0.00472	1.776	4650000	0.00302	2,20433

Table 23 IBM Reduced Data, P-9

TEST P-9	KG2 A	KG2 ATMGSPHERIC COMPC	COMPOSITION CONTROL EVALUATION	EVALUATION	DATE 1-16-1962	2 1-16-1962
	(KØZ IN SCREEN	REGN A	^	(2/3 KG2 +	+ 1/3 M.S. X13 IN	SCREEN )
TIME	02 GENERATED	CO2 ABSORBED	H20 ABSORBED	02 GENERATED	CO2 1650RBED	H20 ABSORBED
(MINOTES)	E 100	CSCFR	(GR/MIN)	(SCFM)	(SCFM)	(GR/MIN)
<u>.</u>	0.01274	0.00661	10.025	0.01145	0.00716	12,80581
	0.01128	0.00329	3.181	0.01274	0.00519	10.56784
. 25	0.01265	0.00328	4.736	0.01411	0.00376	10.74578
15.	0.01264	0.00328	5.820	0.01268	0.00517	10.72591
145.	0.01265	0.00328	1.976	0.01269	0.00423	4.58069
175.	0.01123	0.00328	3.608	0.01132	0.00424	5.69494
205.	0.01264	0.00187	7.011	0.03273	0.00189	7.68473
240.	0.01263	0.00327	5.579	0.01272	0.00283	7.54091
265.	0.00874	0.00340	1.024	0.00886	0.00148	2.75894
310.	0.00854	0.00237	6.620	0.01016	0.00242	7,17959
355.	0.00419	0.00140	2.676	0.00841	0.00187	3,80567
385.	0.00710	0.00331	2,752	0.00571	0.00333	h. h2761
445.	0.00567	0.00189	3.846	0.00423	0.00188	11.05238
505.	0.00708	0.00283	3,00	0.00712	0.00285	h 07180
565.	0.00713	0.00808	5.048	0.00718	0.00958	5.46847
TEST P-9	K02 A1	KOZ ATMOSPHERIC COMPOSITION CONTROL EVALUATION	SITION CONTROL	EVALUATION	DATE 1-16-1962	1-16-1962
		CANISIER A-			CANISTER MSA	
	(2/3 K62 +	1/3 LIOH IN SCREEN	EEN )	_		^
TIME	02 GENERATED	CO2 ABSORBED	H2G ABSGRED	G2 CENEDATED	CAS ABCADRED	01000000
(MINUTES)	(SCFM)	( SCFM)	(GR/MIN)		COC AGSONOFO (ATEM)	COLMINI
20.	0.02260	0,00800	9.420	0.01427	0-00732	11.30556
55.	0.01403	0.00515	2,382	0.01207	20:00:0	3,13482
80.	0.01557	0.00378	7.519	0.01205	0.00329	6.23159
135.	0.01415	0.00613	7.513	0.01098	9240000	5.39202
.05	0.01273	0.00519	1.990	0.00958	0.00390	1.90567
000	0.01152	0.00424	2.230	0.00806	0.00302	2,19431
.012	0.01273	0.00424	5.455	0.00965	0.00354	4.43161
240	18600.0	0.00517	5.849	0.00618	0.00265	3,15565
*0.2 212	0.00884	0.00393	1.500	0.00517	0.001,15	0.64432
240	0.00883	0.00392	5.388	0.00657	0.00192	3.44136
200	0.0100	0.00279	3.754	0.00110	0.00037	1.65486
# C 12	0.00.0	0.00374	3.773	0.00327	0.00182	2.11409
515	0.00432	0.000.0	4.502	0.00381	0.00222	2.58765
575	0.00560	0.00000	334	0.00386	0.00150	0.73188
,	>	1 7000 0	2.370	0.00386	0.00150	1.55524

Table 24 IBM Reduced Data, P-10

	"(K02(CAT.)	KO2(CAT.) IN TWG SCREENS AT 10 1N.)	AT 10 1N.)			
TIME	62 GENERATED	CG2 ABSGRBED	H2G ABSGRBED			
(MINUTES)	(NECEN)	(SCFM)	(GR/MIN)			
15.	0.01865	0.00452	13.094			
45.	0.01909	0.00492	8.842			
75.	0.01907	0.00462	9.428			
35.	0.01907	64500*0	049.6		-	
195.	0.01687	0.00478	5.633			
255.	0.01687	0.00478	7.704			
315.	0.01687	0.00450	5.198			
375.	0.01602	0.00450	6.793			
55.	0.01518	0.00394	609-6			
5.	0.01352	0.00451	5.519			
55.	0.01352	0.00394	5.519			
615.	0.01267	0.00366	5.872			
	0.01174	0.00364	4.821			
735.	0.01018	0.00368	3.991			
795.	0.01014	0.00366	6.222			
855.	0.01014	0.00338	407			
ų.	0.00930	0.00338	6.222			
975.	0.00930	0.00338	5.171			
1035.	0.00846	0,00338	6.247			
1095.	0.00848	0.00368	7.341			
55.	0.00843	0.00309	3.825			
1215.	0.00761	0.00310	3.502			
ď.	0.00770	0.00314	3.943			
335.	0.00770	0.00314	3.563			
	0.00513	0.00313	3.977			
.55	0.00599	0.00314	2.896			
5.	0.00598	0.00256	2,582			
1575。	0.00513	0.00257	3,980			
635.	0.00599	0.00257	3.647			
5.	0.00513	0.00257	2.230			

Table 24 (cont.) IBM Reduced Data, P-10

1-23-1962		1-23-1962	
1-23-1962		1-23-1962	
DATE		DATE	
EVALUATION		EVALUATIGN	
OSITION CONTROL CANISTER )	H20 ABSORRED (GR/MIN) 12.852 10.520 11.263 8.5504 5.004 1.177	ISITION CONTROL	H2G ABSGRBED (GR/MIN) 18.067 12.934 14.323 14.325 11.845 11.472 6.004 2.368
KOZ ATMOSPHERIC COMPOSITION CONTROL EVALUATION CANISTER E (KOZ (CAT) IN 5 IN* CANISTER )	Cd2 ABSGRBED (SCFM) 0.00472 0.00554 0.00553 0.00531 0.00531	KØ2 ATMØSPHERIC COMPØSITIØN CØNTRØL EVALUATIØN CANISTER F (KØ2(CAT.) IN 20 IN. CANISTER)	CG2 ABSGRBED (SCFM) 0.00546 0.00564 0.00557 0.00560 0.00265
KG2 1	02 GENERATED (SCFM) 0.02034 0.01837 0.01748 0.01657 0.00994 0.00570	K G 2 A.	02 GENERATED (SCFM) 0.02241 0.02284 0.02174 0.02174 0.0226 0.0226 0.0226 0.0226 0.00836 0.00836
TEST P-10	FIME (MINUTES) 15. 45. 75. 135. 135. 255. 315. 375.	TEST P-10	TIME (MINUTES) 20. 50. 60. 140. 200. 260. 320.

	Ţ	Table 24 (cont.)		IBM Reduced Data,	P-10		
TEST P-10	KG2 AT	KG2 ATMOSPHERIC COMPOSITION CONTROL EVALUATION	ISITION CONTROL	EVALUATION	DATE	1-23-1962	1-24-1962
	(2/3K02+1/3	(2/3K02+1/3LIOH IN 5 IN CANISTER	INISTER )				
TIME	02 GENERATED	CGZ ABSGRBED	H20 ABSURBED	_			
(MINUTES)	(SCFM)	( SCF,M)	(GR/MIN)				
.5	0.03414	0.00555	154°4				
65.	0.01589	47400.0	8.088				
125.	0.01245	0.00470	7.356				
185.	0.01001	5 4400 0	4.377				
245.	0.00918	5 4400 0	7.396				
305	0.00918	0.00417	1.07				
365.	0.00839	0.00448	6.431				
425.	0.00071	8 4 4 00 0	\$ .594				
485.	0.00419	0.00419	B . 3				
545.	0.00168	0.00308	5.993				
605.	48000*0	0.00196	1.937				
665.	0.00140	0.00168	1.960				
TEST P-10	K02 AT	KOZ ĄTMOSPHERIC COMPOSITION CONTROL EVALUATION CANISTER MSA	SITION CONTROL	EVALUATION	DATE	1-23-1962	1-24-1962
	_		-				
171	AT CONCOATED	0.00.00.00.00.00.00.00.00.00.00.00.00.0					
	( SCC )	CUZ ABSURBED	HZU ABSUKBED				
20.	( SCTA)	(EL)()	(NIW/NI)				
•		0.00000	00000				
•	0.01298	0.00340	5.303				
	0.01505	0.0054	188.5				
• • • • • • • • • • • • • • • • • • • •	0.01205	0.00377	1 49° 4				
	A-100-0	0.00281	1.433				
.00Z	62.00.0	0.00238	2.771				
320	#1.700.0	0.00216	1.194				
280	0.00520	0.00217	0,143				
•00	0.00455	0.00152	1.227				
440	0.00453	0.00194	0.746				
.000	0.00450	0.00171	0.487				
620.	0.00387	0.00150	1.781				
.089	0.00195	0.00151	0.760				
740°	0.00263	0.00153	0.254				
800	0.00197	0.00153	1.537				
860.	0.00263	0.00110	1.698				
950	0.00197	0.00110	1.031				
•086	0.00132	0.00066	1,285				
1040	0.00131	0.00066	1.296				
.001	•	•0	2.579				

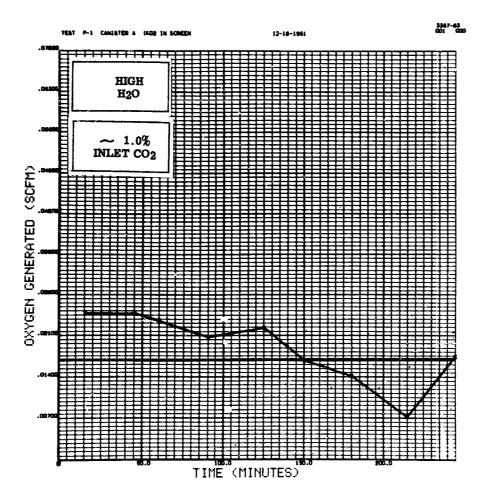


Figure 42 SC-4020 Curve, O2, P-1, A

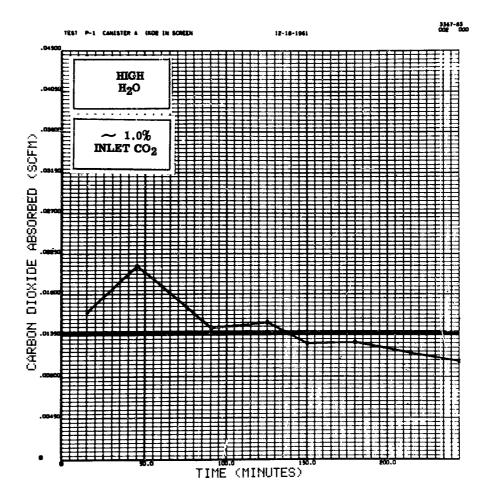


Figure 43 SC-4020 Curve, CO<sub>2</sub>, P-1, A

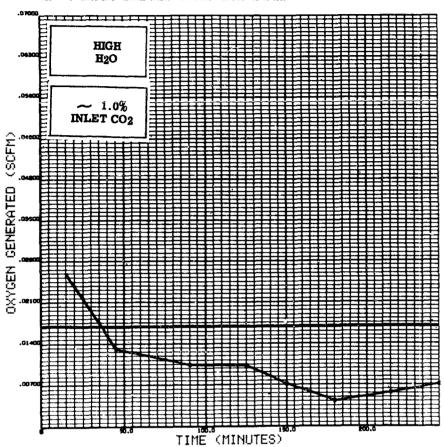
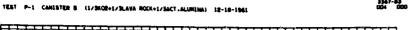


Figure 44 SC-4020 Curve, 0<sub>2</sub>, P-1, B



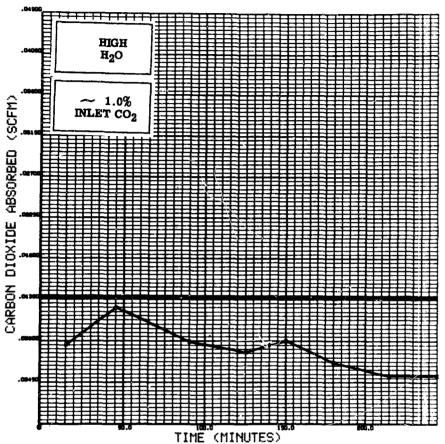


Figure 45 SC-4020 Curve, CO<sub>2</sub>, P-1, B



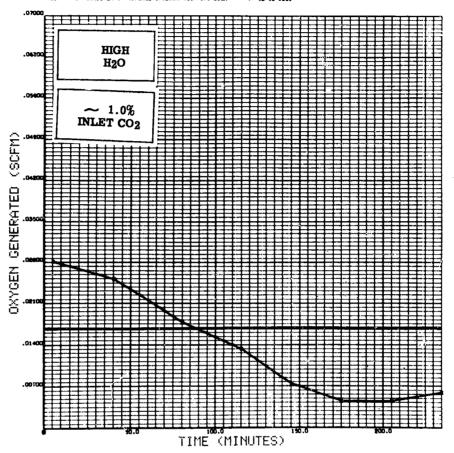


Figure 46 SC-4020 Curve, 02, P-1, C

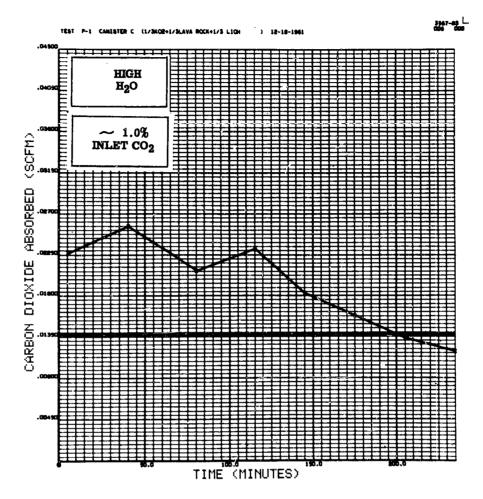


Figure 47 SC-4020 Curve, CO<sub>2</sub>, P-1, C

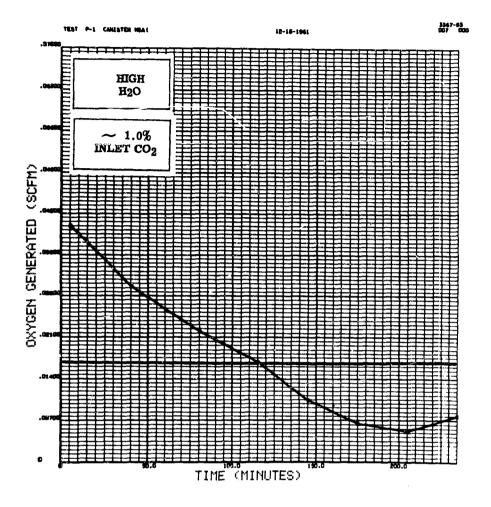


Figure 48 SC-4020 Curve, 0, P-1, MSA

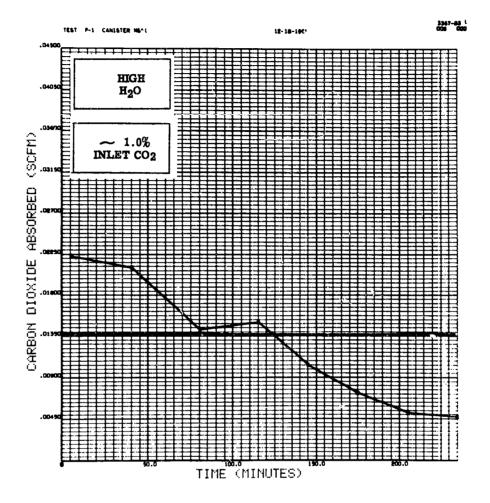


Figure 49 SC-4020 Curve, CO<sub>2</sub>, P-1, MSA

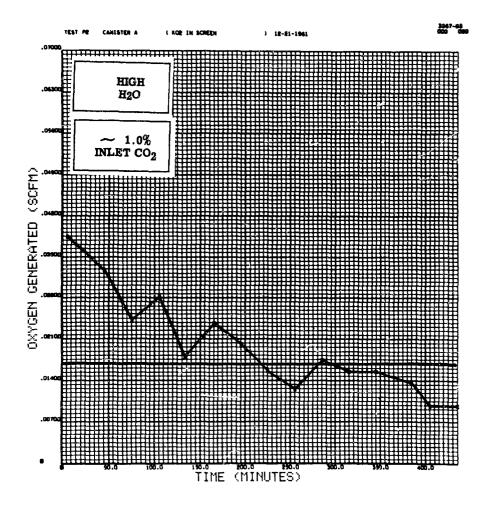


Figure 50 SC-4020 Curve, 0<sub>2</sub>, P-2, A

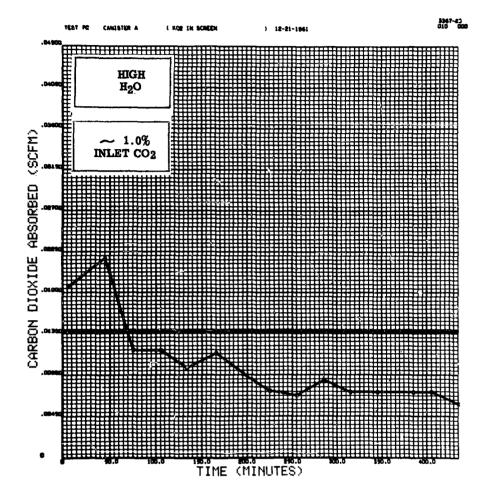


Figure 51 SC-4020 Curve, CO<sub>2</sub>, P-2, A

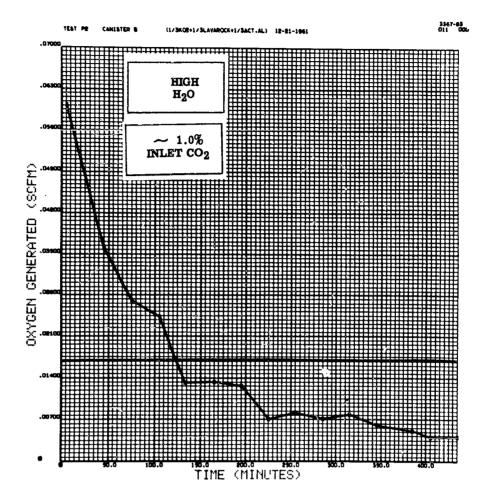


Figure 52 SC-4020 Curve, 0<sub>2</sub>, P-2, B

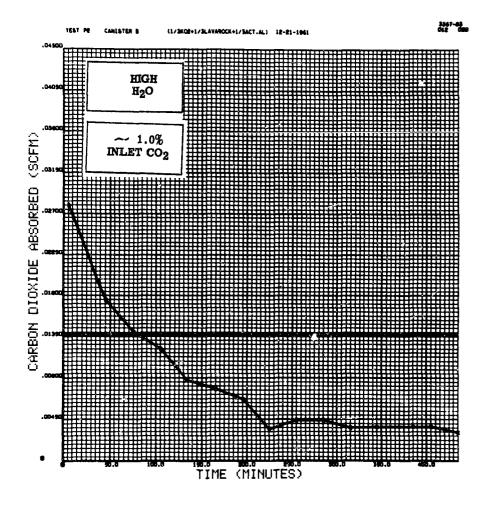


Figure 53 SC-4020 Curve, CO<sub>2</sub>, P-2, B

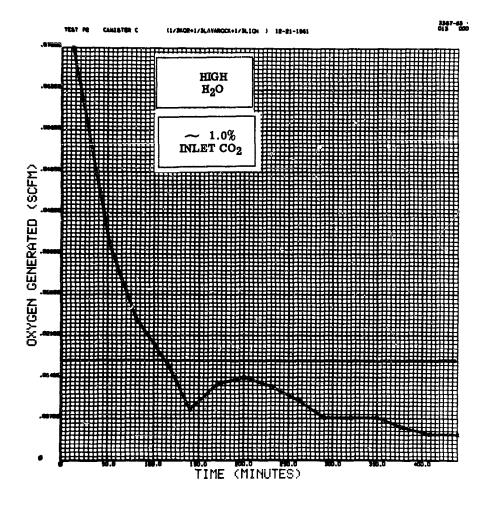


Figure 54 SC-4020 Curve, 0<sub>2</sub>, P-2, C

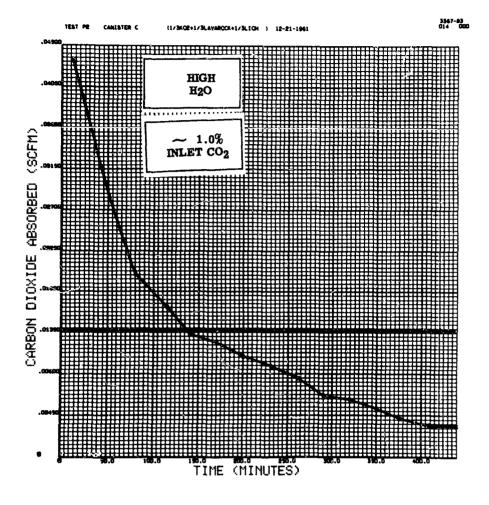


Figure 55 SC-4020 Curve, CO<sub>2</sub>, P-2, C

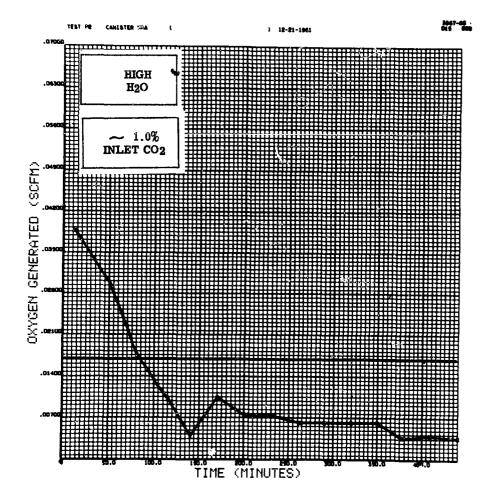


Figure 56 SC-4020 Curve, 02, P-2, MSA

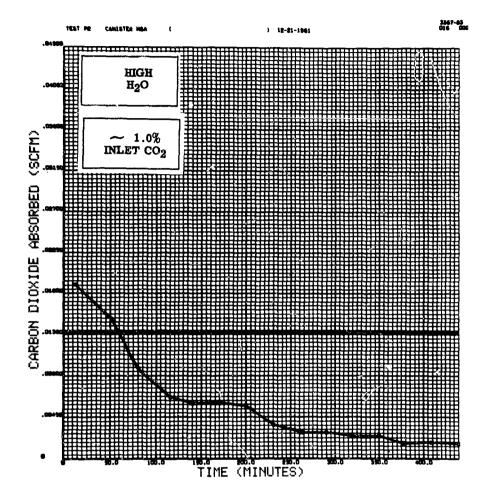


Figure 57 SC-4020 Curve, CO<sub>2</sub>, P-2, MSA

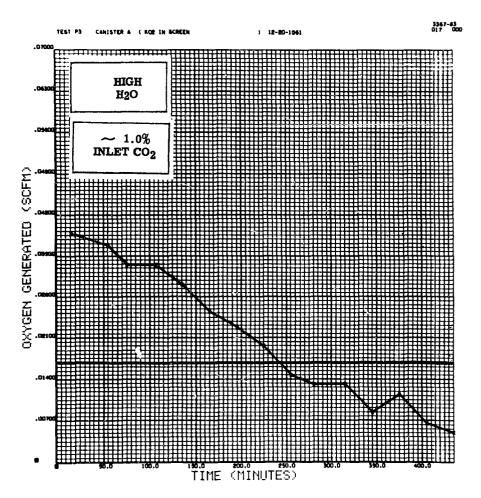


Figure 58 SC-4020 Curve, 0<sub>2</sub>, P-3, A

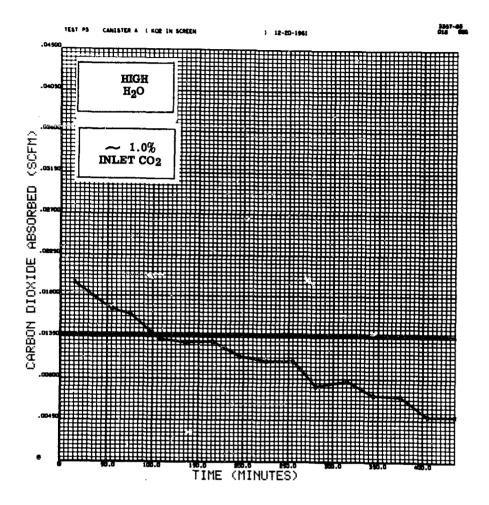


Figure 59 SC-4020 Curve, CO<sub>2</sub>, P-3, A

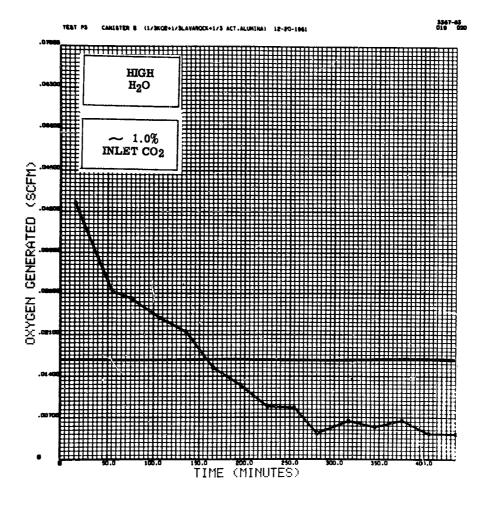


Figure 60 Sc-4020 Curve, 0<sub>2</sub>, P-3, B

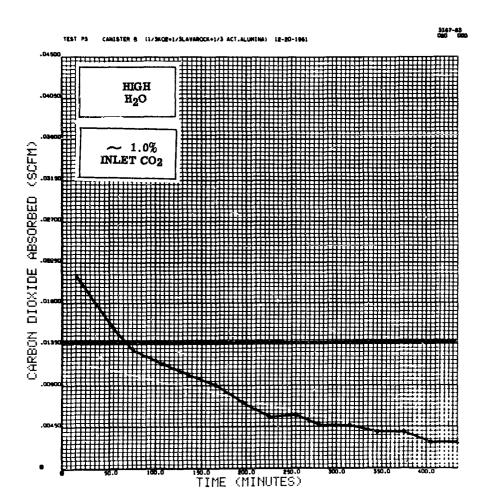


Figure 61 SC-4020 Curve, CO<sub>2</sub>, P-3, B

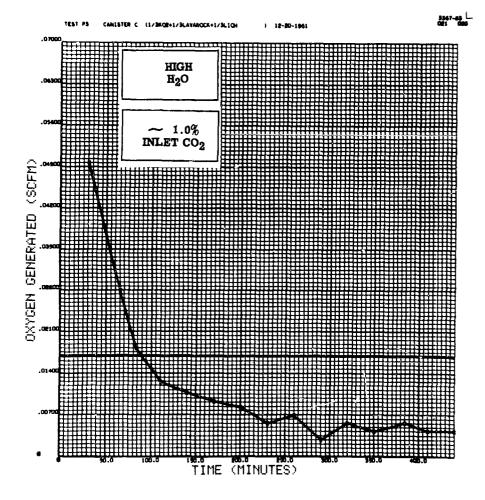


Figure 62 SC-4020 Curve, 0<sub>2</sub>, P-3, C

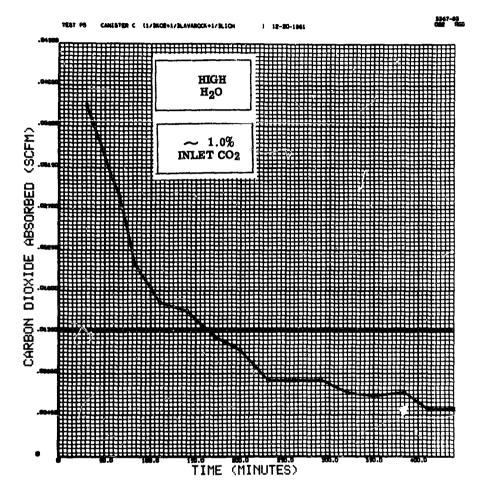


Figure 63 SC-4020 Curve, CO<sub>2</sub>, P-3, C

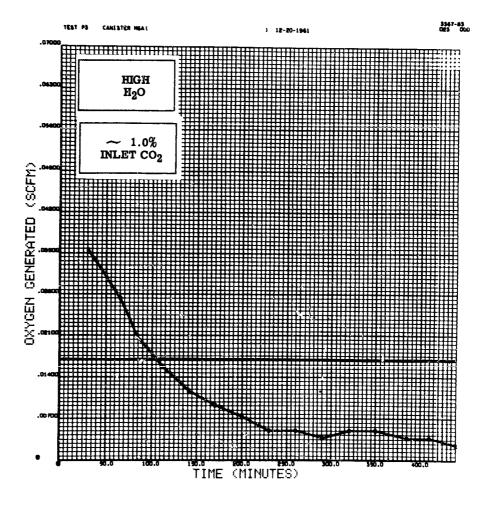


Figure 64 SC-4020 Curve, 0<sub>2</sub>, P-3, MSA

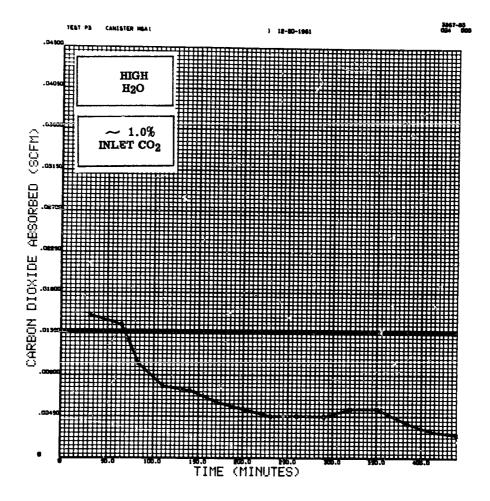


Figure 65 SC-4020 Curve, CO<sub>2</sub>, P-3, MSA

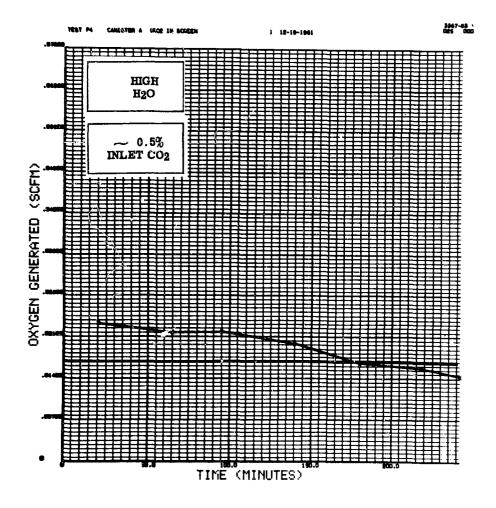


Figure 66 SC-4020 Curve, 0, P-4, A

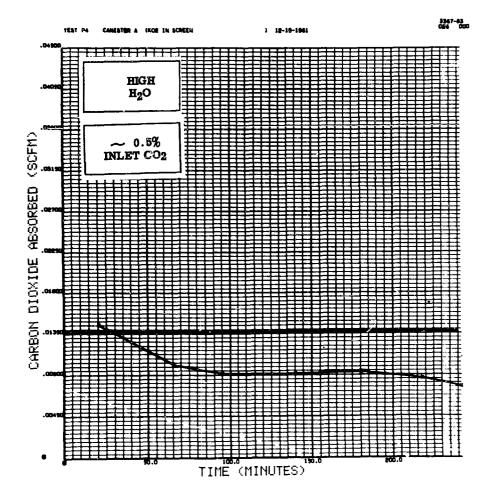


Figure 67 SC-4020 Curve, CO<sub>2</sub>, P-4, A

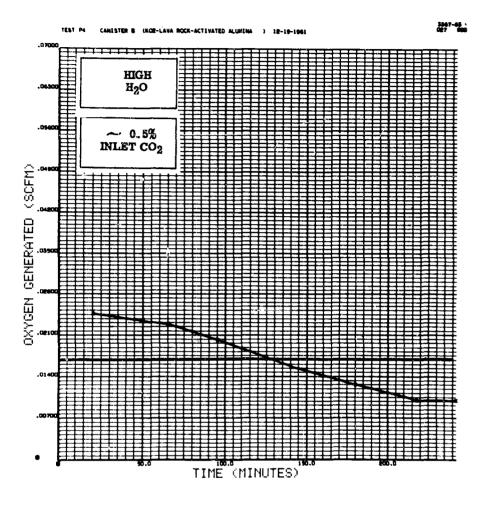


Figure 68 SC-4020 Curve, 0, P-4, B

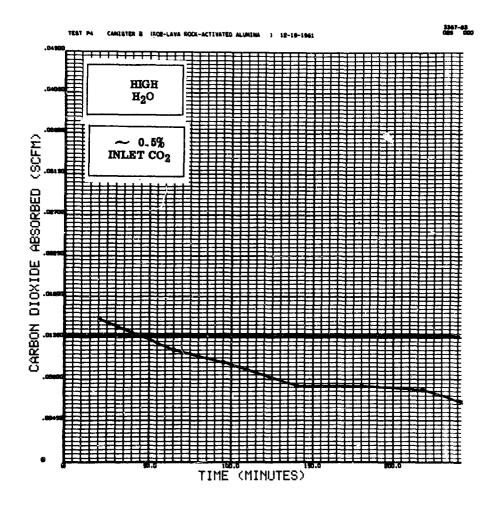


Figure 69 SC-4020 Curve, CO, P-4, B

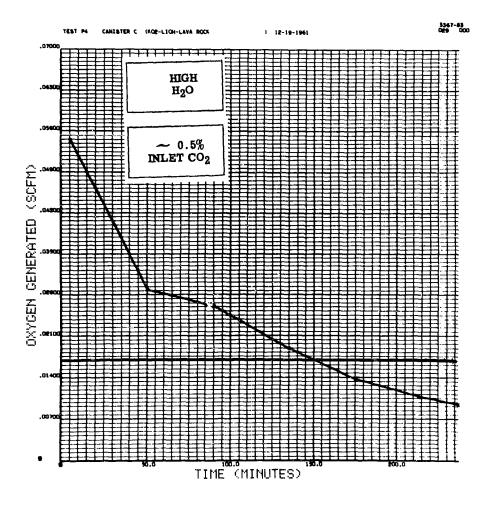


Figure 70 SC-4020 Curve, 0<sub>2</sub>, P-4, C

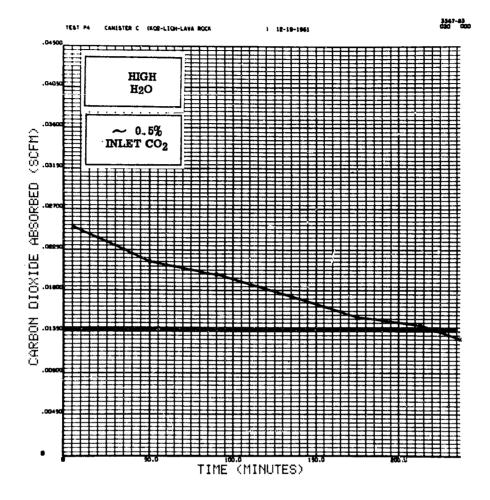


Figure 71 SC-4020 Curve, CO, P-4, C

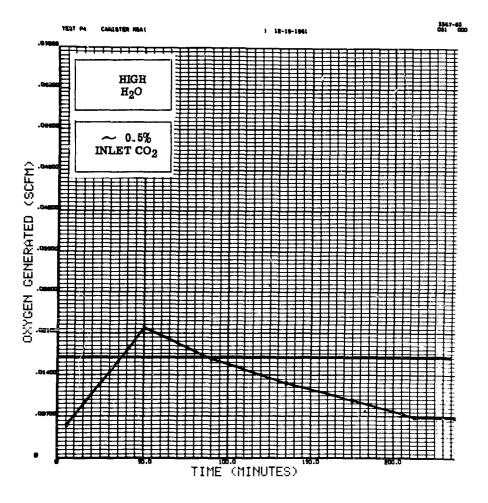


Figure 72 SC-4020 Curve, 0<sub>2</sub>, P-4, MSA

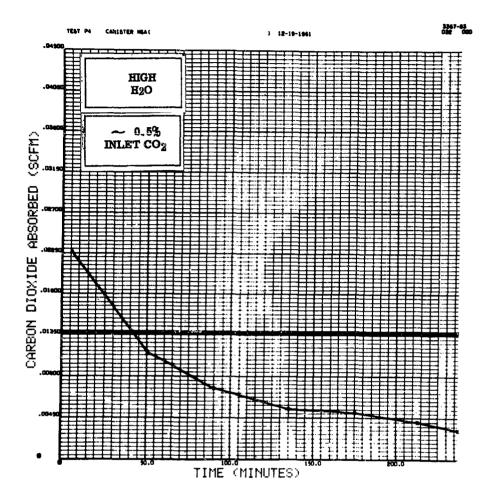


Figure 73 Sc-4020 Curve, CO<sub>2</sub>. P-4, MSA

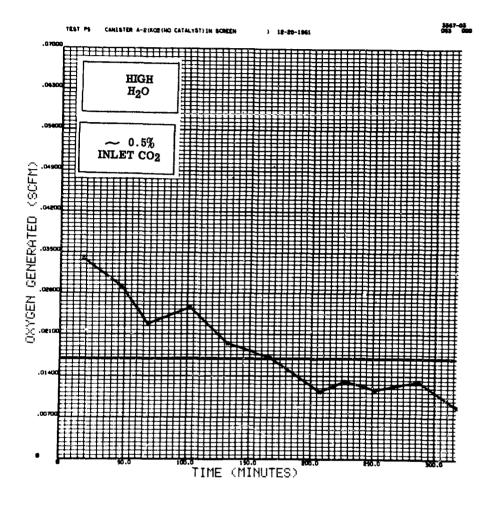


Figure 74 SC-4020 Curve, 0<sub>2</sub>, P-5, A-2

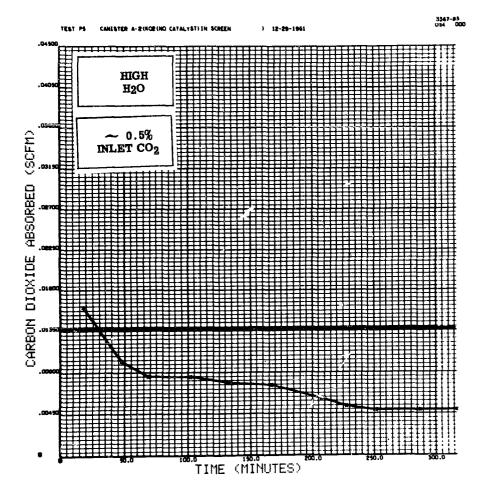
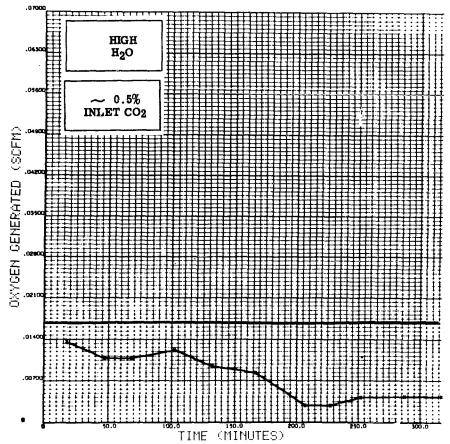


Figure 75 SC-4020 Curve, CO<sub>2</sub>, P-5, A-2





CANISTER A-3(KOR-LAVA ROCK-MS X-13 IN SCREEN ) 12-29-1961

Figure 76 SC-4020 Curve, 0<sub>2</sub>, P-5, A-3



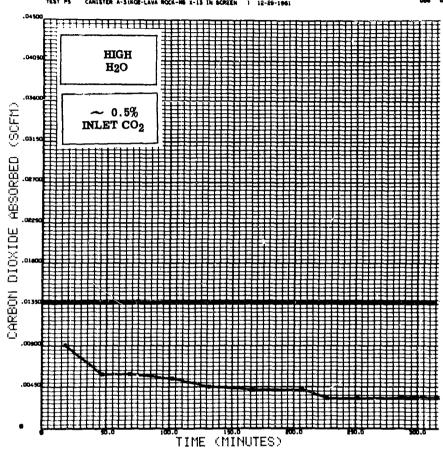


Figure 77 SC-4020 Curve, CO<sub>2</sub>, P-5, A-3

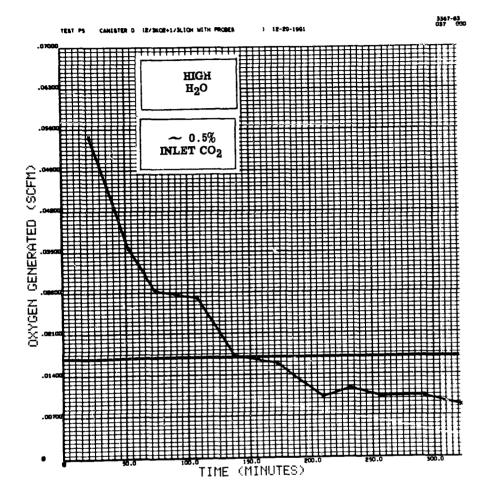


Figure 78 SC-4020 Curve, 0<sub>2</sub>, P-5, D

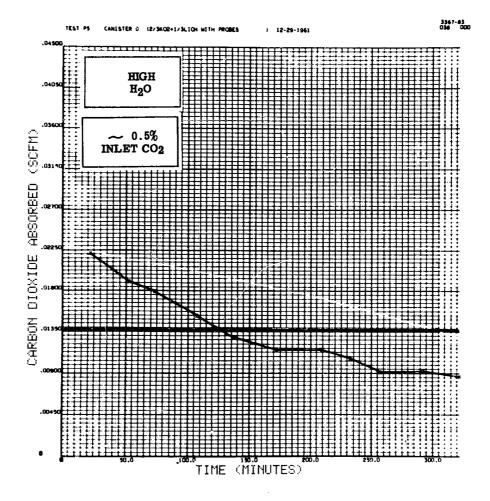


Figure 79 SC-4020 Curve, CO<sub>2</sub>, P-5, D

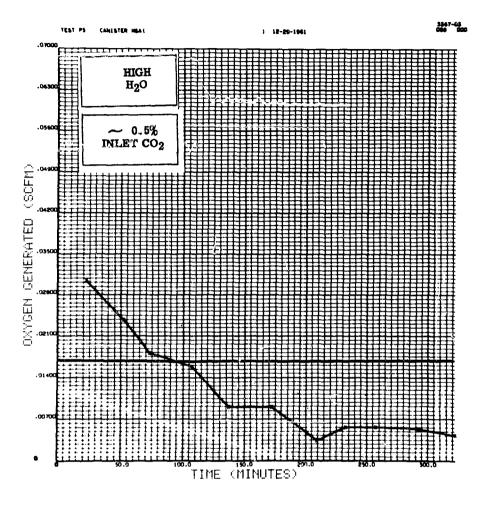


Figure 80 SC-4020 Curve, 02, P-5, MSA

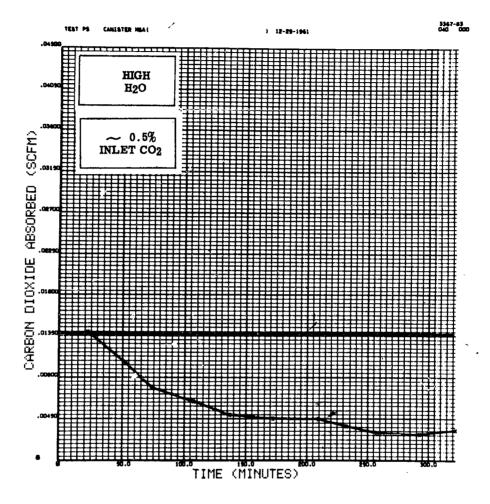


Figure 81 SC-4020 Curve, CO<sub>2</sub>, P-5, MSA

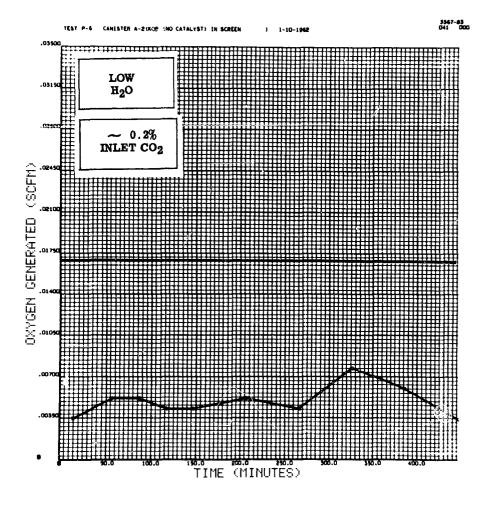


Figure 82 SC-4020 Curve, 0<sub>2</sub>, P-6, A-2

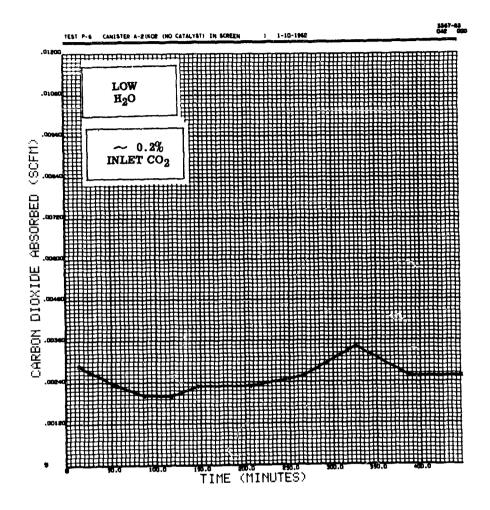


Figure 83 SC-4020 Curve, CO, P-6, A-2

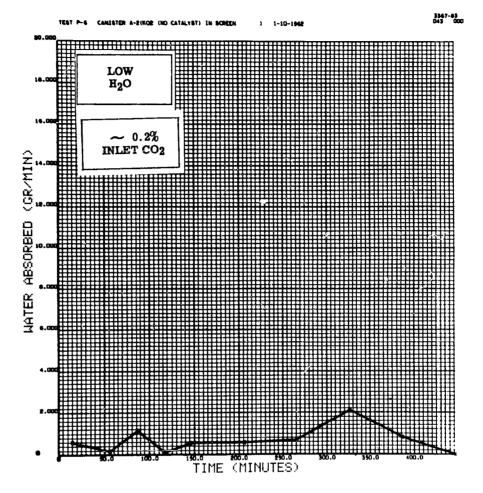


Figure 84 SC-4020 Curve, H<sub>2</sub>0, P-6, A-2

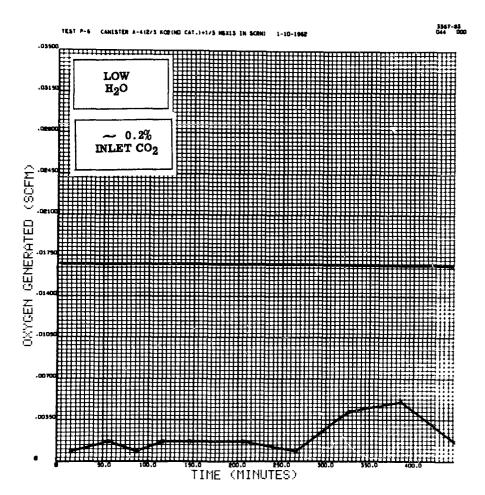


Figure 85 SC-4020 Curve, 0, P-6, A-4

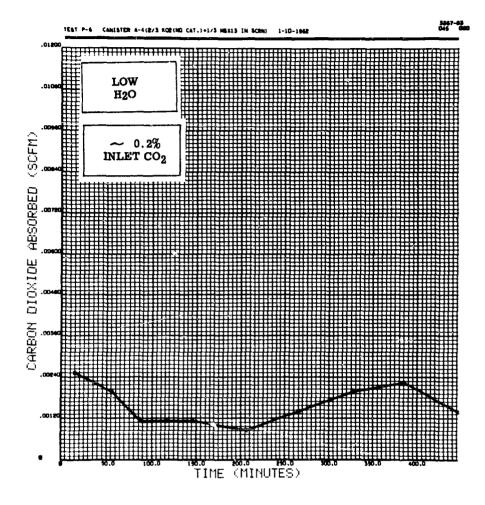


Figure 86 SC-4020 Curve, CO<sub>2</sub>, P-6, A-4

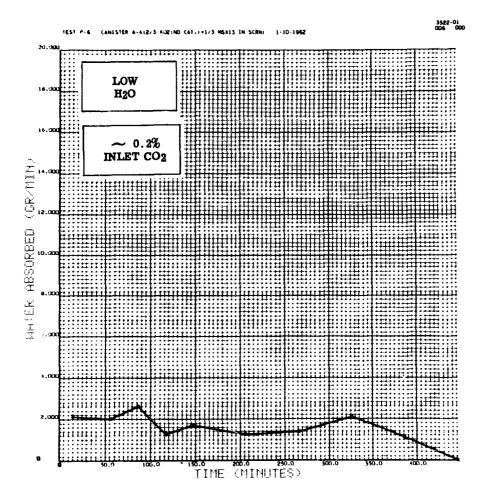


Figure 87 SC-4020 Curve, H<sub>2</sub>0, P-6, A-4

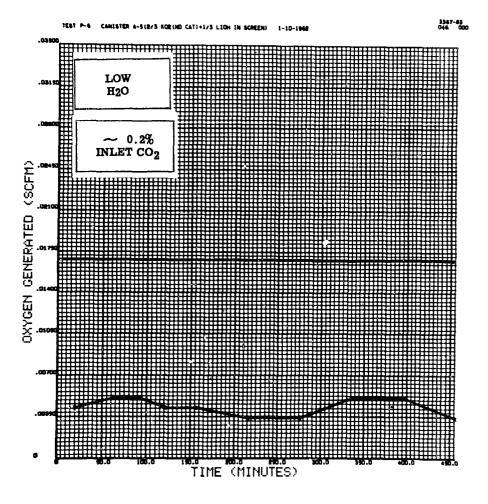


Figure 88 SC-4020 Curve, 02, P-6, A-5

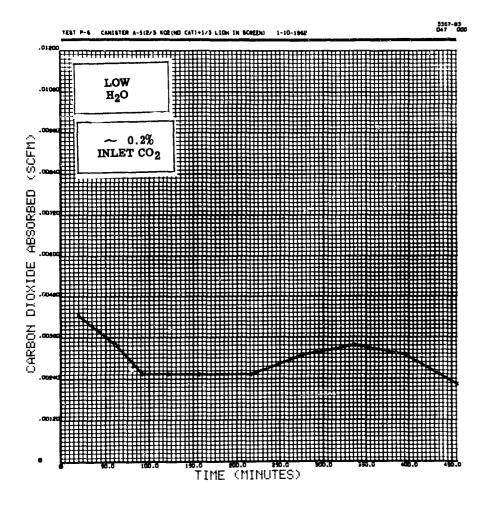


Figure 89 SC-4020 Curve, CO<sub>2</sub>, P-6, A-5

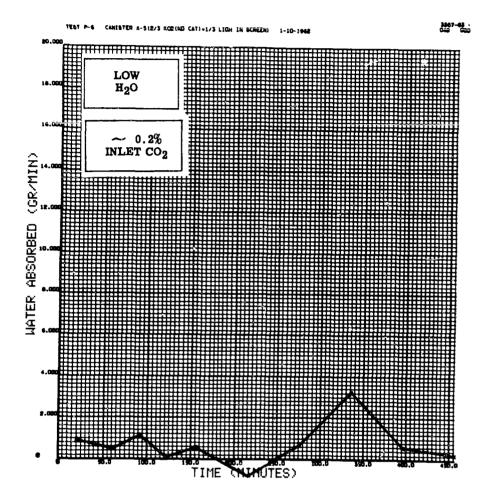


Figure 90 SC-4020 Curve, H<sub>0</sub>0, P-6, A-5

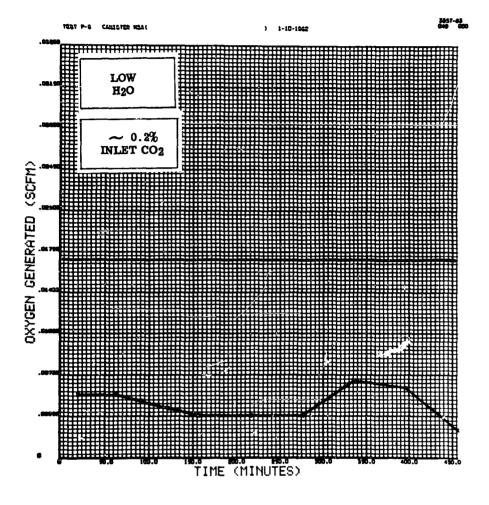


Figure 91 SC-4020 Curve, 0, P-6, MSA

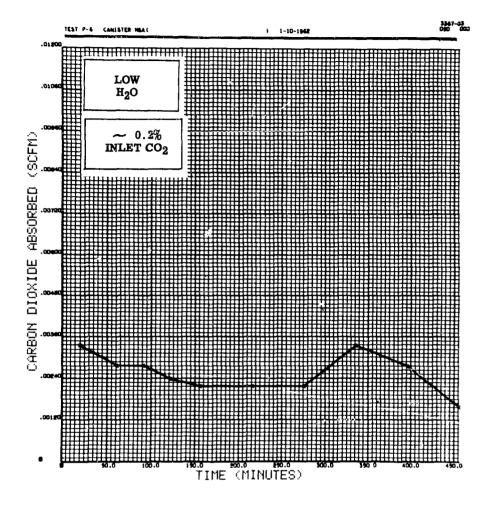


Figure 92 SC-4020 Curve, CO, P-6, MSA

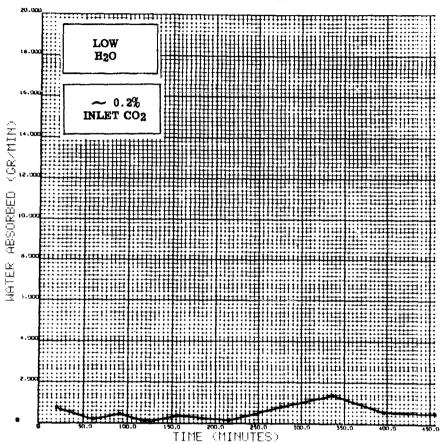


Figure 93 SC-4020 Curve, H 0, P-6, MSA

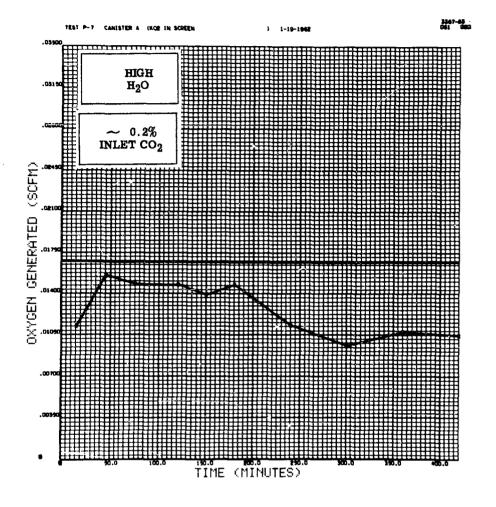


Figure 94 SC-4020 Curve, 0, P-7, A

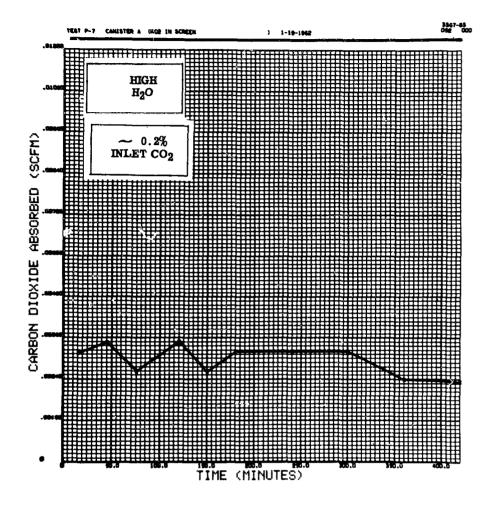


Figure 95 SC-4020 Curve, CO, P-7, A

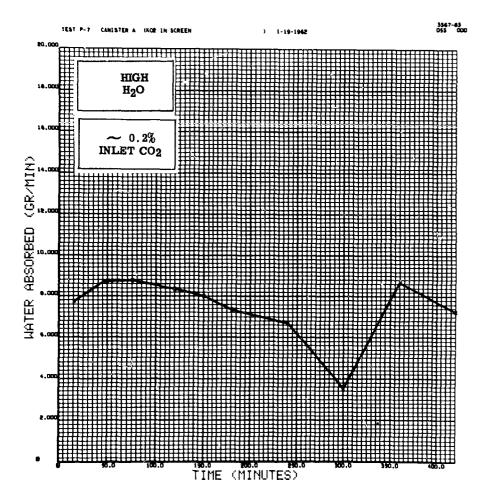


Figure 96 SC-4020 Curve, H<sub>2</sub>0, P-7, A





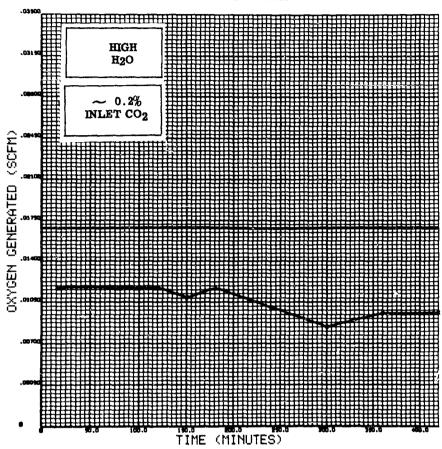


Figure 97 SC-4020 Curve, 0<sub>2</sub>, P-7, A-6

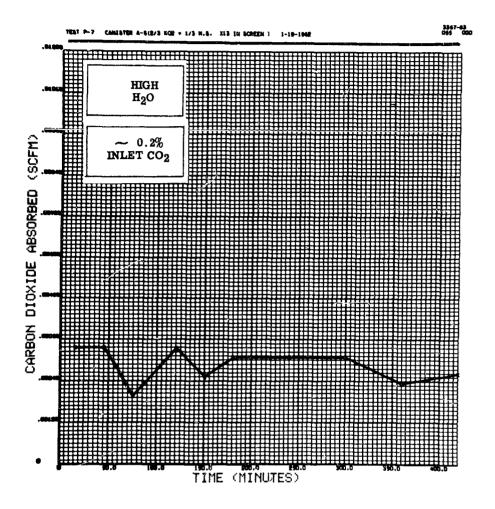


Figure 98 SC-4020 Curve, CO<sub>2</sub>, P-7, A-6

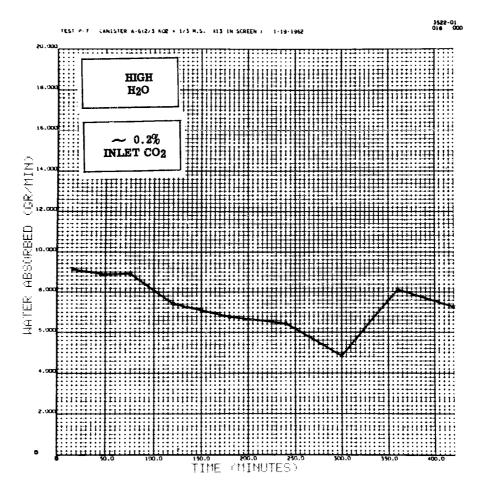


Figure 99 SC-4020 Curve, H<sub>2</sub>0, P-7, A-6

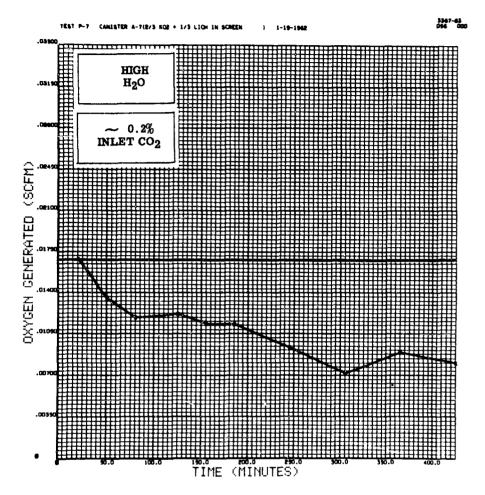


Figure 100 SC-4020 Curve, 0, P-7, A-7

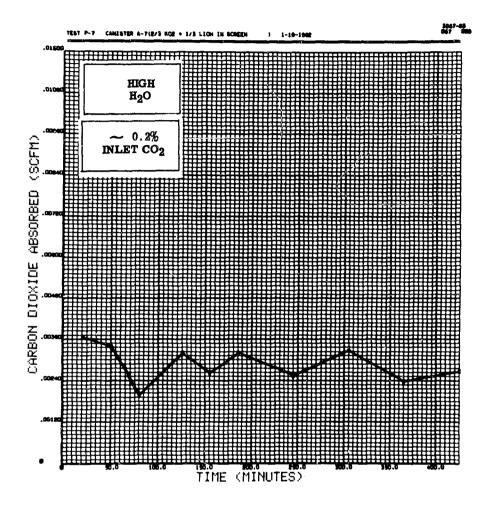


Figure 101 SC-4020 Curve, CO<sub>2</sub>, P-7, A-7

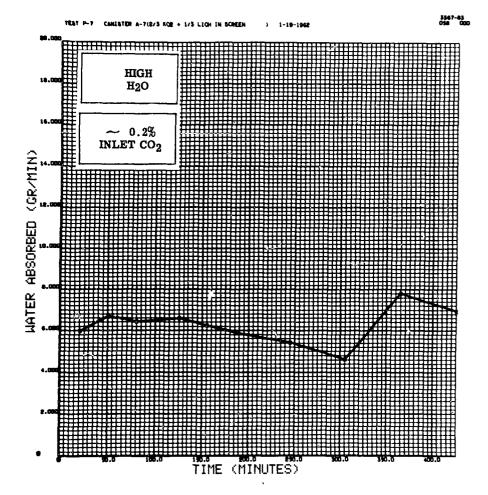


Figure 102 SC-4020 Curve, H<sub>2</sub>O, P-7, A-7

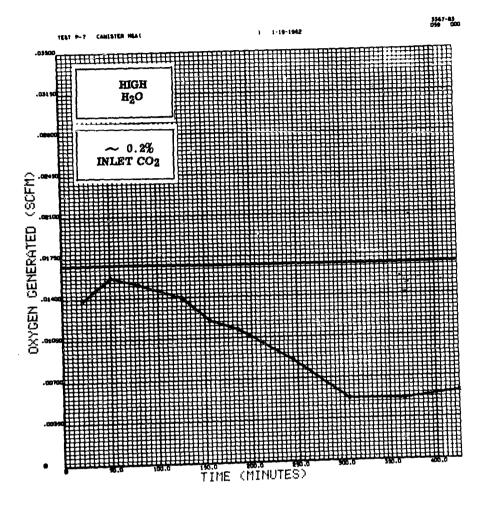


Figure 103 SC-4020 Curve, 02, P-7, MSA

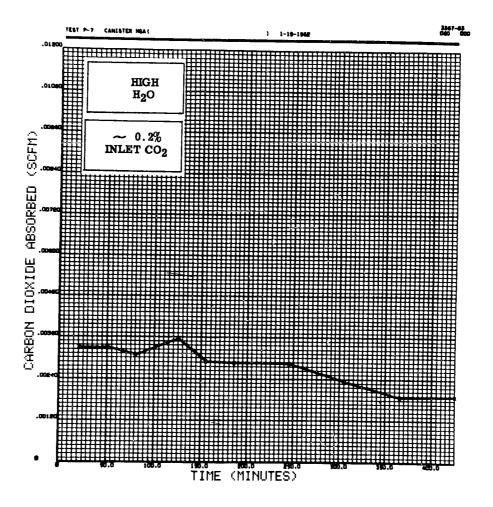


Figure 104 SC-4020 Curve, CO, P-7, MSA

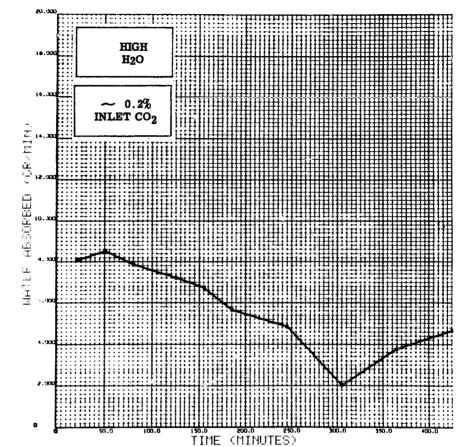


Figure 105 SC-4020 Curve, H O, P-7, MSA

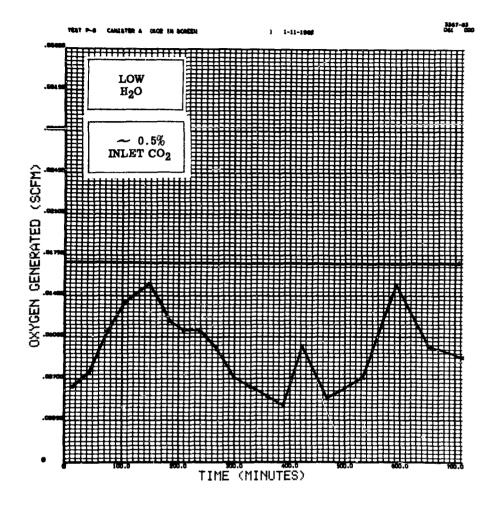


Figure 106 SC-4020 Curve, 0<sub>2</sub>, P-8, A

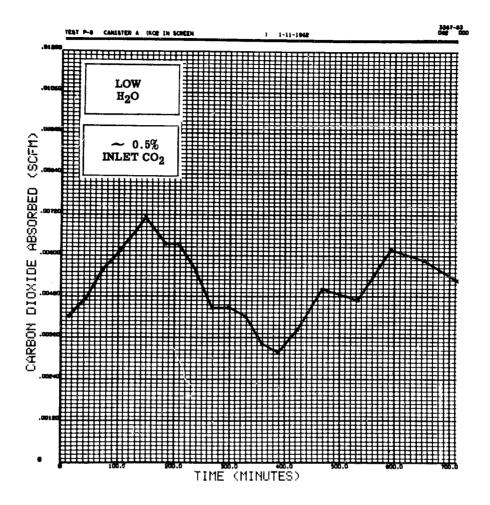


Figure 107 SC-4020 Curve, CO<sub>2</sub>, P-8, A

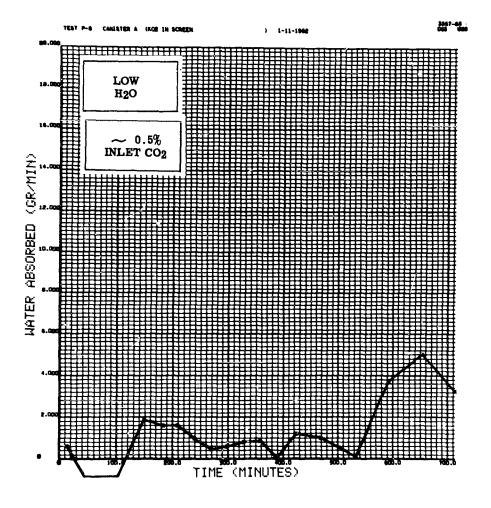


Figure 108 SC-4020 Curve, H<sub>0</sub>, P-8, A

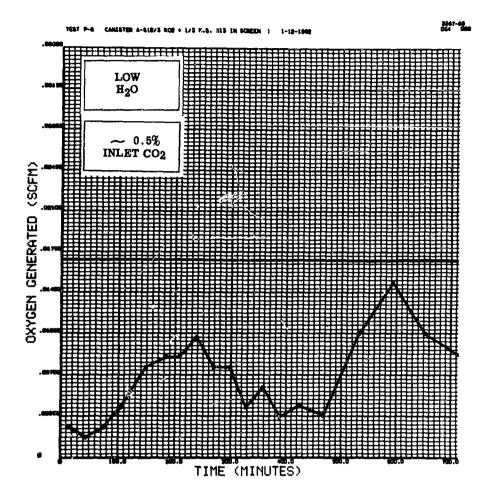


Figure 109 SC-4020 Curve, 02, P-8, A-6

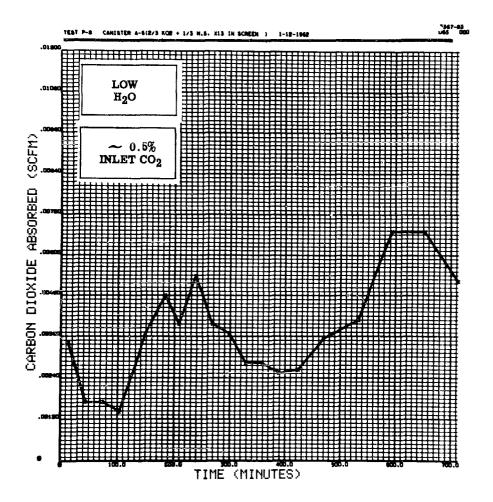
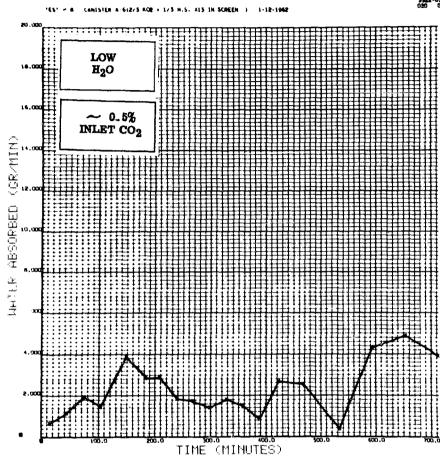


Figure 110 SC-4020 Curve, CO, P-8, A-6





SC-4020 Curve, H<sub>2</sub>O, P-8, A-6 Figure 111

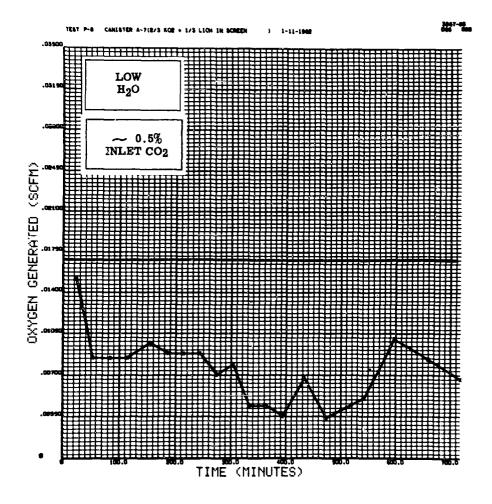


Figure 112 SC-4020 Curve, 0<sub>2</sub>, P-8, A-7

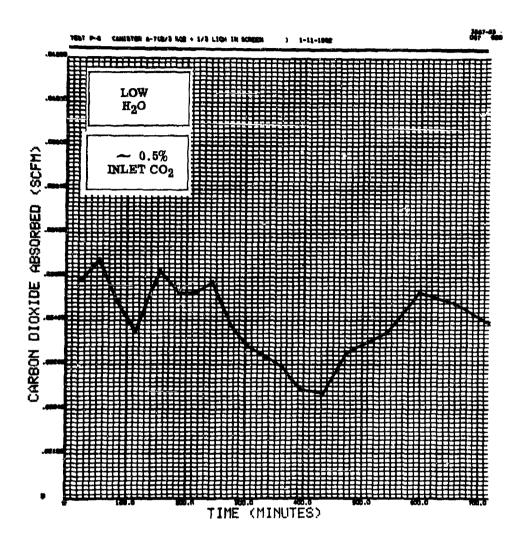


Figure 113 SC-4020 Curve, CO<sub>2</sub>, P-8, A-7

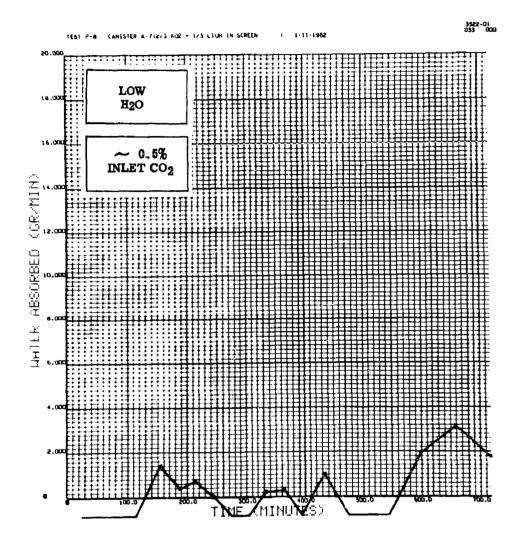


Figure 114 SC-4020 Curve, H<sub>2</sub>0, P-8, A-7

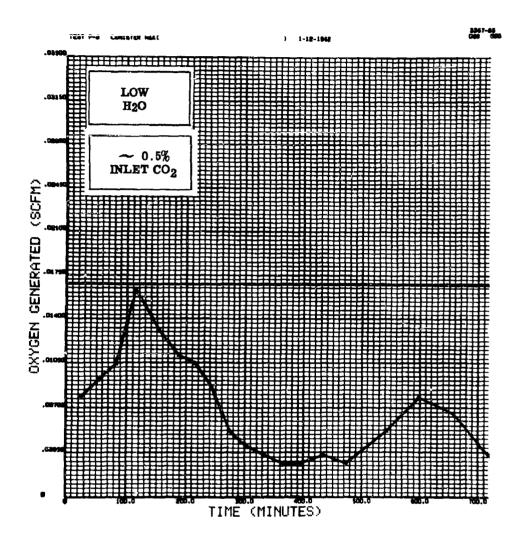


Figure 115 SC-4020 Curve, 0, P-8, MSA

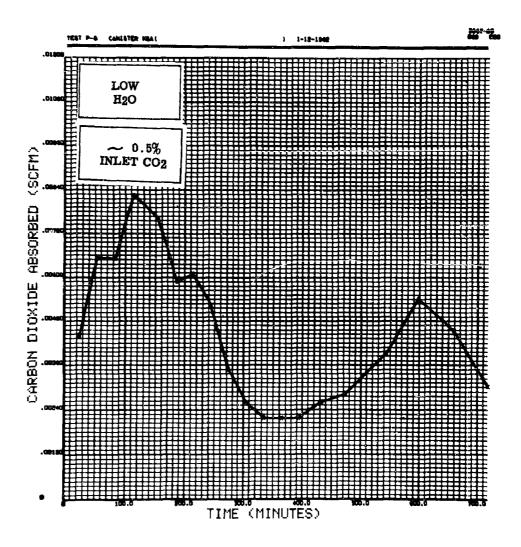


Figure 116 SC-4020 Curve, CO<sub>2</sub>, P-8, MSA

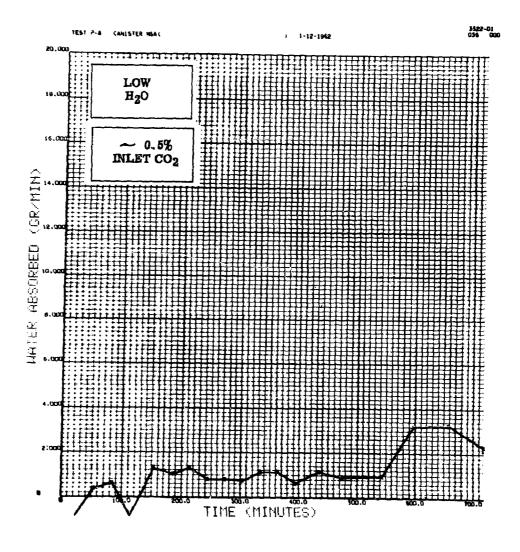


Figure 117 SC-4020 Curve, H<sub>2</sub>0, P-8, MSA

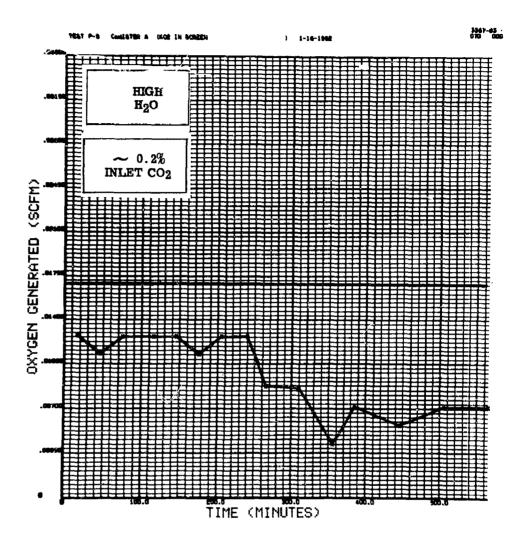


Figure 118 SC-4020 Curve, 0, P-9, A

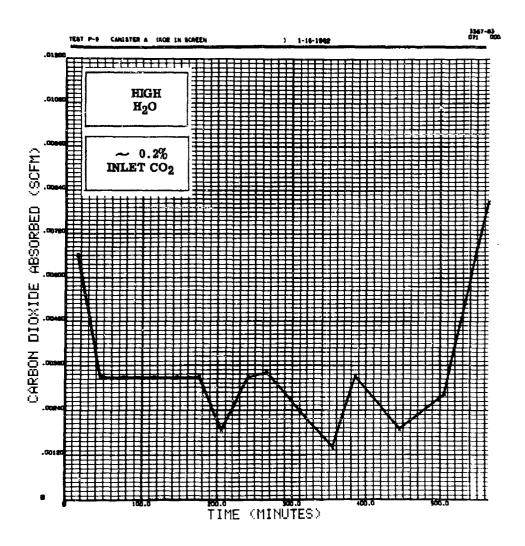


Figure 119 SC-4020 Curve, CO<sub>2</sub>, P-9, A

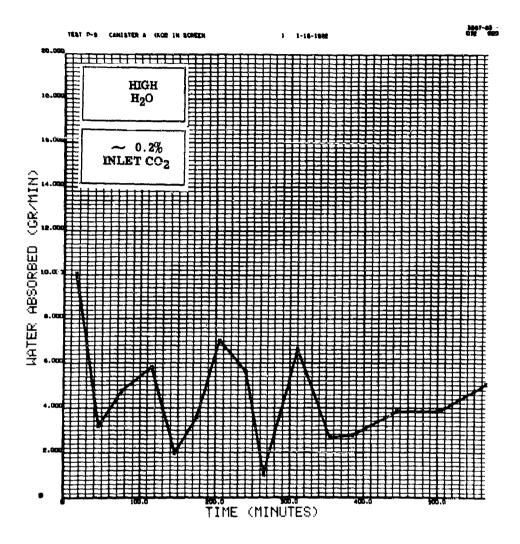


Figure 120 SC-4020 Curve, H 0, P-9, A

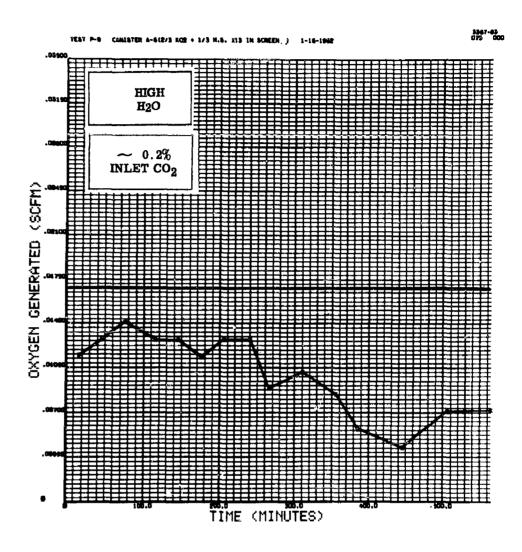


Figure 121 SC-4020 Curve, 0, P-9, A-6

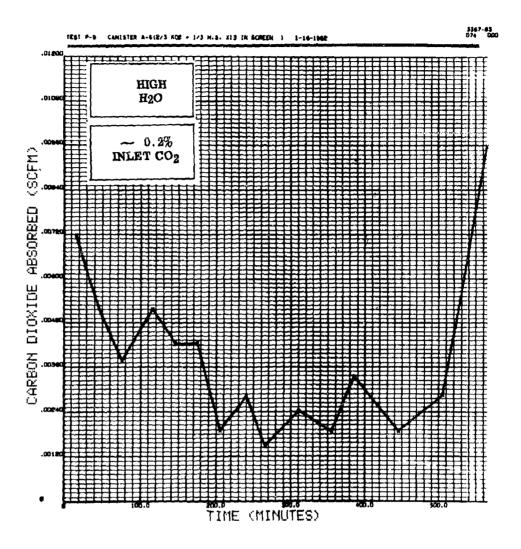


Figure 122 SC-4020 Curve, CO<sub>2</sub>, P-9, A-6





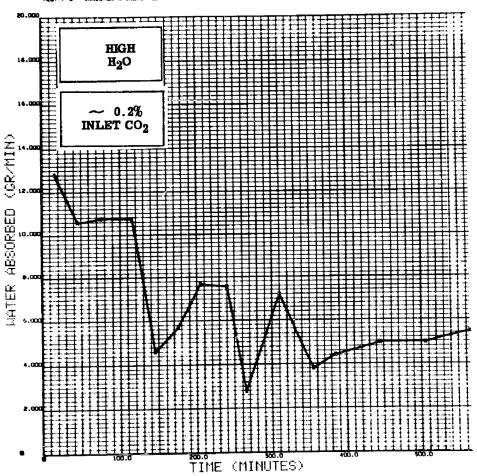


Figure 123 SC-4020 Curve, H O, P-9, A-6

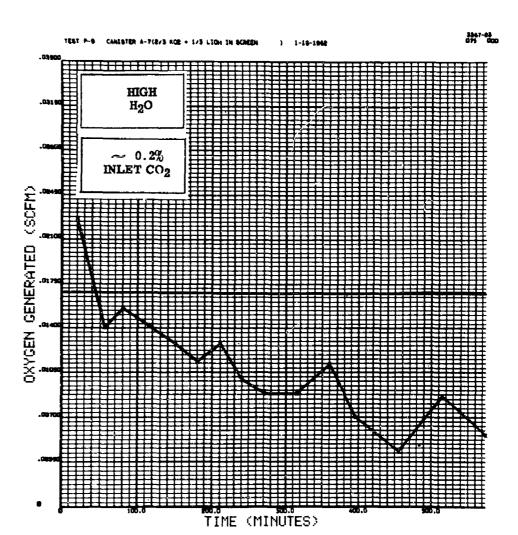


Figure 124 SC-4020 Curve, 0<sub>2</sub>, P-9, A-7

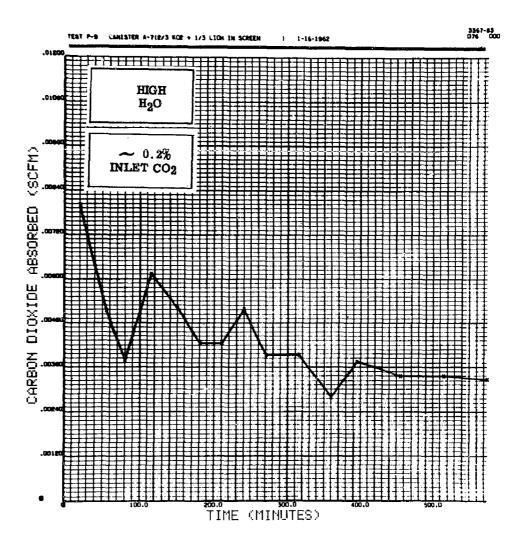


Figure 125 SC-4020 Curve, CO , P-9, A-7

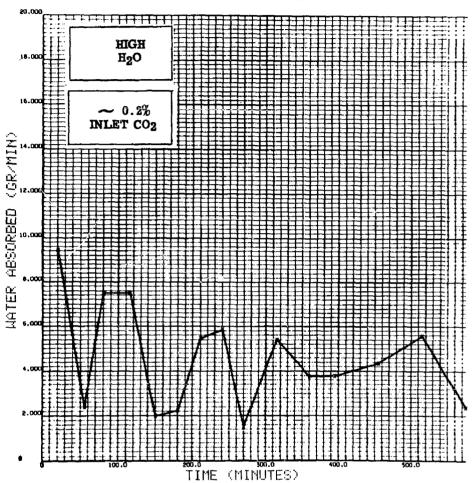


Figure 126 SC-4020 Curve, H<sub>0</sub>, P-9, A-7

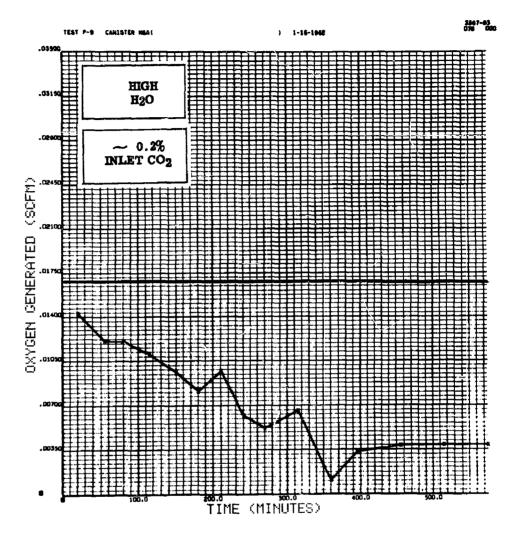


Figure 127 SC-4020 Curve, 0, P-9, MSA

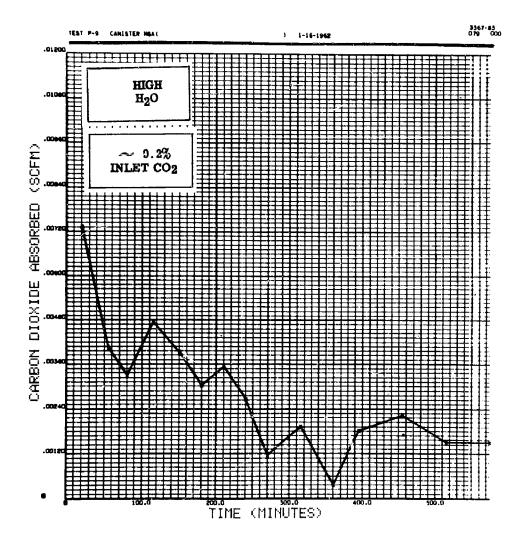


Figure 128 SC-4020 Curve, CO , P-9, MSA

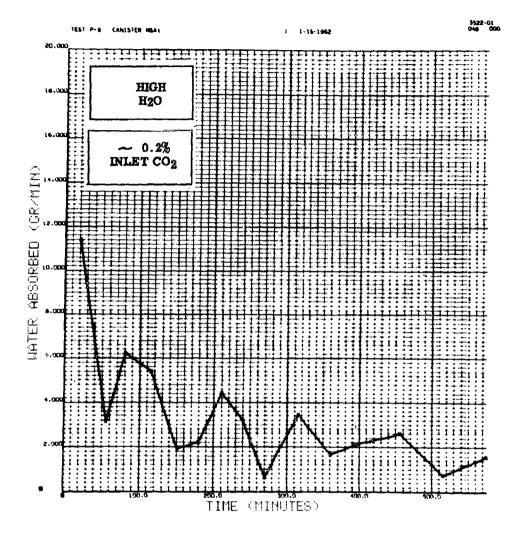


Figure 129 SC-4020 Curve, H<sub>0</sub>0, P-9, MSA

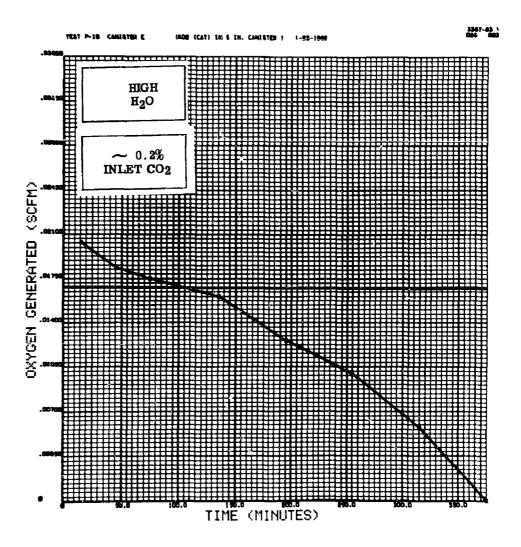


Figure 130 SC-4020 Curve, 0, P-10, E

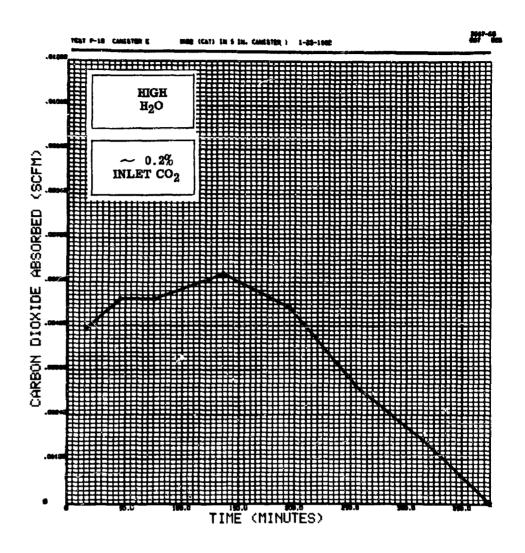


Figure 131 SC-4020 Curve, CO<sub>2</sub>, P-10, E

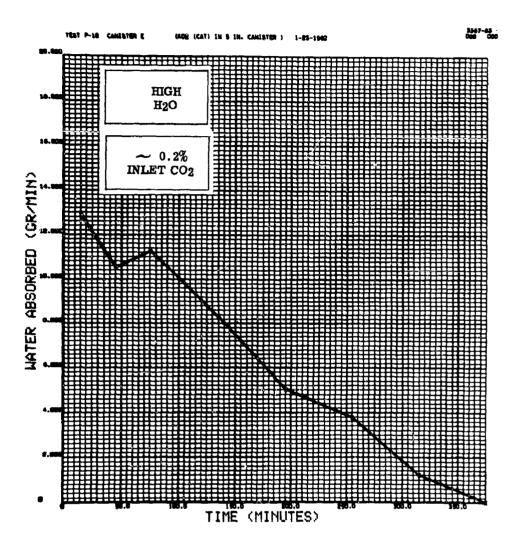


Figure 132 SC-4020 Curve, H<sub>2</sub>0, P-10, E



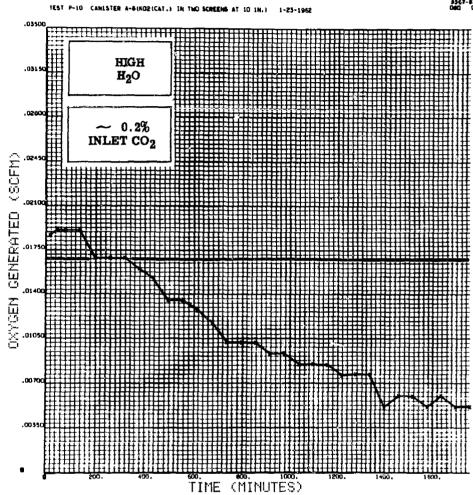


Figure 133 SC-4020 Curve, 0, P-10, A-8



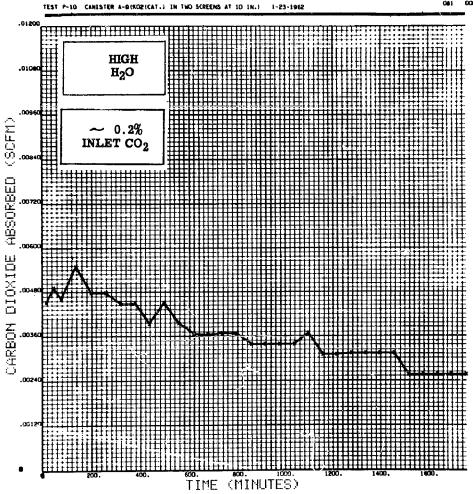


Figure 134 SC-4020 Curve, CO<sub>2</sub>, P-10, A-8

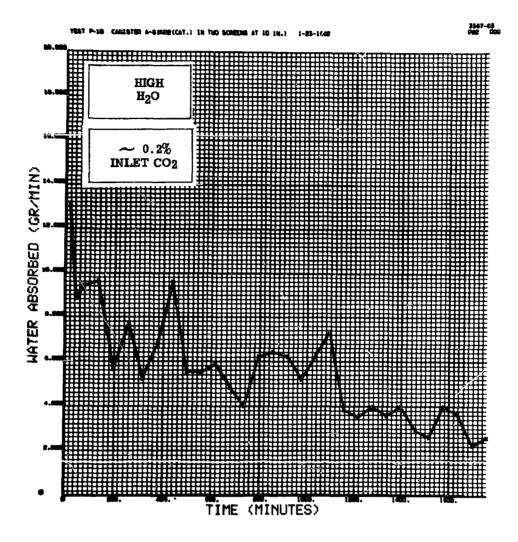


Figure 135 SC-4020 Curve, H<sub>2</sub>0, P-10, A-8

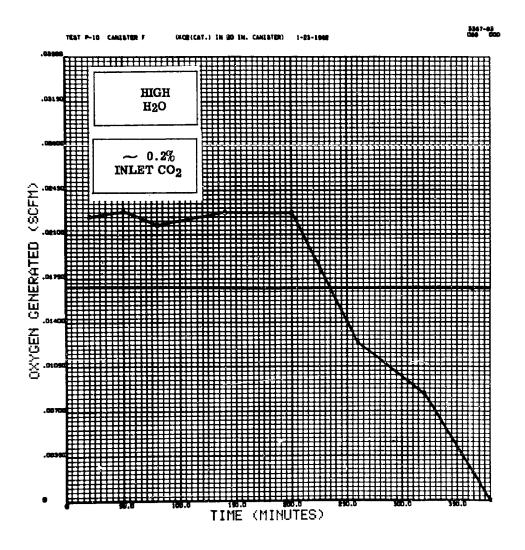


Figure 136 SC-4020 Curve, 0<sub>2</sub>, P-10, F

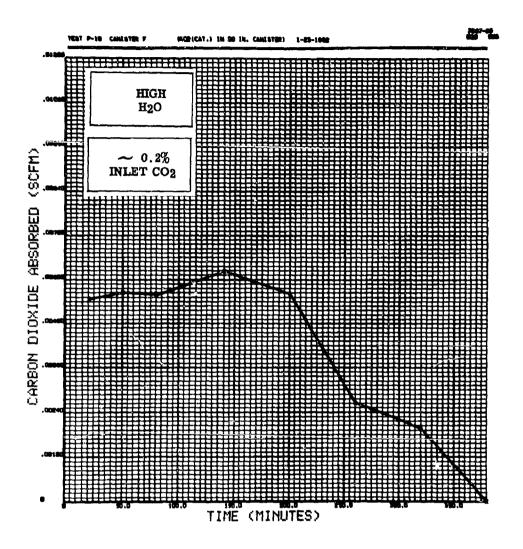


Figure 137 SC-4020 Curve, CO<sub>2</sub>, P-10, F

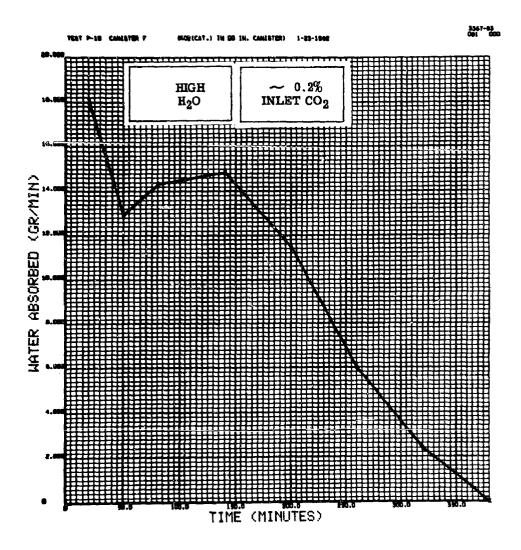


Figure 138 SC-4020 Curve, H<sub>2</sub>0, P-10, F

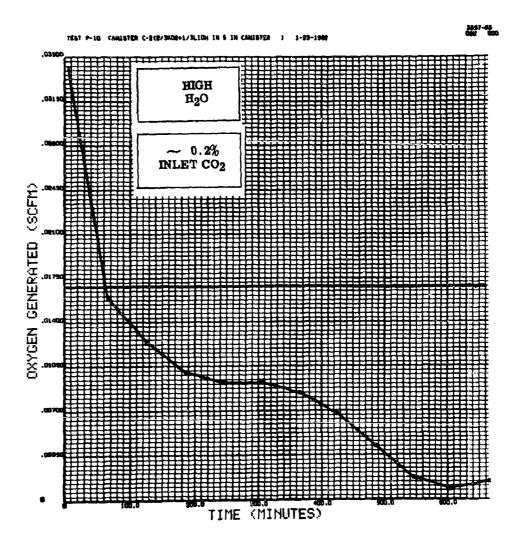


Figure 139 SC-4020 Curve, 0, P-10, C-2

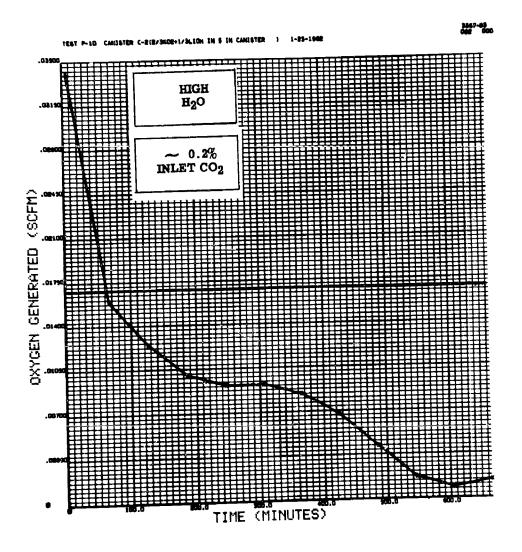


Figure 139 SC-4020 Curve, 0, P-10, C-2

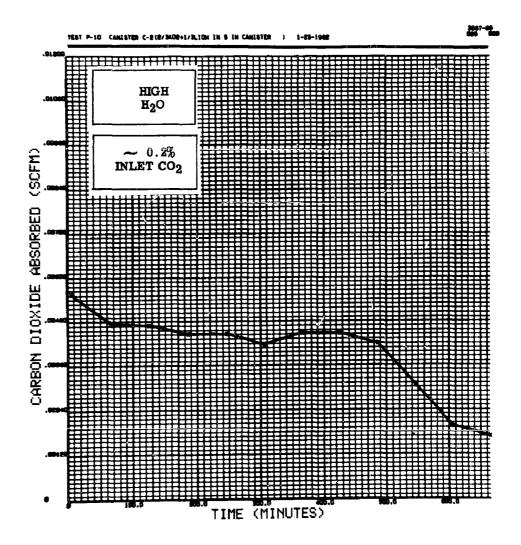


Figure 140 SC-4020 Curve, CO<sub>2</sub>, P-10, C-2

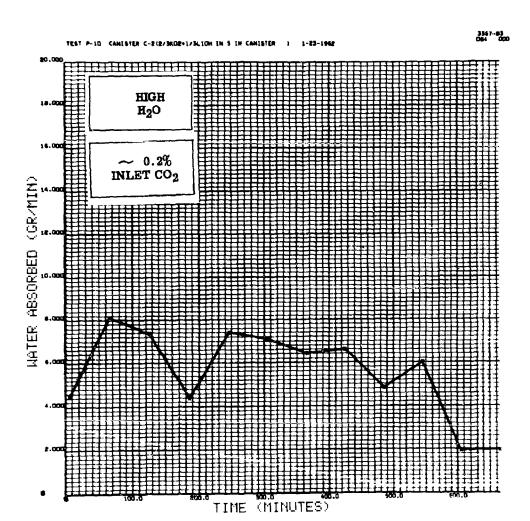


Figure 141 SC-4020 Curve, H<sub>2</sub>0, P-10, C-2

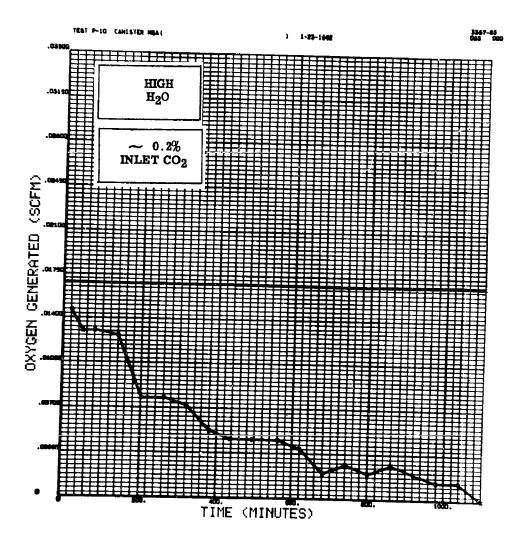


Figure 142 SC-4020 Curve, 0<sub>2</sub>, P-10, MSA

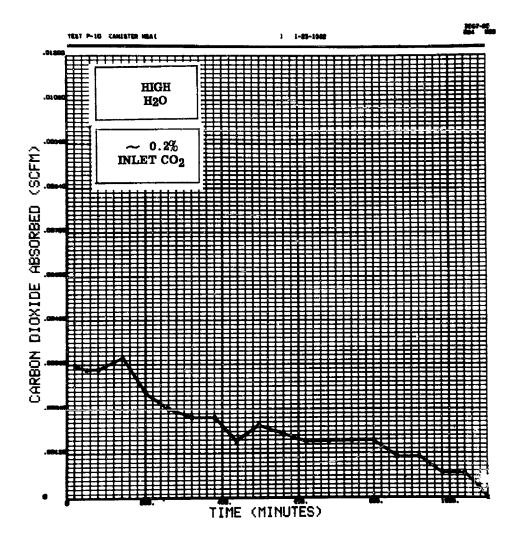


Figure 143 SC-4020 Curve, CO<sub>2</sub>, P-10, MSA

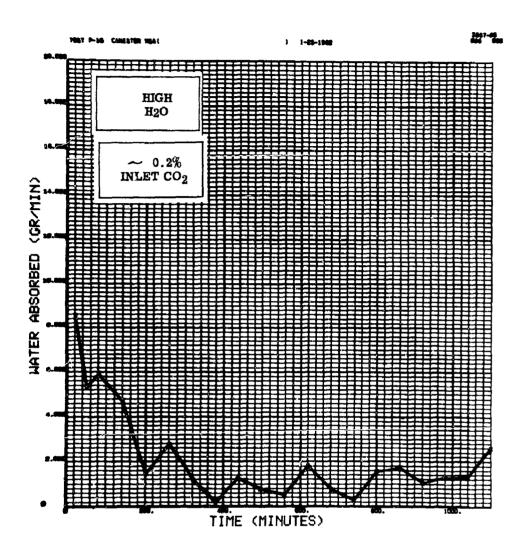


Figure 144 SC-4020 Curve, H<sub>2</sub>O, P-10, MSA

Table 25 Cumulative Data, P-1

Time Min.	OXYCE CONSUMED	en scf Generated	CARBON DI PRODUCED	OXIDE SCF ABSORBED	CUM. R.Q.
Canister A 0 15 45 90	0.25 0.76 1.51	0.19 0.93 1.95	0.21 0.62 1.24	0.12 0.68 1.47	0.65 0.73 0.75
125 150 180 215 245	1.51 2.10 2.52 3.02 3.61 4.11	1.955 2.71 3.19 3.655 4.03	1.72 2.06 2.48 2.96 3.37	1.98 2.33 2.70 3.13 3.46	0.73 0.73 0.74 0.78 0.79
Canister B 0 15 45 90 125 150 180 215 245	0.25 0.76 1.51 2.10 2.52 3.61 4.11	 0.19 0.77 1.30 1.66 1.87 2.04 2.22 2.41	0.21 0.62 1.24 1.72 2.06 2.48 2.96 3.37	0.06 0.38 0.86 1.15 1.50 1.81	0.33 0.49 0.67 0.70 0.78 0.81 0.81
Canister C 0 40 80 115 145 175 205 235	0.67 1.93 1.93 2.94 2.94 3.95	0.07 1.00 1.86 2.40 2.71 2.89 3.02 3.18	0.07 0.55 1.10 1.58 2.00 2.41 2.82 3.23	0.90 1.83 2.60 3.22 3.73 4.17	0.81 0.90 0.98 1.08 1.19 1.29 1.38 1.43
Canister MS	A				
0 5 40 80 115 175 205 235	0.67 1.33 1.933 2.94 3.95	0.10 1.32 2.35 3.02 3.68 3.68 3.85	0.07 0.55 1.10 1.58 2.00 2.41 2.82 3.23	0.06 0.81 1.52 2.03 2.41 2.67 2.86 3.01	0.56 0.62 0.65 0.67 0.70 0.73 0.74

Table 26 Cumulative Data, P-2

TIME MIN.	OXYG CONSUMED	en scf Generated	CARBON DI PRODUCED	OXIDE SCF ABSORBED	cum. R.Q.
Canister A 455 4755 1033 10955 12255 22815 23485 3485 4035	0.75 0.75 0.72 0.72 0.72 0.72 0.72 0.72 0.72 0.73 0.73 0.73 0.73 0.73 0.73 0.73 0.73	0.10 1.52 2.37 3.80 3.47 5.66 6.53 7.49 8.30 8.30 8.59	0.07 0.62 1.03 1.83 2.68 3.15 2.68 3.35 2.35 4.35 5.59 4.35 5.59	0.05 0.85 1.36 1.71 2.34 2.65 2.89 3.13 3.56 3.56 3.56 4.20 4.39	0.48 0.56 0.57 0.53 0.52 0.52 0.51 0.51 0.551 0.551
Canister B 5 45 105 105 133 1655 2255 2355 2345 3455 3465 435	0.84 0.765 11.276 2.727 2.728 2.728 4.729 4.729 4.556 6.30	0.0728 0.073132 0.073132 0.073132 0.03555 0.03563 0.03563	0.623443780011.83782835535935599	0.98 1.845 1.843 2.461 2.461 2.77 2.801 2.98 3.96 3.96 3.55	0.478 00.449 00.55555555555555555555555555555555555
Canister C 12 51 80 115 140 170 200 230 262 290 320 350 377 408 440	001122355660778359	0.451990402228813341660 2.5344.91228813341660 6.666666666666666666666666666666666	0.17 0.70 1.10 1.58 1.92 2.75 3.69 4.82 5.61 4.82 5.66	012233750 1444555555555555555555555555555555555	0.62 0.66 0.70 0.73 0.79 0.80 0.80 0.80 0.81 0.81
Canister MSA					
12 51 80 115 120 1700 230 260 260 350 3708 440	0.0643556667 0.01112233344556667	0123643018018065891571 22333333344444444444444444444444444444	0.17 0.70 1.10 1.53 2.717 3.45 2.717 3.44 2.717 3.44 5.161 5.65	0.178 1.142 1.454 1.73 1.904 1.23 2.338 2.484 2.484 2.484	0.49 0.55 0.55 0.55 0.55 0.55 0.55 0.55 0.5

Table 27 Cumulative Data, P-3

TIME MIN.	OXYŒ CONSUMED	en scf Generated	CARBON DI PRODUCED	OXIDE SCF ABSORBED	CUM. R.Q.
Canister A 0 15 55 77 107 135 165 165 225 2255 2250 315 345 3475 405	0.2526 0.2926 1.280 2.277788 4.270 2.334 4.5556 6.30	0.29 0.851 2.558 4.502 6.66 7.597 8.88 9.03	0.21 0.70 1.47 1.86 2.68 3.10 3.85 3.75 4.17 5.15 5.15 5.15 5.15	0.15 0.89 1.66 2.02 2.48 3.11 3.67 3.97 4.22 4.60 4.75	-5087655555555555555555555555555555555555
Canister B 0 15 75 75 107 135 165 195 225 225 2315 345 375 405		-336 12.731 23.731 33.733 4.733 55.54663 1730 6.30	0.26 0.46 11.86 11.88 15.83 15.53 15.59 15.59	0.16 0.88 1.16 1.54 1.83 2.37 2.56 2.786 2.786 2.786 3.15	-4800 0.5549 0.555555555555555555555555555555555555
Canister C 0 30 65 83 110 140 170 200 230 260 290 320 350 350 410 440	0.509 1.0395 1.0	0.151 0.151 2.593 2.593 3.583 4.023 4.550 4.681 4.93 5.06	194 194 1593715990 1111225948360 111225948360	-748 -57189 -1697 -21697 -21697 -21697 -21699 -1458 -1697 -1	0.77 0.83 0.87 0.92 0.97 1.04 1.06 1.07 1.11 1.12 1.13
Canister MSA 0 30 65 83 110 140 170 200 230 260 290 320 350 385 410 440	0.099 11.835 555 667 11.838 3837 455 668 7	012.593 012.593 012.593 012.593 012.593 012.50 012.50 012.50 012.50 012.50 012.50	0.41 0.44 0.14 1.51 1.59 2.77 2.76 3.58 94 4.36 5.66 5.66 5.66	0.2768 0.768 0.9255 1.465 1.866 1.103 1.033 1.03	-446789901346899122 0000000000000000000000000000000000

Table 28 Cumulative Data, P-4

TIME HIN.	OXYO CONSUMED	en scf Cenerated	CARBON DI PRODUCED	OXIDE SCP ABSORBED	cum. R. Q.
Canister A					
0 20 655 140 180 218 242	0.34 1.09 1.60 2.35 3.02 3.66	0.23 1.24 1.89 2.83 3.56 4.17 4.53	0.28 0.89 1.31 1.93 2.48 3.00 3.33	0.14 0.70 0.99 1.40 1.77 2.11 2.30	0.62 0.56 0.52 0.49 0.50 0.51
Canister B 0 20 65 95 140 180 218 242	0.34 1.09 1.60 2.35 3.66 4.06	0-25 1.31 1.96 2.76 3.76 4.00	0.28 0.89 1.31 1.93 2.48 3.00	0.15 0.78 1.12 1.55 1.88 2.18 2.35	0.63 0.59 0.57 0.56 0.56 0.58 0.59
Canister C 50 90 135 175 213 237	0.08 0.84 1.51 2.27 2.94 3.58 3.98	0.14 2.01 3.11 4.11 4.76 5.23 5.47	0.07 0.69 1.24 1.86 2.41 2.93 3.26	0.06 0.99 1.80 2.64 3.29 3.84 4.16	0.46 0.49 0.58 0.64 0.69 0.73 0.76
Canister MS# 50 90 135 175 213 237	0.08 0.84 1.51 2.57 2.58 3.98	0.01 0.63 1.40 2.05 2.50 2.82 2.98	0.07 0.69 1.24 1.86 2.41 2.93 3.26	0.06 0.83 1.21 1.51 1.72 1.89	4.18 1.32 0.87 0.73 0.69 0.67

Table 29 Cumulative Data, P-5

TIME MIN.	OXYGE CONSUMED	N SCF GENERATED	CARBON DI		CUM. R.Q.
Canister A	-2				
0 17 47 68 102 132 167 207 227 252 287	0.29 0.79 1.14 1.71 2.22 2.80 3.48 3.48		0.235 0.255 0.94 1.40 1.80 2.85 3.12 3.475 4.36	0.14 0.53 0.73 1.02 1.26 1.53 1.92 2.05	0.47 0.43 0.43 0.39 0.39 0.40 0.41
287 317	4.82 5.32	5.47 5.79	4.36	2.22 2.36	0.41
Canister A					
0 17 47 68 102 132 167 207 227 252 287 317	0.29 0.79 1.14 1.71 2.22 2.80 3.48 3.483 4.82 5.32	0.11 0.48 0.71 1.10 1.42 1.73 1.96 2.01 2.25 2.38	0.23 0.65 0.94 1.40 1.82 2.30 2.85 3.12 3.47 3.95	0.08 0.30 0.42 0.61 0.76 0.91 1.08 1.16 1.24 1.35	0.67 0.62 0.660 0.54 0.53 0.557 0.60 0.61
Canister D	1				
0 22 52 73 107 137 172 209 232 257 292 322	0.37 0.23 0.23 1.80 2.83 2.85 2.85 3.99 4.32 4.94 5.41	0.66673 12.6773 12.6773 14.829 14.829 16.55566	0.30 0.72 1.00 1.47 1.89 2.37 2.88 3.19 3.54 4.02	0.24 0.86 1.25 1.81 2.66 3.34 3.58 3.58 3.90	0.444 0.533 0.558 0.663 0.663 0.663
Canister M	SA				
0 22 73 107 137 172 209 232 257 292 322	0.37 0.87 1.23 1.830 2.89 3.51 3.92 4.90 5.41	0.133 1.586 1.586 2.887 2.131 3.55 3.65	0.30 0.72 1.00 1.47 1.89 2.37 2.88 3.19 3.54 4.02 4.43	0.15 0.52 0.71 0.95 1.12 1.29 1.45 1.63 1.73 1.82	0.45 0.45 0.45 0.44 0.44 0.47 0.49 0.49 0.49 0.5

Table 30 Cumulative Data, P-6

TIME	OXYGE	n SCF Generated	CARBON DI PRODUCED	OXIDE SCF ABSORBED	CUM. R.Q.
MIN.	CONSUMED	GEMERATED	INODUCED	ADDOX DOD	
Canister A 14 56 86 116 146 206 266 326 386 446	1-2 0.24 0.94 1.45 12.45 3.46 7.49	0.02 0.20 0.36 0.50 0.63 0.91 1.19 1.96 2.25	0.19 0.77 1.18 1.60 2.81 2.83 3.66 4.49 5.31	0.02 0.13 0.19 0.25 0.32 0.45 0.60 0.78 0.96	0.84 0.63 0.59 0.50 0.50 0.50 0.50 0.50
Canister 1 14 56 86 116 146 206 266 326 386 446	4-4 0.24 0.94 1.45 1.45 1.47 1.47 1.47 1.47	0.01 0.06 0.10 0.13 0.18 0.28 0.36 0.35 0.79	0.19 0.77 1.18 1.60 2.83 3.66 4.49 5.31 6.14	0.02 0.11 0.16 0.19 0.28 0.35 0.45 0.58	2.98 1.90 1.63 1.42 1.22 1.00 0.97 0.88 0.73 0.69
Canister 19 61 91 121 156 216 276 336 456	A-5 0.32 1.02 1.53 2.62 3.63 4.63 5.65 7.66	0.04 0.24 0.39 0.53 0.67 0.90 1.10 1.36 1.66	0.26 0.25 1.66 2.15 2.97 4.62 5.27	0.04 0.20 0.29 0.36 0.45 0.61 0.77 0.97 1.16 1.32	1.00 0.85 0.75 0.69 0.67 0.70 0.71 0.70
Canister 19 61 91 121 156 216 276 336 336 456	MSA 0.32 1.02 1.53 2.62 3.63 4.63 4.65 7.66	0.05 0.27 0.42 0.56 0.69 0.90 1.12 1.42 1.79	0.264 0.264 1.665 1.62 2.980 4.62 2.984 5.27	0.03 0.16 0.24 0.32 0.40 0.53 0.66 0.82 1.01	0.63 0.58 0.577 0.58 0.58 0.598 0.556

Table 31 Cumulative Data, P-7

TIME	oxygen	SCF	CARBON DI		CUM.
MIN.	CONSUMED	GENERATED	PRODUCED	ABSORBED	R.Q.
Canister A	0.25	0.08	0.21	0.02	0.28
∜5 75	0.76 1.26	0.48 0.94	0.62 1.03	0.12 0.21	0.25 0.23
120	2.01	1.61	1.65	0.36	0.22
150 180	2.52 3.02	2.04 2.46	2.06 2.48	0.44 0.53 0.72	0.22 0.21
240 300	4.03 5.04	3.25 3.87	3, 30	0.72	0.22
360 360	6.04	3.07 4.48	4.13 4.95	0.41 1.08	0.24
420	7.05	5.11	5.78	1.22	0.24
Canister A	-6		_		
15 45	0.25 0.76	0.09 0.44	0.21 0.62	0.03 0.13	0.29 0.29
75 120	1.26	0.79	1.03	0.21	0.26
150	2.01 2.52	1.32 1.66	1.65 2.06	0.32 0.41	0.25 0.25
180 240	3.02 4.03	2.00 2.65	2,48	0.50 0.68	0.25 0.26
300	5.04	3.20	3.30 4.13	0.86	0.27
360 4 <b>20</b>	6.04 7.05	3.74 4.31	4.95 5.78	1.03 1.17	0.27 0.27
		**54	5.10	1.1	****
Canister A	0.34	0.17	0.28	0.04	0.22
50 80	0.84 1.34	0.62 1.00	0.69 1.10	0.14 0.22	0.23 0.22
125	2.10	1.54	1.72	0.34	0.22
125 155 185	2.60 3.11	1.90	2,13 2,55	0.43	0.22 0.23
245	4.11	2.24 2.85	3,37	0.51 0.68	0.24
305 365	5.12 6.13	3.35 3.83 4.33	4.70	0.86 1.03	0.26 0.27
365 425	7.14	4.33	5.02 5.85	1.18	0.27
Canister M	SA .				
20 50	0.34 0.84	0.14 0.58	0.28 0.69	0.03 0.13	0.24 0.23
50 80	1.34	1.04	1,10	0.23	0.22
125 155 185 245	2.10 2.60	1.70 2.09	1.72 2.13	0.23 0.38 0.47	0.22 0.23
185	3.11 4.11	2.43 3.03	2.55 3.37	0.56	0.23 0.24
305	5.12	3.44	4,20	0.88	0.26
365 425	6.13 7.14	3.75 4.07	5.02 5.85	1.01	0.27 0.28
,	145.	,	2002		-,

Table 32 Cumulative Data, P-8

TIME MIN.	OXYGE CONSUMED	N SCF GENERATED	CARBON DI PRODUCED	OXIDE SCF ABSORBED	CUM. R.Q.
Contaton					
Canister A	0.22	0.04	0.18	0.03	0.67
ИЗ	0.72	0.25	0.59	0.16	0.65
73 103	1.23 1.73 2.48	0.53 0.89	1.00	0.32	0.60
103	1.73	o.89	1.42	0.49	0.55
148	2.48	1.53	2.04	0.79	0.52
183	3.07	2.00	2.52 2.86	1.03	0.51
208 238	3.49 4.00	2.29	2.00	1.19	0.52
268 268	4.50	2.02	3.27 3.69 4.10	1.36 1.51 1.65 1.78 1.89	0.52 0.52
298	5,00	2.93 3.18 3.38	4.10	1.65	0.52
328	5.51	3.38	4.51	1.78	0.53
358	6.01	3.56	4.93	1.89	0.53
388	6.51 7.10	3.56 3.71 3.96	4.51 4.53 5.84 5.84	1.99	0.53 0.54 0.53 0.54
423	7.10	3.96	5.82	2.11	0.53
468 533	7.86 8.95 9.96	4.30	5.44 7.22	2.31	0.54 0.56
222	0.95	4.70 5.36	7.33 8.16	2.63 3.05	0.50
593 653	10.96	5.36 6.10	8.99	2.95 3.31	0.55 0.54
713	11.97	6.66	9.81	3.65	0.55
Canister A					
13	0.22	0.02	0.18	0.02	1.33
43	0.72	-0.08	0.59 1.00	0.10	1.22
103	1.23	0.14 0.25	1.42	0.15 0.20	1.04 0.80
73 103 148	1.73 2.48	0.51	3 011	0.31	0.60
183	3,07	0.80	2.52 2.86	0.46	0.58
208	3.49	1.01	2.86	0.57	0.56
238	4.00	1.29	3.27 3.69 4.10	0.71 0.85	0.55 0.55
268	4.50	1.55 1.78	3.69	0.85	0.55
298	5.00	1.78	4.10	0.96	0.54
328 3 <u>5</u> 8	5.51 6.01	1.96 2.11	4.51 4.93	1.06 1.15	0.54 0.54
388	6.51	2.25	5-34	1.23	0.54
423	6.51 7.10	2.39	5.34 5.82 6.44	1.32	0.55 0.56
468	6,86	2.57	6.44	1.43	0.56
533	8.95	3.02 3.78	7.33 8.16	1.65	0.54
593 653	9.96	3.78	8.16	1.97	0.52
653 713	10.96 11.97	4.54 5.12	8.99 9.81	2.37 2.73	0.52 0.53
Canister A	-7				
23	0.39	0.17	0,32	0.07	0.39
53 83	0.39 0.89	0.53	0.73	0.25	0.48
83	1.39	0.78	1.14	0.43	0.55
113	1.90	1.03	1.55 2.11	0.58 0.79	0.56
153 186	2.57	1.40 1.70	2.11	0.79	0.57
213	2.57 3.12 3.58 4.08	1.94	2.56 2.93 3.34 3.76	0.98 1.13	0.58 0.59
243 243	4.68	2.20	3,34	1.30	0.59
273	4.58	2.44	3.76	1.30 1.46	0.59 0.60
<u>3</u> 03	5.09	2,66	44.17	1.59	0.60
333	5.59	2.85	4.58	1,71	0.60
363	6.09	2.98	4.99	1.82	0.61
393	6.60	3.10	5.41	1.92	0.62

Table 32 (cont.) Cumulative Data, P-8

TIME MIN.	OXYGEI CONSUMED	N SCF GENERATED	CARBON DI	OXIDE SCF ABSORBED	cum. R.Q.
Canister A 433 473 543 598 658 718	7.27 7.94 9.12 10.04 11.05 12.06	3.30 3.50 3.80 4.21 4.76 5.21	5.96 6.51 7.47 8.23 9.05 9.88	2.03 2.17 2.46 2.74 3.06 3.36	0.62 0.62 0.65 0.65 0.64 0.65
Canister M 23 53 83 113 153 1243 2473 2473 303 333 363 393 473 548 658 718	0.39 0.89 1.39 1.39 2.57 2.57 2.58 4.58 5.09 5.69 7.27 7.94 9.04 11.05	95443320155542463951 0001102202033333334453951	0.32 32 32 31,75 31,55 3	0.05 0.21 0.41 0.63 0.95 1.17 1.33 1.63 1.63 1.86 1.93 2.02 2.13 2.92 3.15	55241888888999901244556 0000000000000000000000000000000000

Table 33 Cumulative Data, P-9

TIME	oxygen scf			CARBON DIOXIDE SCF	
MIN.	CONSUMED	GENERATED	PRODUCED	ABSORBED	R.Q.
Canister A 15 45 75 115 145 205 240 265 310 355 385 445 505	27263334435066789 0011223445556789	0.46 0.46 0.82 1.306 2.813 2.58 935 4.76	0.21 0.62 1.03 1.58 2.00 2.41 2.83 3.65 4.83 4.83 56.95 7.77	0.05 0.20 0.43 0.53 0.63 0.79 0.88 1.01 1.12 1.32 1.46	0.5436 0.43621 0.3331 0.2288 0.2298 0.3315 0.3315
Canister 455 455 1155 12050 2465 2555 565	566334435066789 2729494035066789	0.45 0.45 0.45 0.839 1.713 2.493 2.63 2.63 3.63 4.026 4.56 4.56 4.59 5.33	0.21 0.62 1.58 2.00 2.41 2.30 3.65 4.27 4.830 6.95 7.77	0.05 0.24 0.37 0.55 0.69 0.89 1.00 1.05 1.14 1.23 1.31 1.47 1.61	0.63 0.544 0.40 0.39 0.37 0.34 0.331 0.330 0.312 0.332
Canister 4 20 555 80 115 150 180 210 240 270 315 360 395 455 575	1-7 0.34 2.3934 1.5932 2.3.5053 3.5053 4.5294 5.665 7.8.65	0.23 0.24 0.24 1.76 2.59 2.59 2.57 2.57 2.57 2.57 2.57 2.57 2.59 2.57 2.59 2.59 2.59 2.59 2.59 2.59 2.59 2.59	0.28 0.76 1.58 2.06 2.48 23.30 2.49 4.95 4.95 56.29 7.91	0.08 0.31 0.42 0.59 0.79 0.93 1.06 1.20 1.34 1.52 1.78 1.99 2.11	33344 33333333333333333333333333333333

Table 33 (cont.) Cumulative Data, P-9

TIME HIN.	OXYGE CONSUMED	n scp Generated	CARBON DI PRODUCED	OXIDE SCF ABSORBED	Cum. R. Q.
Canister M	ISA				
20	0.34	0.14	0.28	0.07	0.51
55 80	0.92	0.60	0.76	0.27	0.45
8õ	1.34	0.91	1.10	0.36	0.40
115	1.93	1.31	1.58	0.50	0.39
15Ó	2.52	1.67	2.06	0.66	0.39
115 150 180	3.02	1.93	2.48	0.76	0.39
210	3.53	2.20	2.89	0.86	0.39
240	4.03	2.44	3.30	0.95	0.39
270	4.53	2.61	3.72	1.01	0.39
<b>31</b> 5	5.29	2.87	4.33	1.08	0.38
360	6.04	3.04	4.95	1.13	0.37
395	6.63	3.12	5.44	1.17	0.37
455	7,64	3.33	6,26	1.29	0.39
395 455 5 <b>15</b>	8.65	3.58	7.09	1.41	0.39
575	9.65	3.81	7.91	1.50	0.39

Table 34 Cumulative Data, P-10

TIME MIN.	OXYGE CONSUMED	n scf Generated	CARBON DI PRODUCED	OXIDE SCF ABSORBED	CUM. R.Q.
Canister E					
0 15 45 735 1355 255 315 375	0.25 0.76 1.26 2.27 3.27 4.78 5.29 6.30	0.15 0.73 1.27 2.28 3.14 3.81 4.28 4.45	0.62 0.62 1.03 1.86 2.68 3.51 4.33	0.04 0.19 0.36 0.71 1.05 1.31 1.50	0.23 0.26 0.28 0.31 0.34 0.34 0.34
Canister A	-8				
0 15 75 135 135 255 315 375	0.25 0.76 1.26 2.27 3.28 4.28 5.30	0.14 0.71 1.28 2.42 3.50 4.51 5.53	0.62 0.62 1.03 1.86 2.68 3.51 4.33 5.16	0.03 0.18 0.32 0.62 0.93 1.22 1.50	0.24 0.25 0.25 0.26 0.27 0.27 0.27
Canister F					
0 20 50 140 260 320 380	0.84 1.34 2.335 3.35 3.37 2.37 5.38	0.90 1.57 2.91 4.27 5.94 5.19	0.28 0.69 1.10 1.93 2.758 3.58 4.40 5.23	0.22 0.39 0.74 1.10 1.35 1.48	0.24 0.24 0.25 0.26 0.26 0.25 0.25
Canister C	-2				
655 1255 1855 2455 3055	0.08 1.09 2.10 3.11 4.11 5.12 6.13	0.09 1.59 2.44 3.11 3.24 4.24	0.89 0.89 1.72 2.55 3.37 4.20 5.02	0.01 0.32 0.61 0.88 1.15 1.41 1.67	0.16 0.20 0.25 0.28 0.31 0.33
Canister M	3A				
0 20 50 80 140 200 260 320 380	0.34 0.84 1.34 2.35 3.36 4.37 5.37	0.15 0.55 0.95 1.724 2.80 3.62	0.69 1.10 1.93 2.75 3.58 4.40 5.23	0.04 0.14 0.24 0.46 0.66 0.82 0.95	0.24 0.25 0.26 0.27 0.28 0.29 0.29

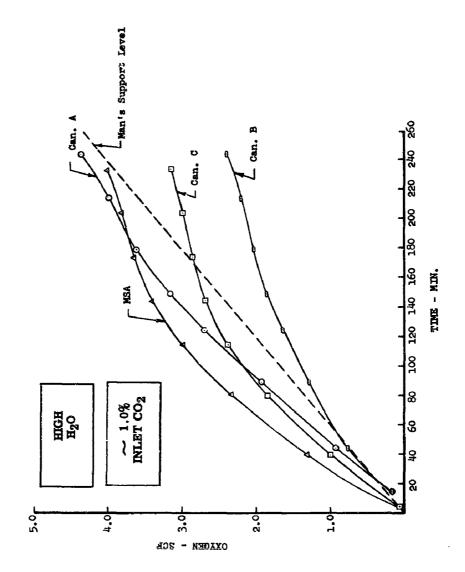
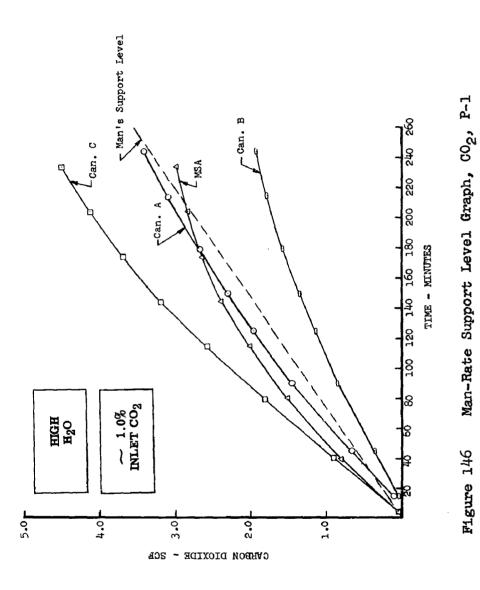


Figure 145 Man-Rate Support Level Graph, O2, P-1



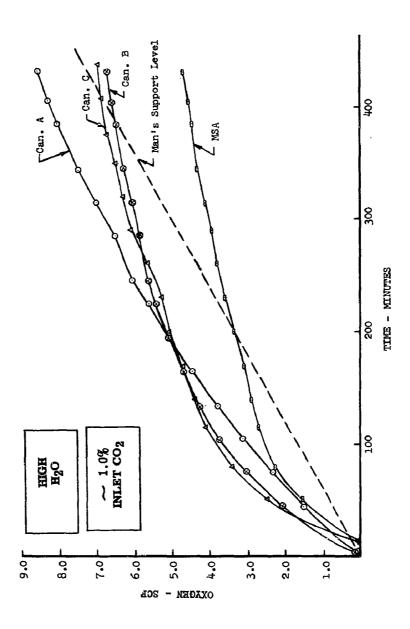


Figure 147 Man-Rate Support Level Graph, O2, P-2

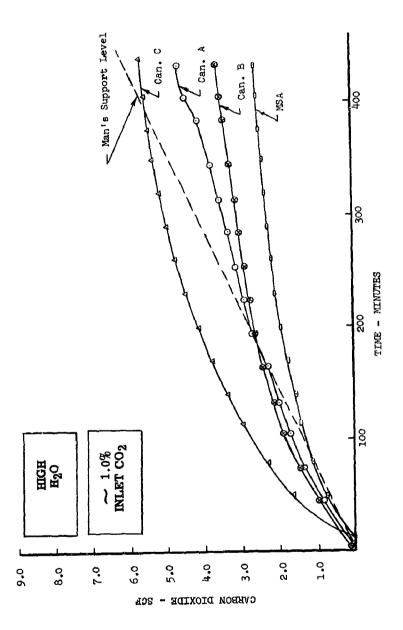


Figure 148 Man-Rate Support Level Graph, CO2, P-2

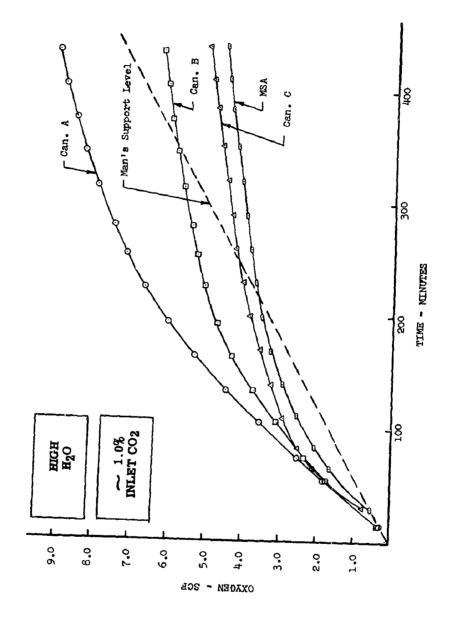


Figure 149 Man-Rate Support Level Graph, O2, P-3

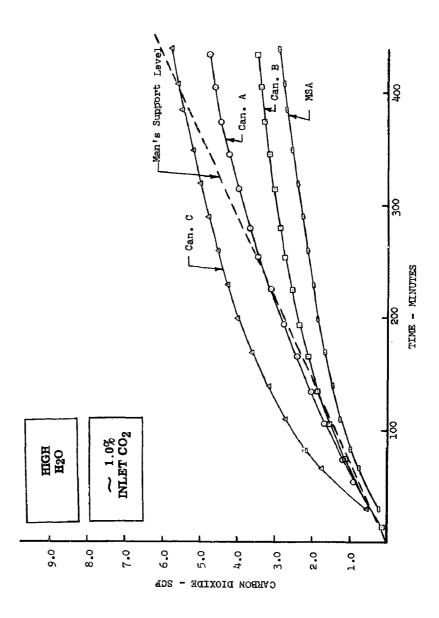


Figure 150 Man-Rate Support Level Graph, CO2, P-3

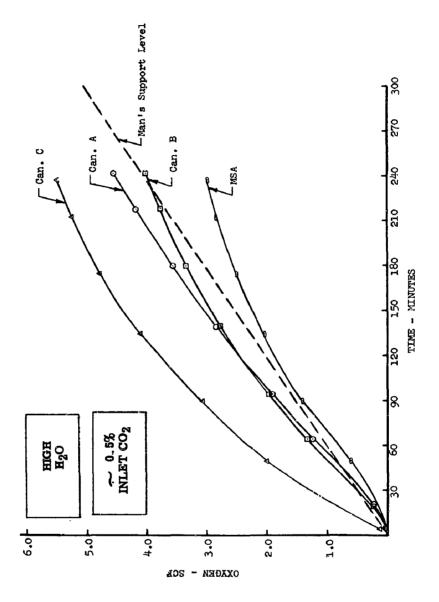
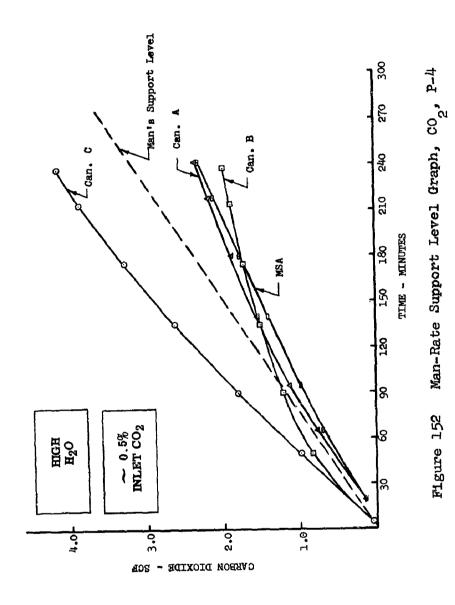


Figure 151 Man-Rate Support Level Graph, 02, P-4



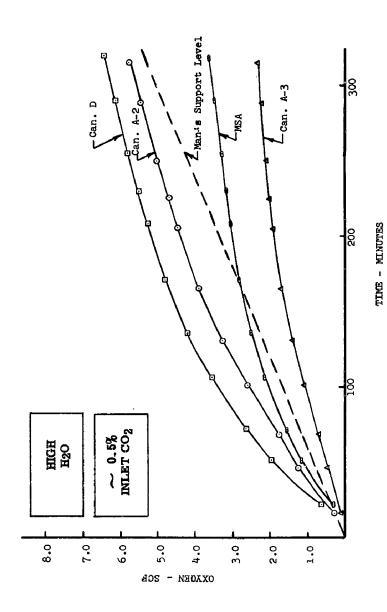


Figure 153 Man-Rate Support Level Graph, 02, P-5

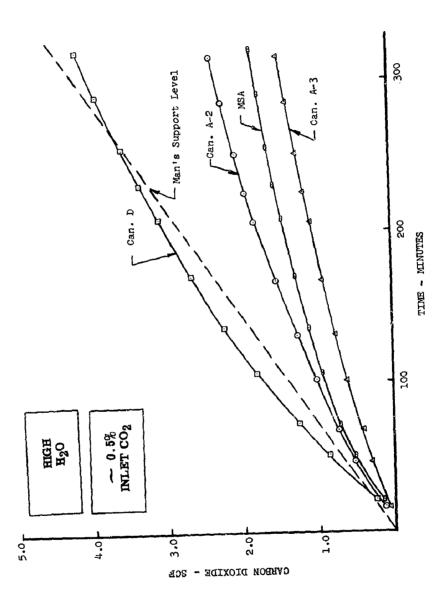
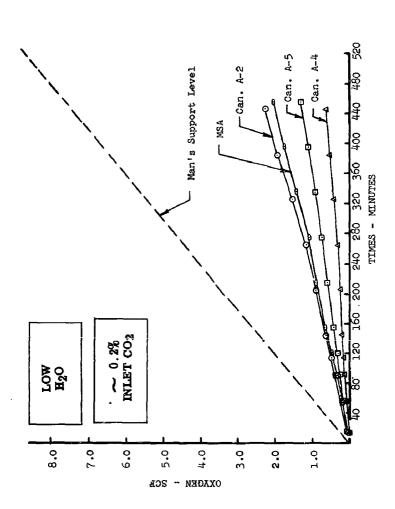
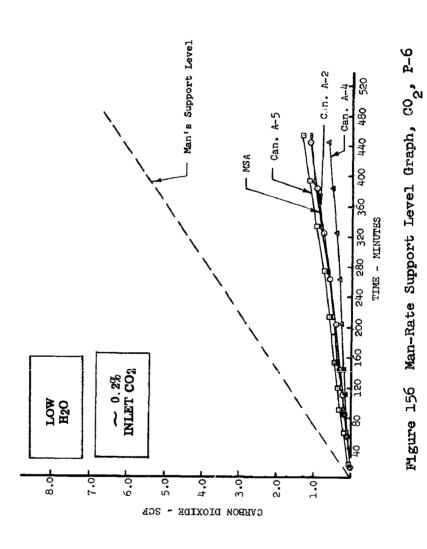


Figure 154 Man-Rate Support Level Graph, CO2, P-5





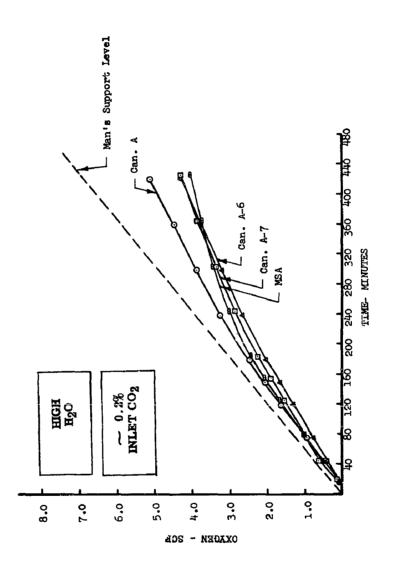


Figure 157 Man-Rate Support Level Graph, 02, P-7

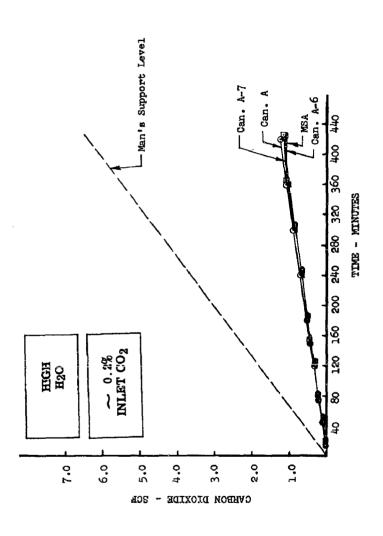


Figure 158 Man-Rate Support Level Graph, CO, P-7

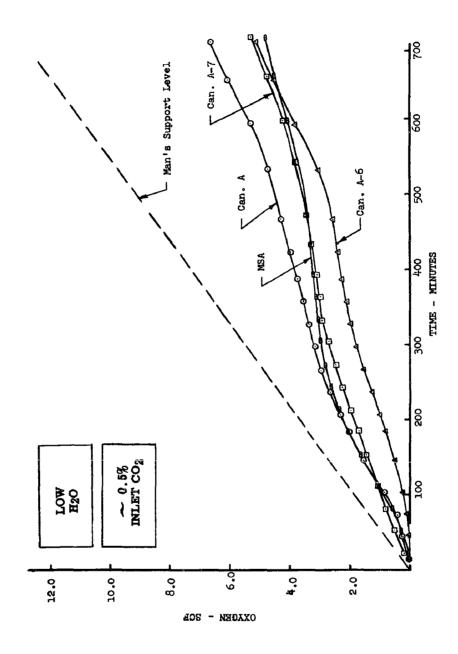


Figure 159 Man-Rate Support Level Graph, 02, P-8

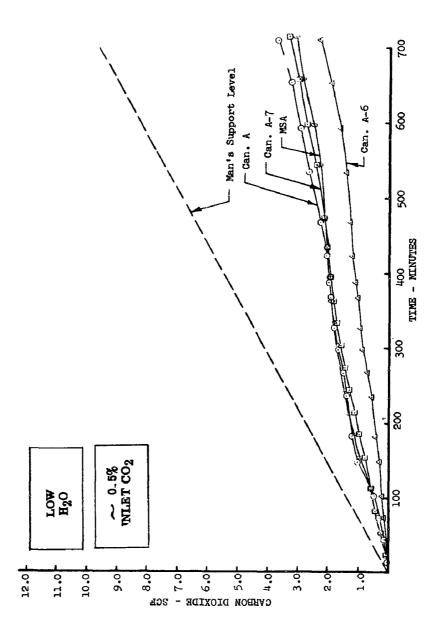


Figure 160 Man-Rate Support Level Graph, CO, P-8

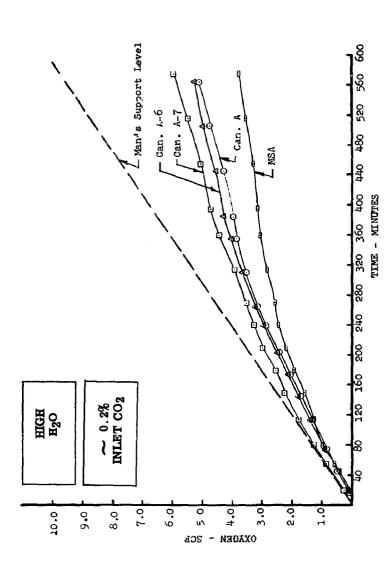


Figure 161 Man-Rate Support Level Graph, O2, P-9

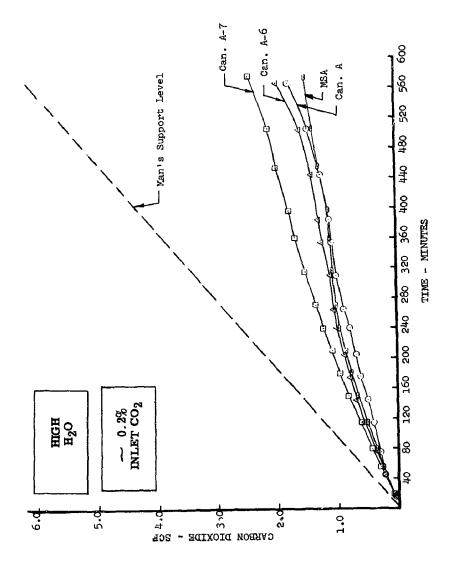


Figure 162 Man-Rate Support Level Graph, CO2, P-9

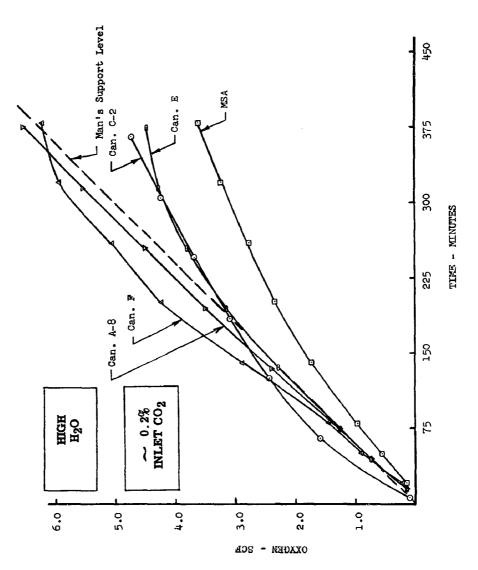
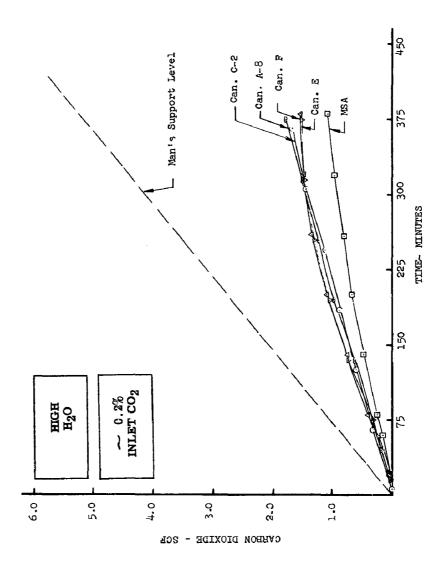


Figure 163 Man-Rate Support Level Graph, O2, P-10



Hgure 164 Man-Rate Support Level Graph, CO2, P-10

Ppt No. ASD-TDR-62-583, POTASSIUM SUPEROXIDE CANISTER EVALUATION FOR MANNED SPACE VEHICLES. mechanics, Flight Accessories Laboratory, Aeronautical Systems Division, Dir/Aero-Final report, Sep 62, 275p, incl illus, Wright-Patterson Air Force Base, Ohio. tables, 66 refs.

Unclessified Report

II. Contract AF33(616)-I. AFSC Project 6146

rask 614608

III. North American Aviation, Inc. A. W. Opticen

composition control

5. Atmospheric

removal

space vehicles, are presented. Actual exper-iments using rodents, men, and simulated-men, are described and compared. Experimental peremeters on potassium superoxide ( $KO_2$ ) canisters for life support systems in manned effects of simultaneous multiple operating results, used to develop and evaluate the New experimental techniques, and their

V. NA-62-283 VI. Not evel fr OTS

V.

collection

In ASTIA

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basis. Dominant role of PCO2 in establishing the CO<sub>2</sub> adsorption rate and the O<sub>2</sub> generation rate is shown as well as roles of absolute humidity and catalysts in establishing the O<sub>2</sub> generation rate. Empirical induction was used to try and substantiate previously device for a chemical canister which prevents canister sizes for atmosphere composition control system tests on a real-time, one-man characteristics of single and duel canister blocking of the airflow in a granuler solid method was developed to determine basic KO2 accepted theoretical equation of a KO2 bed. perimental definitive Respiratory Quotient bed due to high absolute humidity systems are analyzed. Best method for exmatching is described. New annular screen Canister design (demend) stmosphere composition control is shown and described. chemical

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mechanica, Flight Accessories Leboratory, Wright-Patterson Air Force Base, Ohio. Rpt No. ASD-IDR-62-583, POTASSIUM SUPEROXIDE CANISTER EVALUATION FOR MANNED SPACE VEHICLES. Aeronautical Systems Division, Dir/Aero-Final report, Sep 62, 275p, incl illus, tables, 66 refs.

¢

Environmental control

1. Breathing apperatu 2. Environmental cont 3. Oxygen equipment 4. Carbon dioxide

Breathing apparatus

Unclessified Report

space rehicles, are presented. Actual exper-iments using rodents, men, and simulated-men, are described and compared, byperimpatel parameters on potessium superoxide ( $KO_2$ ) canisters for life support systems in manned effects of simultaneous multiple operating results, used to develop and evaluate the New experimental techniques, and their

/ ( over )

basis. Dominant role of PCO2 in establishing the CO<sub>2</sub> adsorption rate and the O<sub>2</sub> generation rate is shown as well as roles of absolute humidity and catalysts in establishing the v2 generation rate. Empirical induction was used to try and substantiate previously as matching is described. New annular screen device for a chemical canister which prevents control system tests on a real-time, one-man method was developed to determine basic EO2 blocking of the sirflow in a granular solid accepted theoretical equation of a NO2 bed. characteristics of single and duel canister chemical bed due to high absolute humidity (demend) stmosphere composition control systems are analyzed. Best method for experimental definitive Respiratory quotient canister sizes for atmosphere composition is shown and described. Canister design

**Environmental** control Breathing apparetus

Φ

Oxygen equipment Cerbon dioxide

5. Atmospheric removal

II. Contract AF33(616)composition control I. AFSC Project 6146 Task 614608

III. North American Avistion, Inc.

IV. A. W. Optican V. NA-62-283

VI. Not sval fr OTS collection In ASTIA

Wright-Petterson Air Force Base, Ohio. Rut No. ASD-TDR-62-583, FOTASSIUM SUPEROXIDE CANISTER EVALUATION FOR MANNED SPACE VEHICLES. mechanics, Flight Accessories Laboratory, Aeronautical Systems Division, Dir/Aero-Final report, Sep 62, 275p, incl illus, tables, 66 refs.

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III. North American

composition control

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space vehicles, are presented. Actual experiments using rodents, men, and simulsted-men, are described and compared, Experimental canisters for life support systems in menned effects of simultaneous multiple operating results, used to develop and evaluate the parameters on potassium superoxide (KO2) New experimental techniques, and their

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Environmental control

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Oxygen equipment Carbon dioxide

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